

AperTO - Archivio Istituzionale Open Access dell'Università di Torino

Block copolymers for the synthesis of pure and Bi-promoted nano-TiO₂ as active photocatalysts

This is the author's manuscript

Original Citation:

Availability:

This version is available <http://hdl.handle.net/2318/122708> since 2016-10-07T15:59:33Z

Published version:

DOI:10.1007/s11051-012-1086-z

Terms of use:

Open Access

Anyone can freely access the full text of works made available as "Open Access". Works made available under a Creative Commons license can be used according to the terms and conditions of said license. Use of all other works requires consent of the right holder (author or publisher) if not exempted from copyright protection by the applicable law.

(Article begins on next page)

This is the author's final version of the contribution published as:

D. Meroni; V. Pifferi; B. Sironi; G. Cappelletti; L. Falciola; G. Cerrato; S. Ardizzone. Block copolymers for the synthesis of pure and Bi-promoted nano-TiO₂ as active photocatalysts. JOURNAL OF NANOPARTICLE RESEARCH. 14 pp: 1086-1100.
DOI: 10.1007/s11051-012-1086-z

The publisher's version is available at:

<http://www.springerlink.com/index/pdf/10.1007/s11051-012-1086-z>

When citing, please refer to the published version.

Link to this full text:

<http://hdl.handle.net/2318/122708>

Block copolymers for the synthesis of pure and Bi-promoted nano-TiO₂ as active photocatalysts

D. Meroni¹ V. Pifferi,¹ B. Sironi,¹ G. Cappelletti,¹ L. Falciola,¹ G. Cerrato² and S. Ardizzone^{1*}

¹*Dipartimento di Chimica Fisica ed Elettrochimica, Università di Milano, Via Golgi 19, 20133
Milano, Italy, tel. +39/0250314225, fax: +39/0250314228, e-mail: silvia.ardizzone@unimi.it*

²*Dipartimento di Chimica IFM & NIS-Interdept. Excellence Centre, Università di Torino, via
Pietro Giuria, 7, 10125 Torino, Italy*

** Corresponding author*

Abstract

Nano-TiO₂ is certainly the most studied semiconductor for environmental purposes. Here, a template synthesis using block copolymers is combined to Bi promotion to the purpose of enhancing the TiO₂ photocatalytic activity by modulating the oxide surface area and porosity as well as by slowing down the electron-hole recombination. Three block copolymers of the Pluronic family, characterized by different micelle size in water as determined by light scattering analysis, are employed to induce mesoporosity in nano-TiO₂. The surfactants are removed by combining UV and thermal treatments in order to avoid pore collapse while obtaining a good oxide crystallinity. A fine modulation of pore size and total volume is obtained by changing polymer type and concentration, effectively enhancing the ability of the oxide to promote the mineralization of methylene blue stains. The mesoporous oxides are then used as scaffolds to obtain Bi₂O₃-TiO₂ composites. X-ray diffraction, N₂ adsorption at subcritical temperatures, high resolution transmission microscopy, Fourier transform infrared spectroscopy, and zeta potential determinations give insight on the composite structure and on the specificity of the Bi-mesoporous TiO₂ composites with respect to traditional sol-gel TiO₂ nanomaterials. All samples are tested for the photodegradation of methylene blue stains and of formic acid under dry and wet conditions, respectively. The presence of Bi promotes the photocatalytic activity of the final samples in both tested reactions (about 30% mineralization enhancement with respect to unpromoted TiO₂). The top performing photocatalyst is the Bi₂O₃-mesoporous TiO₂, which shows the largest recombination time of photogenerated electrons and holes as determined by photocurrent measurements.

Keywords

Pluronic block copolymers, mesoporous nano-TiO₂, Bi-promoted titania, photocatalysis

Introduction

In the past decade, there has been a great interest in nanocrystalline titania materials. The photoelectronic properties of nanotitania are very promising for environmental or clean energy applications, such as photovoltaics and photocatalysis (Ardizzone et al. 2008; Crepaldi et al. 2003; Meroni et al. 2011). Crystalline structure and particle morphology are responsible for an efficient light semiconductor interaction (Angelome et al. 2007).

Mesoporous TiO₂ materials with tailored pore size, high specific surface area, and well-defined crystalline frameworks have, therefore, great potential applications in these fields. The pore size and specific surface area greatly impact the physical properties of a mesoporous material and can play a relevant role in its activity (Bleta et al. 2010). For instance, Saadoun et al. (2000) reported that mesoporous anatase powder with larger pore diameter shows a higher photocatalytic activity, than the same system with smaller pores, for the degradation of formaldehyde under UV illumination. Even though formaldehyde is a small molecule, large pore diameters could provide a faster reaction pathway.

The synthesis of ordered mesoporous oxides goes back to the beginning of the nineties when low molecular weight cationic surfactants were employed as structure directing agents for the synthesis of mesoporous silica. In time this approach was extended to numerous systems by using both anionic and non-ionic surfactants (Bates and Fredrickson 1999). A critical point of this kind of syntheses, especially in the case of titanium dioxide, is the control of the nucleation and growth processes and consequent crystallization taking place within the inorganic walls upon thermal treatment. When crystal size exceeds wall thickness, the continuous mesoporous structure usually collapses, leading to a fragile packing of nanocrystallites. A growing number of works was dedicated to the optimization of conditions leading to the production of ordered mesoporous nanotitania with efficient photocatalytic or photovoltaic performances. To this aim recently block copolymers have been consistently proposed (Wu and Rankin 2011). Block copolymers are made of

two or more chemically distinct polymer blocks covalently bonded together (Alexandridis and Hutton 1995; Taylor et al. 2007). These blocks may be thermodynamically incompatible with each other supporting micro-phase separation on the molecular scale (in the nanometers range) to produce complex nanostructures with various morphologies depending on the relative volume fraction of one block with respect to the other (Tadros 2009).

The large majority of preparations of ordered mesoporous TiO₂ by block copolymers follows the evaporation-induced self-assembly (EISA) process, in which the polymers are introduced in the reaction mixture simultaneously with the titanium precursor (Brinker et al 1999; Choi et al 2004). Thus, condensation of the metal ions and self-assembly of the block copolymer templates occur simultaneously.

A novel method, termed the “nanoparticle route” (Bleta et al. 2010, Bosc et al. 2003), appears to be promising for the synthesis of stable mesostructured titania. In comparison with the classical EISA process, in the nanoparticle route, the copolymer is added to the titania nanoparticulate sol, and only the self-assembly process between copolymer and nanoparticles is allowed to take place. By this approach collapse problems that can be associated with template removal and crystallization may be reduced (Sanchez et al. 2008). Although the nanoparticles route method may imply a larger number of synthetic intermediate steps, it is inherently more flexible as every separate step can be tailored to produce the desired effect (promote surface area, given pore size, etc.). Further, pre-synthesized nanoparticles are used allowing a fine control of the material features (*e.g.*, lattice location of dopants) (Wang et al. 2010).

In this work, a modified nanoparticle route is adopted. The TiO₂ un-calcined nanoparticles, separately prepared, are mixed with the selected copolymer in water. No mesophase of copolymer is allowed to form; since no long range order of mesostructure is sought. The desired mesoporosity will, actually, be the result of the pores (voids) created into the oxide matrix by the micellar structures, decomposed during the calcination treatment. Three different block copolymers of the Pluronic family are employed and their self-aggregation behavior in water is studied. The role

played by the different block molecular structure on the final morphological features of the oxide is investigated. The effect of mesoporosity on the resulting photocatalytic performances of the mesoporous oxides is examined, both directly onto the “dry” oxide and in solution. Further, since the control over the morphological features of the host oxide may give rise to composite (*i.e.*, mixed) oxides with improved features, mesoporous TiO₂ samples, obtained by the present procedure, are employed to obtain Bi-promoted TiO₂ samples. These latter materials are recently raising great interest in several fields like photocatalysis, photovoltaics, and hydrogen production (Jing et al. 2009; Shamaila et al. 2010). Indeed, Bi₂O₃ is a non-toxic and environmentally friendly material which, thanks to its band structure (Bian et al. 2008), could trap photo-generated electrons, and thus improve the overall quantum efficiency of the material. In fact, theoretical calculations have shown that the specific band structure of Bi₂O₃-TiO₂ (Long and English 2009) significantly improves the oxide photocatalytic efficiency. Bi₂O₃ can thus represent a cheaper alternative to noble metals in enhancing the photocatalytic performances of TiO₂. Such theoretical results are supported, in the literature, by photoluminescence experiments (Shamaila et al. 2010) which showed a slowed radiative recombination process of photogenerated electrons and holes in Bi₂O₃-TiO₂ composites. However, no direct evidence of the actual reduced recombination rate can be found in the literature. Here, photocurrent measurements of Bi₂O₃-TiO₂ composites are performed in order to assess any effect of the Bi addition on the fate of the photogenerated electron-hole pair.

The use of mesoporous titania as a scaffold for Bi₂O₃ formation is an issue scantily investigated in the literature (Shamaila et al. 2010). Mesoporosity could enhance the dispersion of heteroatoms/oxide particles, determining a further improvement of the photocatalytic performances. In this work, the mesoporosity of the oxide is used to obtain highly dispersed Bi₂O₃-TiO₂ composites with enhanced photocatalytic activity. Their structural features and activity are compared with Bi₂O₃-TiO₂ composites, in which the titanium oxide is obtained by a traditional sol-gel reaction.

Experimental section

All of the chemicals were of reagent grade purity and were used without further purification; doubly distilled water passed through a Milli-Q apparatus was used to prepare solutions and suspensions.

The Pluronics were obtained by courtesy of BASF Company (USA).

Synthesis of Mesoporous Titania. Titanium dioxide nanoparticles were synthesized by mixing the TiO₂ nanoparticle sol with a Pluronic copolymer in water.

The sol synthesis was performed as follows. A solution of 0.1 mol of Ti(OC₃H₇)₄ in 38 mL of 2-propanol was stirred for 10 min at 300 rpm at 25 °C. Then, 180 mL of water was added, dropwise, fast, to the alkoxide solution, in order to obtain a water/alkoxide molar ratio of 100 and a water/2-propanol molar ratio of 20. The slurry was stirred for 90 min in order to complete the hydrolysis. Subsequently, hydrochloric acid (HCl 37%) was added to peptize the hydroxide precipitate in order to obtain an homogeneous milky suspension. The mixture was maintained under reflux at 80 °C for 5 h. The final product was a stable translucent suspension of TiO₂ nanoparticles. Preliminary studies showed that a separate HCl addition followed by a hydrothermal treatment allow to maximize the final sample surface area and porosity.

A suitable amount of the chosen Pluronic (PEO_PPO_PEO, PEO = polyethylene oxide and PPO = polypropylene oxide) triblock copolymer (P104, F127 or P123) was then added to the sol after being liquefied in 2-propanol, and the mixture was maintained under stirring for 3 h at 50 °C. The ratio between the number of ethoxy units and the number of titanium moles (EO/Ti) was varied from 0 to 1.6. Xerogels were obtained after drying the sol by evaporation in a vacuum oven at 40 °C and 40 mbar, and they were calcined at 350 °C in static air for 5 h using a heating ramp of 3 °C min⁻¹. Finally the powders were irradiated with UV light (iron halogenide lamp Jelosil HG 500, irradiation intensity reaching the sample: 30 mW cm⁻²) for 5 h to remove the remaining traces of organic compounds. Samples are identified according to this notation: T_copolymer_(EO/Ti). For example, T_P104_1.2 indicates the use of the P104 copolymer, and an EO/Ti ratio of 1.2. Sol-gel

TiO₂ samples were prepared by a classical sol-gel route (Boiadjieva et al. 2004; Carp et al. 2004) for the sake of comparison.

Synthesis of Bi-TiO₂ composites. Bi-TiO₂ materials were prepared by a direct impregnation method (Di Paola et al. 2002; Shamaila et al. 2010). Two calcined bare TiO₂ samples (a classical sol-gel sample, Tsg, and T_P104_1.2) were treated with a suitable amount of 0.1 M Bi(NO₃)₃ acidic solution. The starting molar ratio Bi/Ti was 0.02, on the grounds of previous literature results (Shamaila et al. 2010). Then, the powders were dried in oven and subsequently calcined at 350 °C in static air for 5 h using a heating ramp of 3 °C min⁻¹.

Samples Characterization. Room-temperature X-ray powder diffraction (XRPD) patterns were collected between 10 and 80° with a Siemens D500 diffractometer, using Cu K α radiation. Rietveld refinement has been performed using the GSAS software suite and its graphical interface EXPGUI. The average diameter of the crystallites, d , was estimated from the most intense reflection (101) of the TiO₂ anatase phase using the Scherrer equation.

Specific surface areas were determined by the classical BET procedure using a Coulter SA 3100 apparatus. Desorption isotherms were used to determine the pore size distribution using the Barret-Joyner-Halander (BJH) method.

Diffuse reflectance spectra (DRS) of the powders were measured on a UV-vis scanning spectrophotometer (Perkin-Elmer, Lambda 35) equipped with a diffuse reflectance accessory. A “total white” Perkin-Elmer reference material was used as reference.

Critical micelle concentrations in water of the employed block copolymers were determined at room temperature by surface tension measurement using the DuNouy method (TSD, Gibertini).

Micelle size distribution of different block copolymers was evaluated by Malvern Nano Zetasizer (ZEN 1600).

The temperature at which the complete decomposition of organics occurs in the TiO₂ xerogel was determined by TGA analysis (Perkin Elmer TGA 7).

Particles morphology was obtained by high resolution transmission electron microscopy (HR-TEM), using a JEOL JEM 3010UHR (300 kV) microscope fitted with a LaB₆ single crystal filament and an Oxford INCA Energy TEM 200 energy dispersive X-ray (EDX) detector. All samples were dry deposited on Cu “holey” carbon grids (200 mesh).

Scanning Electron Microscopy (SEM, HITACHI TM-1000) equipped with energy-dispersive X-ray spectroscopy (EDX, Hitachi ED3000) was employed to estimate the Bi content in the composite samples.

Electrophoretic mobility and zeta potential data were obtained by a Coulter DELSA apparatus. The ionic strength of the suspensions was kept constant by KNO₃ 10⁻² M. Measurements were performed at pH 8 by a diluted KOH solution.

FTIR spectra (128 scans, 4 cm⁻¹ resolution) were collected on a Bruker IFS 113v spectrometer, equipped with MCT detector. Samples were pressed into self-supporting pellets (approx 10 – 15 mg cm⁻²) and placed in a quartz cell fitted with KBr windows. Prior to any FTIR measurements, all samples were activated in vacuo at either room temperature (RT) or at 250°C, connecting the quartz cell to a conventional high vacuum glass line capable of a residual pressure <10⁻⁴ Torr (1 Torr = 133.32 Pascal).

The photocatalytic activity of the samples was tested towards the mineralization of methylene blue on “dry” TiO₂ films and of formic acid in liquid medium, according to procedures previously described (Paoli et al. 2010; Spadavecchia et al. 2010). Films for formic acid degradation were prepared by depositing 0.1 g TiO₂ onto a rough Al substrate according to Ardizzone et al. 2011. Photocatalytic tests were carried out in an 850 mL thermostatic reactor containing 5 · 10⁻³ M formic acid solution. All degradation tests were carried out at 20 °C and at spontaneous pH in the experimental set-up under UV-A irradiation (3 h irradiation time). Methylene blue degradation tests were performed using TiO₂ films (0.08 g) deposited over glass slides (Milanesi et al. 2010) and stained with 0.25 mL of a methylene blue solution (0.02 M). The films were irradiated for 7 h under UV-A irradiation. The photon sources for both photocatalytic tests were provided by two 500 W

iron halogenide lamps (Jelosil, model HG 500, 85 mW effective total power) emitting in the 315-400 nm wavelength range (UV-A). The actual mineralization of both formic acid and methylene blue were monitored by chemical oxygen demand (COD) measurements (Spectroquant Merck).

Photocurrent transient measurements (Hagfeldt et al. 1995), were performed in a suitable electrochemical cell with two platinum counter-electrodes placed in line with the working electrode in the middle. The saturated calomel reference electrode was placed orthogonally with a Luggin capillary, in order to minimize the ohmic drop. The electrodes were immersed in 0.5 M aqueous NaCl at a spontaneous pH and the cell was deaerated with nitrogen. The TiO₂ photoanodes were alternatively exposed to UV light (Jelosil HG 500). The photocurrent and the photocurrent transient were measured by a microIII Autolab potentiostat/galvanostat (EcoChemie, The Netherlands), at a constant potential. After a 60 s equilibration time in the dark under nitrogen degassing, dark/light alternations (200 s/100 s) were repeated for at least 6 times, in order to obtain reproducible transient patterns. The working electrodes were prepared depositing a film of 8 thin layers (3 cm² surface area) of titania nanoparticles on an ITO (Indium Tin Oxide) glass slide (by Aldrich), using a spin coating technique. The as-prepared films were finally sintered at 350 °C in air for 1 h.

Results and Discussion

Synthesis of mesoporous TiO₂ samples

The presently adopted synthetic method is based on the occurrence of the self-assembly process between the copolymer and the nanoparticle precursors, separately prepared. The advantage of using preformed TiO₂ nanoparticles is that pre-condensed species exhibit a lower hydrolytic reactivity than molecular precursors, so that the chemistry, self-assembly, and processing parameters are more separated (Sanchez et al. 2008).

Three different copolymers of the Pluronic family (PEO_PPO_PEO, PEO = poly(ethylene oxide) and PPO = poly(propylene oxide)), characterized by diverse compositions and average molar mass, are adopted in template syntheses. The first part of this work, *i.e.*, the optimization of the synthetic

route and the study of the role played by EO/Ti ratio, was performed by using P104 as the copolymer due to its relatively low molecular weight (around 6000) and an intermediate HLB value (HLB = 8). This copolymer has not been previously adopted for the synthesis of TiO₂.

A tailored strategy for the removal of the template, after the synthesis, involving a mild heat treatment followed by UV irradiation, was developed in order to maximize the sample porosity and surface area while allowing a complete removal of the template agent. The TG analysis under N₂ atmosphere of P104 shows that the copolymer is completely decomposed at about 300°C (Fig. 1a). The TG analysis of un-calcined T_P104_1.2 shows a mass loss recorded below 150 °C, which can be attributed to the loss of physically adsorbed water, followed by two more separated mass losses (centered at about 220 and 270 °C) in the region where the polymer decomposition takes place. Although most of the copolymer is removed between 200 and 300 °C, a second much smaller weight loss takes place at around 400 °C (Fig. 1b). The degradation process ends at 420 °C, in agreement with what observed by Bleta et al. (2010) in the case of TiO₂ templated by a different Pluronic polymer (F127). A calcination temperature of 350 °C with slow heating rate, under static air, was presently adopted in order to prevent a pore collapse due to the template removal. These calcination conditions are often employed in the literature (Bleta et al. 2004; Shamaila 2010) to remove surfactants in oxide precursors, however the TG analyses clearly show that not all the copolymer is decomposed at 350 °C. Thus, a subsequent UV irradiation treatment was presently performed to promote a complete template removal avoiding higher temperature induced crystallite growth and a consequent pore collapse. The optimal time of UV irradiation was determined by surface area determinations: after 5 h UV irradiation, the sample surface area and pore volume reach a maximum value that is only negligibly improved by further irradiation. Moreover, HRTEM images (see the following) show that the sample pore walls are highly crystalline and that the crystallites are not aggregated.

The addition of the block copolymer results in a marked change in the morphological features of the oxide (Fig. 2). The reference sample (T), peptized in HCl without the addition of the copolymer and

calcined under the same conditions, shows a N₂ adsorption-desorption isotherm with a type H2 hysteresis loop, which is typical of bottle-neck pores (Barrett et al. 1951; de Boer et al. 1966; Rouquerol et al. 1999). On the contrary, T_P104 samples present type IV isotherms (Fig. 2), characteristic of mesoporous materials (Bleta et al. 2011), with a type H1 hysteresis loop, typical of spherical or pseudospherical agglomerates with homogenous size and cylindrical pores opened at both sides (Barrett et al. 1951; de Boer et al. 1966; Rouquerol et al. 1999). The addition of the copolymer also results in a remarkable increase in specific surface area and pore volume (Tab. 1). Such a large increase is mainly related to a narrowly distributed mesoporosity (Fig. 2, inset), that is almost absent in the reference sample.

Effect of the EO/Ti ratio

Different P104 copolymer concentrations, expressed as the ratio between the number of ethoxy units and the number of titanium atoms (EO/Ti), were tested. All obtained samples are anatase – brookite composites. No rutile phase was found as expected from the adopted low calcination temperatures. The formation of the brookite polymorph usually occurs together with anatase for crystallites of small size (Boiadjieva et al. 2004). The surfactant addition seems to cause a slight enrichment in the brookite phase (Tab. 1, column 2-3) and a slight decrease in the average crystallite diameter (Tab. 1, column 4). Varying the copolymer amount modifies significantly the sample morphological features: at increasing EO/Ti ratios, a steady increase of the sample surface area and pore volume is observed (Tab. 1, columns 5-6). However, the average pore dimension remains almost unchanged (7 nm) for all tested EO/Ti ratio, and a narrow size distribution (65 - 70% pores with a size in the 6-10 nm range) is obtained for all samples.

At the highest tested concentrations, the copolymer removal becomes more demanding and progressively longer UV irradiation times are requested to obtain its complete removal. Therefore, an optimal concentration (EO/Ti 1.2), allowing a fast and complete removal together with high surface area and mesoporosity, was adopted for testing the other two copolymers.

Effect of the copolymer type

Three different Pluronic copolymers were tested in order to investigate the role played by the different block structure on the oxide final features. The adopted copolymers were two widely employed structure directing agents (P123, F127) (Bleta et al. 2010; Bosc et al. 2003; Crepaldi et al. 2003; Sanchez 2008), together with the already described P104. The characteristics of the three copolymers are reported in Tab. 2. The samples obtained using the same EO/Ti ratio but different templating agents show different shapes of N₂ adsorption-desorption hysteresis loops (Online Resources Fig. 1).

During the syntheses, the adopted copolymer concentrations were at least two orders of magnitude higher than the measured critical micelle concentration (Tab. 2, Online Resources, Fig. 2). The copolymers were characterized for their CMC by surface tension measurements (Tab. 2). The values reported in the literature fall in broad ranges and are strongly affected by the method adopted for the CMC determination (Perry et al. 2011; Tsui et al. 2008; Tsui et al. 2010). In the case of P123 the present value closely compares with the values reported by Kadam and coauthors (2011) and by Lee et al. (2011). In the case of F127 both larger and lower CMC values, with respect to the present one, can be found in the literature (Alexandridis et al. 1994; Hecht and Hoffmann 1995; Zana et al. 2006) while for P104 the literature data are very limited and significant comparisons cannot be made.

The rapid solvent evaporation adopted in this study does not allow a mesostructure to be formed. Therefore the pore morphology is determined only by the micelle features. Fig. 3a shows the micelle size distribution for the three polymers as determined by light scattering at the polymer concentrations adopted in the syntheses. Both P104 and P123 present a monomodal size distribution centered around 10-15 nm, in agreement with literature data (Alexandridis et al. 1995b; Lee et al. 2011). Instead, in the case of F127, more than one component can be appreciated besides the main population at small sizes. Actually the sequence shown by the pore sizes in the mesoporous oxides

(Fig. 3b) reflects the sequence among the principal populations of micelle sizes. In the case of F127, only one pore size component is appreciable possibly because the larger micelle sizes lead to pores out of the mesoporosity range. The obtained oxides show a narrowly distributed mesoporosity that reflects the sequence of the micelle dimensions, as determined by light scattering (Fig. 3a).

All obtained samples were tested towards the photocatalytic oxidation of methylene blue. All the mesoporous samples, with the exception of T_P104_0.5, give mineralization degrees higher than those of reference sample T (mineralization percentage 45%) and of the traditional photocatalytic commercial sample Degussa P25 (38%). Fig. 4 reports the mineralization percentage as a function of the sample surface area. The mineralization degree shows a good linear correlation with the sample surface areas, that is irrespective of the type of copolymer employed. Thus the mineralization degree seems, in this case, to be directly affected by the specific surface area and only indirectly by the sample pore size distribution. It is not surprising that the surface area of the photocatalyst plays a pivotal role in the case of self-cleaning of dry surfaces, like in the tested degradation of methylene blue stains. A higher photocatalyst surface area leads to a larger extent of pollutant directly adsorbed at the oxide surface and thus to a faster oxidation as the degradations probably occurs via a direct oxidation at the catalyst surface. Therefore, in this case the diffusion of the pollutant into the photocatalyst pore network is less crucial.

Bi-TiO₂ composites

Due to the environmentally friendly nature of Bi (and of its compounds) and to the favorable semiconductor features of Bi₂O₃, Bi-promoted TiO₂ samples were obtained with the aim to improve their photocatalytic performance. Bi₂O₃-TiO₂ composites were synthesized by using as scaffolds both the mesoporous TiO₂ and also a TiO₂ powder obtained by a classical sol-gel route to highlight possible effects introduced by the mesoporous structure of the host oxide. Comparing the effect of the different adopted copolymers, it emerges (Tab. 1) that the samples treated with P104 have the largest total pore volume: all the T_P104 samples, with the exception of T_P104_0.5, show a higher

pore volume than T_P123 and T_F127 samples. Moreover, P104 leads to larger pores with respect to the other tested copolymers in agreement with the observed unimodal micelle sizes. The P104 copolymer was thus chosen to prepare Bi-doped mesoporous TiO₂ composites.

For all Bi-promoted oxides, peaks relative to separate Bi₂O₃ phases are not appreciable in X-ray diffraction patterns, while Bi species can be observed by EDX in Bi/Ti atomic ratios (corresponding to a weight percentage of about 10%) comparable with those adopted in the synthesis. The presence of Bi slightly promotes the TiO₂ brookite polymorph while no variations on the crystallite sizes are observed (Tab. 3). Although the surface area of the mesoporous sample is larger (about 20%) than the sol-gel one, the effects produced by the Bi addition are opposite for the two oxide synthetic routes. In the case of the sol-gel sample, the presence of Bi produces a slight increase in the specific surface area and an appreciable increase in the total pore volume possibly due to a better dispersion of the TiO₂ particles in the strong acid medium used for the impregnation. Instead, in the case of the mesoporous TiO₂ both the surface area and the total pore volume decrease appreciably. This effect could be traced back to the desired partial filling of the mesopores by Bi₂O₃. Fig. 5 reports the comparison among the N₂ adsorption/desorption isotherms of two bare TiO₂ samples (sol-gel and mesoporous) and of the two relative samples promoted by Bi. In the case of the sol-gel samples the isotherms show the presence of capillary condensation in the mesopores but eventually give rise to the sharp increase of adsorbed volume at relative pressures close to 1, indicating the presence in the samples of both mesoporosity and of a non porous surface (Barrett et al. 1951; de Boer et al. 1966; Rouquerol et al. 1999). The pore shape changes drastically from type H1 in mesoporous samples, typical of cylindrical pores, to H2 in sol-gel samples, representative of bottle-neck pores (Barrett et al. 1951; de Boer et al. 1966; Rouquerol et al. 1999). The presence of Bi does not modify the shape of the isotherms loop but the shift of the T_P104_1.2Bi curve to lower volumes with respect to T_P104_1.2, is a direct indication of a partial filling of the mesopores by the bismuth oxide.

The addition of Bi to the synthesis slightly modifies the shape of DRS curves (Online Resources, Fig. 3). The observed shift of the Bi promoted curve suggests the formation of intra-gap localized levels (Shamaila et al. 2010).

HRTEM (Fig. 6) investigations indicate that all samples are characterized by a small size of the crystallites (5-8 nm), in good agreement with the data obtained by means of X-ray diffraction (see Tab. 1). Smooth edges and high transparency to the electron beam are appreciable as well as a good crystallinity, as indicated by the high incidence of fringe patterns, generated by the presence of crystalline planes exhibiting low hkl indexes. A detailed inspection of the fringes evidences that for all samples the particles exhibit the same type of crystal planes. The lattice spacings indicate as main termination the (101) crystal plane of anatase (JCPDS file n. 21-1272), with a less frequent presence of the (111) crystal plane of brookite (JCPDS file n. 2-514) (Xu and Zhang 2009). In any case, but more frequently for the systems synthesized employing the P104 copolymer, the presence of Bi is linked to a general underestimation of the lattice distances ascribable to the (111) crystal plane of the brookite phase. Although no clear diffraction pattern relative to bulk Bi₂O₃ can be appreciated in both HRTEM and XRD, the above mentioned distortion of the lattice distances can be most likely attributed to the presence of a highly nanodispersed oxide phase containing Bi species. This conclusion is also supported by other indirect determinations described in the following.

In order to characterize the surface terminations present onto the nanoparticles of these various TiO₂-based systems, FTIR spectroscopy of the powders themselves has been resorted to. All the results have been reported in Fig. 7, in which the full lines refer to a plain vacuum activation at RT (namely, IR beam temperature, some 30°C higher than RT), whereas the bold lines refer to a vacuum activation carried out at 250°C (in order to get rid of all physisorbed and/or weakly adsorbed surface species). Some spectral features deserve comments, as reported in the following:

1. For all spectra relative to the plain RT evacuation, a broad envelope is present in the 2500-3500 cm⁻¹ range. This is not surprising, as at the surface of all these TiO₂-based systems a huge

amount of OH species interacting by H bonding are present and this envelope represents the ν_{OH} mode of all these species (Little 1966). The spectroscopic δ_{HOH} counterpart of these species is represented by the band located at $\sim 1620 \text{ cm}^{-1}$ and indicated by an arrow in the Fig.. All the above spectral components are absent (or drastically reduced) after the activation at $250 \text{ }^\circ\text{C}$, as this thermal treatment carries out a dehydration (up to a medium degree) of the oxidic surface.

2. At higher frequency ($\nu > 3600 \text{ cm}^{-1}$) an envelope of singled out components is evident in all cases: these components are better observable in the bold curves, when the surface has reached a medium degree of dehydration. In fact, all these bands are ascribable to the ν_{OH} modes of OH species free from H-bonding interaction (Morterra et al. 1989) and are better observable in the blown-up section reported in the inset to Fig. 7. In particular, it can be noted that, for the plain TiO_2 system (see the uppermost curve in the inset) obtained by sol-gel synthesis, at least three different components can be singled out. When Bi species are added, the three components are still observable, even though with a different intensity ratio. On the other hand, in the case of the systems obtained employing the P104 copolymer, there is a general lowering of the quality of the spectra (see the lower couple of curves in the inset), much more marked when Bi species are present. In the latter case, only the highest component is still observable and its intensity represents only a small fraction of the original band. These observations allow us to draw some preliminary conclusions about the location of Bi species; in particular, their presence seems to be linked to the covering and/or disappearance of some specific band ascribable to OH species. This may indicate that the grafting of Bi species is most likely due to the “consumption” of surface OH species in the TiO_2 matrix.

3. In the cases of the pure TiO_2 oxides, no matter the synthetic route followed for the preparation, some peculiar absorption located in the $3000\text{-}2750 \text{ cm}^{-1}$ range are evident. These are ascribable to the ν_{CH} mode of hydro-carbonaceous species chemisorbed at the surface of the

oxide as a result of the synthetic route and/or to the exposure to the atmosphere after the calcination treatment (Magnacca et al. 2003).

The interfacial electrification features of the oxides were analyzed by zeta potential measurements in aqueous suspensions. At pH 8 the Bi composites show zeta potential values which are less negative than those of the un-doped TiO₂ samples (Tab. 3). Such effect can be traced back to a shift of the isoelectric point (i.e.p.) in the alkaline direction, with respect to the i.e.p. of bare TiO₂ (pH ≈ 6) due to the presence of Bi₂O₃. Actually the isoelectric point of pure Bi₂O₃ is located at pH 9 (Kosmulski et al. 2009). The effect is larger in the case of the sol-gel sample, possibly due to a larger surface extension of the Bi oxide as suggested by the larger surface area of T_Bi with respect to both T_{sg} and T_P104_1.2_Bi.

Bi₂O₃-TiO₂ composites photocatalytic activity

The photocatalytic activity of Bi₂O₃-TiO₂ composites was tested with respect to the mineralization of both MB deposited onto the dry oxide and of formic acid in the liquid phase (Tab. 4). The mineralization (total degradation) of MB to CO₂ for the sol-gel sample is lower than in the case of the mesoporous one. In both cases the promotion by Bi produces an increased mineralization.

In liquid phase all samples show a mineralization degree larger than 80% at 3 h reaction time while conductivity variations allow differences among the various samples to be appreciated (Tab. 4, third column). The slope of the χ vs t plot represents the rate of conductivity decrease due to the progressive disappearance of formic acid. The slope increases for both Bi promoted samples, the more so in the case of T_P104_1.2_Bi.

On the grounds of what reported above Bi₂O₃-TiO₂ composites show much higher photocatalytic activity than bare TiO₂ both as a dry oxide and in liquid phase tests. It may be relevant to recall that, while in the case of the sol-gel sample the addition of Bi induces an increase in the oxide surface area, thus producing a larger extension of photocatalyst in the reactions, the Bi promoted mesoporous sample is the one showing at the same time the lowest surface area and the top photocatalytic performances.

Possibly this effect is to be connected to a lower recombination between electrons and holes due to a better separation between the two oxide phase domains. Photocurrent measurements are performed to further analyze this latter point. The probability in charge recombination was studied by photocurrent kinetics curves (inset of Fig. 8). The photocurrent decay indicates that electron-hole recombination processes are occurring. The photocurrent transient can be evaluated by considering the kinetic equation

$$\ln(D) = -t/\tau,$$

where $D = \frac{I(t) - I(st)}{I(in) - I(st)}$; τ is the transient time constant, which indicates the time of recombination

process; $I(t)$ is the current at time t ; $I(st)$ is the stationary current, and $I(in)$ is the current at $t = 0$.

Fig. 8 shows the comparison between $\ln(D)$ vs. time plot of Tsg and Tsg_Bi. The linear behavior observed for all the tested materials indicates that the decay mechanism is due only to surface recombination leading to a first-order kinetics in electrons surface concentration (Dholam et al. 2010; Tafalla et al. 1990); the different slopes supports a longer recombination time for the Bi promoted sample.

The transient time constant (τ), reported in Tab. 4 for the adopted samples are larger (15-20 s) than those observed by compact single-crystal or polycrystalline electrodes (few milliseconds, Dholam et al. 2010; O'Regan et al. 1990), but comparable with what reported by other authors in the case of both pure and doped nano-titania photocatalysts (Dholam et al. 2010). The τ values of the Bi-promoted samples (Tab. 3) are slightly better than those of the undoped ones, especially in the case of T_P104_1.2_Bi, in which the controlled mesopores are filled by bismuth oxide, leading to a better charge separation efficiency. Thus the addition of bismuth in a tailored mesoporous titania structure enhances the photocatalytic efficiency, as demonstrated by the mineralization tests. The optimal amount of Bi presently selected is the result of a balance between the enhancement of the charge separation efficiency and the reduction of the sample surface area. Increasing Bi/Ti ratio

above 2% results in a marked decrease in surface area (25% passing from 1% to 3%) that reduces the photocatalytic activity of the sample in agreement with literature results (Shamaila et al. 2010).

Conclusions

Block copolymers of the Pluronic family (P104, P123, F127) successfully induce mesoporosity in nano-TiO₂ samples. The procedure adopted for the surfactant removal, a combination of thermal treatments and UV irradiation steps, permits to avoid pore collapse as well as to obtain a good degree of crystallinity. The sequence of the mesopore sizes in the samples obtained by different Pluronic copolymers follows the sequence of the relative micelle size in water, measured in the same conditions of the oxide formation. Copolymers characterized by relatively low molecular weights (around 6000) and by intermediate HLB values (P104 and P123) induce larger surface areas and pore volumes with respect to another Pluronic molecule (F127), characterized by a twice as large MW and a more relevant hydrophilic component. P104, used in the present case for the first time in the synthesis of nano-TiO₂, appears to be the best promoter of mesoporosity, producing increases in both surface area and total pore volume starting from very low copolymer initial amounts.

Bi-promoted oxides were prepared by using as scaffolds pre-calcined TiO₂ samples. For the sake of comparison both the mesoporous and a “classic” sol-gel TiO₂ sample were employed. Powder X-ray diffraction lines do not confirm the presence of bulk Bi₂O₃ but HRTEM micrographs show effects possibly linked to a highly monodispersed bismuth oxide into the titania matrix. Furthermore, EDX analyses support the presence of Bi in the final samples in the same Bi/Ti ratio as the one adopted in the synthesis. Definite differences appear when comparing the sol-gel and mesoporous Bi-promoted samples. In the latter case a decrease in surface area and pore volume suggests the desired formation of Bi₂O₃ in the voids created by the mesoporosity. Both FTIR and zeta potential measurements show that the Bi/Ti interactions are mainly localized at the surface of the titania particles.

The presence of Bi promotes the photocatalytic activity of the final samples. The top performing photocatalyst is the Bi₂O₃-mesoporous TiO₂, both for the degradation of methylene blue onto the dry oxide and for the degradation of formic acid in the liquid phase. Photocurrent measurements show that the recombination time between electrons and holes is the largest in the case of this sample. Apparently the prevailing location of Bi₂O₃ in the mesopores allows the formation of separate Bi₂O₃/TiO₂ domains; this occurrence could promote an increase in charge separation efficiency due to interfacial electron transfer via the junction between TiO₂ and Bi₂O₃ nanocrystals. Bi₂O₃ may further, due to its band structure, act as an electron acceptor which aids the effective separation between electron and hole pairs.

Acknowledgements

The authors wish to thank Prof. G. D'Alfonso and Dr. M. Panigati of the University of Milan for the determination of the micelle size distribution of the different block copolymers. This research has been supported by the University of Milan Research Funds (FIRST, PUR).

References

Alexandridis P, Holzwarth JF, Hatton TA (1994) Micellization of poly(ethylene oxide)-poly(propylene oxide)-poly(ethylene oxide) triblock copolymers in aqueous-solutions - thermodynamics of copolymer association. *Macromolecules* 27:2414-2425. DOI: 10.1021/ma00087a009

Alexandridis P, Hutton TA (1995) Poly(ethylene oxide)-poly(propylene oxide)-poly(ethylene oxide) block-copolymer surfactants in aqueous-solutions and at interfaces - thermodynamics, structure, dynamics, and modeling. *Colloids Surf A* 96:1-46. doi: 10.1016/0927-7757(94)03028-X

Alexandridis P, Nivaggioli T, Hatton, TA (1995) Temperature effects on structural-properties of pluronic P104 and F108 PEO-PPO-PEO block-copolymer solutions. *Langmuir* 11:2847-2847. doi: 10.1021/la00007a085

Angelome P, Andrini L, Calvo ME, Requejo FG, Bilmes SA, Soler-Illia GJAA (2007) Mesoporous anatase TiO₂ films: Use of TiK XANES for the quantification of the nanocrystalline character and substrate effects in the photocatalysis behavior. *J Phys Chem C* 111:10886-10893. doi: 10.1021/jp069020z

Ardizzone S, Bianchi CL, Cappelletti G, Naldoni A, Pirola C (2008) Photocatalytic degradation of toluene in the gas phase: Relationship between surface species and catalyst features. *Environ Sci Technol* 42: 6671-6676. doi: 10.1021/es8009327

Ardizzone S, Cappelletti G, Meroni D, Spadavecchia F (2011) Tailored TiO₂ layers for the photocatalytic ozonation of cumylphenol, a refractory pollutant exerting hormonal activity. *Chem Commun* 47:2640–2642. doi: 10.1039/c0cc05134a

Barrett EP, Joyner LG, Halenda PH (1951) The determination of pore volume and area distributions in porous substances. I. Computations from nitrogen isotherms. *J Am Chem Soc* 73:373–380 doi:10.1021/ja01145a126

Bates FS, Fredrickson GH (1999) Block copolymers - Designer soft materials. *Phys Today* 52: 32-38. doi: 10.1063/1.882522

Bian ZF, Zhu J, Wang SH, Cao Y, Qian XF, Li HX, (2008) Self-assembly of active Bi₂O₃/TiO₂ visible photocatalyst with ordered mesoporous structure and highly crystallized anatase. *J Phys Chem C* 112:6285-6262. doi: 10.1021/jp800324t

Bleta R, Alphonse P, Lorenzato L (2010) Nanoparticle Route for the Preparation in Aqueous Medium of Mesoporous TiO₂ with Controlled Porosity and Crystalline Framework. *J Phys Chem C* 114:2039-2048. doi: 10.1021/jp909646h

Boiadjieva T, Bianchi CL, Cappelletti G, Ardizzone S, Rondinini S, Vertova A (2004) The role of surface electrification on the growth and structural features of titania nanoparticles. *Phys Chem Chem Phys* 6:3535-3539. doi: 10.1039/b402370f

Bosc F, Ayral A, Albouy PA, Guizard C (2003) A simple route for low-temperature synthesis of mesoporous and nanocrystalline anatase thin films. *Chem Mater* 15:2463-2468. doi: 10.1021/cm031025a

Brinker CJ, Lu Y, Sellinger A, Fan H (1999) Evaporation-induced self-assembly: Nanostructures made easy. *Adv Mater* 11:579-585. doi: 10.1002/(SICI)1521-4095(199905)11:7<579::AID-ADMA579>3.0.CO;2-R

Carp O, Huisman CL, Reller A (2004) Photoinduced reactivity of titanium dioxide. *Progr Solid State Chem* 32:33-117. doi: 10.1016/j.progsolidstchem.2004.08.001

Choi SY, Mamak M, Coombs N, Chopra N, Ozin GA (2004) Thermally stable two-dimensional hexagonal mesoporous nanocrystalline anatase, meso-nc-TiO₂: Bulk and crack-free thin film morphologies. *Adv Fun Mater* 14:335-344. doi: 10.1002/adfm.200305039

Crepaldi EL, Soler Illia GJDAA, Grosso D, Castagnol F, Ribot F, Sanchez C (2003) Controlled formation of highly organized mesoporous titania thin films: From mesostructured hybrids to mesoporous nanoanatase TiO₂. *J Am Soc* 125:9770-9786. doi: doi: 10.1021/ja030070g

de Boer JH, Lippens BC, Linsen BG, Broekhof JCP, van den Heuvel A, Osinga TJ (1966) The curve of multimolecular N₂-adsorption. *J Colloid Interface Sci* 21:405–414 doi:10.1016/0095-8522(66)90006-7

Dholam R, Patel N, Santini A, Miotello A (2010) Efficient indium tin oxide/Cr-doped-TiO₂ multilayer thin films for H₂ production by photocatalytic water-splitting. *Int J Hydrogen Energy* 35:9581-9590. DOI: 10.1016/j.ijhydene.2010.06.097

Di Paola A, Marcì G, Palmisano L, Schiavello M, Uosaki K, Ikeda S, Ohtani B (2002) Preparation of polycrystalline TiO₂ photocatalysts impregnated with various transition metal ions: Characterization and photocatalytic activity for the degradation of 4-nitrophenol. *J Phys Chem B* 106:637-645. doi: 10.1021/jp013074l

Hagfeldt A, Lindström H, Södergren S, Linqvist S-E (1995) Photoelectrochemical studies of colloidal TiO₂ films - the effect of oxygen studied by photocurrent transients. *J Electroanal Chem* 381:39-46. doi: 10.1016/0022-0728(94)03622-A

Hecht E, Hoffmann H (1995) Kinetic and calorimetric investigations on micelle formation of block-copolymers of the poloxamer type. *Colloids Surf A* 96:181-197. DOI: 10.1016/0927-7757(94)03044-Z

Jing L, Wang J, Qu Y, Luan Y (2009) Effects of surface-modification with Bi₂O₃ on the thermal stability and photoinduced charge property of nanocrystalline anatase TiO₂ and its enhanced photocatalytic activity. *Appl Surf Sci* 256:657-663. doi: 10.1016/j.apsusc.2009.08.037

Kadam Y, Yerramilli U, Bahadur A, Bahadur P (2011) Micelles from PEO-PPO-PEO block copolymers as nanocontainers for solubilization of a poorly water soluble drug hydrochlorothiazide. *Colloids Surf B* 83:49-57. doi: 10.1016/j.colsurfb.2010.10.041

Kosmulski M (2009) Compilation of PZC and IEP of sparingly soluble metal oxides and hydroxides from literature. *Adv Colloid Interface Sci* 152:14-25. DOI: 10.1016/j.cis.2009.08.003

Lee ES, Oh YT, Youn YS, Nam M, Park B, Yun J, Kim JH, Song H-T, Oh KT (2011) Binary mixing of micelles using Pluronics for a nano-sized drug delivery system. *Colloids Surf B* 82:190-195. DOI: 10.1016/j.colsurfb.2010.08.033

Little LH (1966) *Infrared Spectra of Adsorbed Species*. Academic Press, London

Long R, English NJ (2009) Synergistic Effects of Bi/S Codoping on Visible Light-Activated Anatase TiO₂ Photocatalysts from First Principles. *J Phys Chem C* 113:8373–8377. doi: 10.1021/jp900589k

Magnacca G, Cerrato G, Morterra C, Signoretto M, Somma F, Pinna F (2003) Structural and surface characterization of pure and sulfated iron oxides. *Chem Mater* 15:675–687. DOI: 10.1021/cm021268n

Meroni D, Ardizzone S, Cappelletti G, Oliva C, Ceotto M, Poelman D, Poelman H (2011) Photocatalytic removal of ethanol and acetaldehyde by N-promoted TiO₂ films: The role of the different nitrogen sources. *Catal Today* 161:169-174. doi: 10.1016/j.cattod.2010.08.013

Milanesi F, Cappelletti G, Annunziata R, Bianchi CL, Meroni D, Ardizzone S (2010) Siloxane-TiO₂ Hybrid Nanocomposites. The Structure of the Hydrophobic Layer. *J Phys Chem C* 114:8287-8293. doi: 10.1021/jp1014669

Morterra C, Bolis V, Fiscaro E (1989) The hydrated layer and the adsorption of CO at the surface of TiO₂ (anatase). *Colloids Surf* 41:177–188. DOI: 10.1016/0166-6622(89)80051-4

O'Regan B, Moser J, Anderson M, Grätzel M (1990) Vectorial electron injection into transparent semiconductor membranes and electric-field effects on the dynamics of light-induced charge separation. *J Phys Chem* 94:8720-8726. DOI: 10.1021/j100387a017

Paoli E, Cappelletti G, Falcicola L (2010) Electrochemistry as a tool for nano-TiO₂ deposition and for photoremediation pollutant monitoring. *Electrochem Commun* 12:1013–1016. doi: 10.1016/j.elecom.2010.05.012

Perry CC, Sabir TS, Livingston WJ, Milligan JR, Chen Q, Maskiewicz V, Boskovic DS (2011) Fluorescence of commercial Pluronic F127 samples: Temperature-dependent micellization. *J Colloid Interface Sci* 354:662-669. doi: 10.1016/j.jcis.2010.10.028

Rouquerol J, Rouquerol F, Sing KSW (1999) Adsorption by powders and porous solids. Academic Press, London.

Saadoun L, Ayllon JA, Jimenez-Becerril J, Peral J, Domenech X (2000) Synthesis and photocatalytic activity of mesoporous anatase prepared from tetrabutylammonium-titania composites. *Mater Res Bull* 35:193-202. doi: 10.1016/S0025-5408(00)00204-X

Sanchez C, Boissière C, Grosso D, Laberty C, Nicole L (2008) Design, synthesis, and properties of inorganic and hybrid thin films having periodically organized nanoporosity. *Chem Mater* 20:682-737. doi: 10.1021/cm702100t

Shamaila S, Sajjad AKL, Chen F, Zhang J (2010) Bismuth-Doped Ordered Mesoporous TiO₂: Visible-Light Catalyst for Simultaneous Degradation of Phenol and Chromium. *Chem Eur J* 16:13795-13804. doi: 10.1002/chem.201001099

Spadavecchia F, Cappelletti G, Ardizzone S, Bianchi CL, Cappelli S, Oliva C, Scardi P, Leoni M, Fermo P (2010) Solar photoactivity of nano-N-TiO₂ from tertiary amine: role of defects and paramagnetic species. *Appl Catal B* 96:314–322. doi: 10.1016/j.apcatb.2010.02.027

Tadros T (2009) Polymeric surfactants in disperse systems. *Adv Colloid Interface Sci* 147-148:281-299. doi: 10.1016/j.cis.2008.10.005

Tafalla D, Salvador P, Benito RM (1990) Kinetic approach to the photocurrent transients in water photoelectrolysis at n-TiO₂ electrodes. Analysis of the photocurrent-time dependence. *J Electrochem Soc* 137:1810-1815. DOI: 10.1149/1.2086809

Taylor DJF, Thomas R.K., Penfold J (2007) Polymer/surfactant interactions at the air/water interface. *Adv Colloid Interface Sci* 132:69-110. doi: 10.1016/j.cis.2007.01.002

Tsui HW, Hsu YH, Wang JH, Chen LJ (2008) Novel Behavior of Heat of Micellization of Pluronics F68 and F88 in Aqueous Solutions. *Langmuir* 24:13858-13862. doi: 10.1021/la803272y

Tsui HW, Wang JH, Hsu YH, Chen LJ (2010) Study of heat of micellization and phase separation for Pluronic aqueous solutions by using a high sensitivity differential scanning calorimetry. *Colloid Polym Sci* 288:1687-1696. doi: 10.1007/s00396-010-2308-5

Wang JS, Li H, Li HY, Zou C (2010) Mesoporous $\text{TiO}_{(2-x)}\text{A}_{(y)}$ (A = N, S) as a visible-light-response photocatalyst. *Solid State Sci* 12:490-497. doi: 10.1016/j.solidstatesciences.2009.12.01

Wu Q.L., Rankin S.E (2011) Tuning the Mesopore Size of Titania Thin Films Using a Polymeric Swelling Agent. *J Phys Chem C* 115:11925-11933. doi: 10.1021/jp2021193

Xu H, Zhang L (2009) Controllable One-Pot Synthesis and Enhanced Photocatalytic Activity of Mixed-Phase TiO_2 Nanocrystals with Tunable Brookite/Rutile Ratios. *J Phys Chem C* 113:1785-1790. DOI: 10.1021/jp8089903

Zana R, Marques C, Johner A (2006) Dynamics of micelles of the triblock copolymers poly(ethylene oxide)-poly(propylene oxide)-poly(ethylene oxide) in aqueous solution. *Adv Colloid Interface Sci* 123-126:345-351. DOI: 10.1016/j.cis.2006.05.011

Fig. captions

Fig. 1 TG analyses under N₂ atmosphere of **a)** P104 copolymer and of **b)** un-calcined T_P104_1.2

Fig. 2 N₂ adsorption–desorption isotherms at subcritical temperatures of T and T_P104_1.2 samples calcined at 350°C. Inset: Relative pore volume distribution of the two samples

Fig. 3 a) Micelle size distribution by light scattering for the three types of Pluronic; **b)** pore size distribution of the calcined TiO₂ powders by nanoparticle route in the presence of the adopted copolymers

Fig. 4 % MB mineralization as a function of the surface area of the TiO₂ samples

Fig. 5 N₂ adsorption–desorption isotherms of bare (Tsg and T_P104_1.2) and Bi-promoted (Tsg_Bi and T_P104_1.2_Bi) titania samples.

Fig. 6 HRTEM micrographs of pure and Bi-promoted titania powders

Fig. 7 FTIR spectra of bare (Tsg and T_P104_1.2) and Bi-promoted (Tsg_Bi and T_P104_1.2_Bi) titania samples. Full lines: activation in vacuo at RT, bold lines: activation in vacuo at 250°C)

Fig. 8 lnD versus time plot of Tsg and Tsg_Bi. Inset: photocurrent kinetic curve relative to Tsg_Bi.

Sample	% Anatase	% Brookite	d_{101}^a (nm)	S_{BET} ($\text{m}^2 \text{g}^{-1}$)	V_{pores} (mL g^{-1})
<i>T</i>	73	27	8	92	0.119
<i>T_P104_0.5</i>	62	38	8	147	0.352
<i>T_P104_0.8</i>	61	39	7	175	0.395
<i>T_P104_1.0</i>	74	26	7	178	0.398
<i>T_P104_1.2</i>	61	39	6	187	0.425
<i>T_P104_1.4</i>	61	39	6	205	0.429
<i>T_P104_1.6</i>	70	30	6	197	0.515
<i>T_P123_1.2</i>	59	41	6	199	0.391
<i>T_F127_1.2</i>	59	41	6	172	0.262

Tab. 1. Structural (anatase and brookite content, anatase crystallite sizes d_{101}^a) and morphological (BET surface area S_{BET} , total pore volume V_{pores}) features of bare titania samples.

Pluronic copolymer	MM (g mol⁻¹)	EO_xPO_yEO_x	HLB	CMC × 10⁶ (mol L⁻¹)
<i>P104</i>	5900	x = 25, y = 56	8	4.0
<i>P123</i>	5750	x = 20, y = 70	7 – 12	3.1
<i>F127</i>	12600	x = 106, y = 70	18 – 23	18.0

Tab. 2. Properties of Pluronic P104, P123, and F127 copolymers: average molecular weight, average composition, hydrophilic-lipophilic balance (HLB) (Taylor et al. 2007), and critical micelle concentration (as determined by surface tension measurements, see Online resources Fig.2).

Sample	% Anatase	% Brookite	d_{101}^a (nm)	S_{BET} (mL g ⁻¹)	V_{pores} (ml g ⁻¹)	ζ potential (mV)
<i>Tsg</i>	70	30	7	150	0.338	-38
<i>T_Bi</i>	57	43	7	161	0.420	-10
<i>T_P104_1.2</i>	61	39	6	187	0.425	-35
<i>T_P104_1.2_Bi</i>	53	47	7	148	0.375	-23

Tab. 3. Morphological features (BET surface area S_{BET} , total pore volume V_{pores} , content of pores with sizes ranging from 6 to 20 nm) of bare TiO₂ and Bi-TiO₂ composites.

Sample	% MB mineralization	HCOOH $\frac{d\chi}{dt}$ (mS min ⁻¹)	τ (sec)
<i>Tsg</i>	46 ± 1	0.51 ± 0.02	15
<i>Tsg_Bi</i>	68 ± 1	0.58 ± 0.01	16
<i>T_P104_1.2</i>	54 ± 1	0.48 ± 0.02	15
<i>T_P104_1.2_Bi</i>	78 ± 1	0.76 ± 0.01	19

Tab. 4. Photocatalytic performances of bare and Bi-promoted TiO₂ samples for MB and formic acid mineralization. Transient time constant (τ) extrapolated from the photocurrent measurements.