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INORGANIC MARKERS PROFILING IN WILD TYPE AND GENETICALLY MODIFIED PLANTS SUBJECTED TO ABIOTIC STRESSES

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Abstract

Nicotiana is a well characterized genus of *Solanaceae*. It is widely studied in literature in different research fields, such as biology, environmental chemistry, etc.. In this study *Nicotiana langsdorffii* wilde type plants and transgenic for the rat glucocorticoid receptor gene (*GR*) and *Agrobacterium rhizogenes rolC* gene were used to value the effect of exposure to chemical compounds (Cr) and to harsh physical conditions (dryness, heat stress) on the inorganic compositions in plants. Four markers were selected, namely Na, K, Cl⁻ and NO₃⁻. Different procedures of extraction were tested: water permitted to obtain the better conditions for the determination of four selected markers. The analysis was carried out by emission atomic spectroscopy for the determination of the cations; the anions were determined both by ion chromatography and potentiometry: the results obtained by these two techniques and their potentiality were compared. The total cations content (Al, Ca, Cr, Fe, K, Mg, Mn, Na, P, Si and Zn) was also determined. The effect of the genetic modifications on the capacity of the plant to contrast the stress stimuli was valued: the introduction of GR gene seems to reduce in grater manner the effect caused by the stresses and the adsorption of the inorganic pollutants. The modification with *rolC* increases the absorption capacity of the plant forward Cr and in the same time reduces its release: this behaviour can be adopted in phytoremediation applications.

Keywords: Nicotiana transgenic plants, rat glucocorticoid receptor, rolC, inorganic stress markers, chemometrics

1. Introduction

In the case of plants, we can define “stress” as “a strength that tends to show an influence on the natural mechanism of its functioning” [1]. For example, extreme environmental phenomena such as droughts cause conditions of stress for plants with a significant impact on their physiology, morphology and development [2].

Plants exposed to biotic and abiotic stresses respond with various biochemical reactions to protect themselves against the causal agent [1]: the action of stresses can induce, within the vegetal organisms, the production of compounds able to contrast the effects of the attack [3,4].

Abiotic stresses can be classified in i) hydric (drought or flooding), ii) thermal (heating or chilling/freezing); iii) saline (salinity); iv) mineral (deficiency or excess); v) mechanic (brushing or restriction). Another classification of abiotic stresses can be: i) physical, such as lack of water or presence of extremely high or low temperatures, or ii) chemical, in case of exposure to toxic species.

In particular, in the case of water deficit, which can be due to drought, cold or osmotic shock, an accumulation of various molecular compounds, such as proline, glutamate, glycine-betaine, carnitine, mannitol, sorbitol, fructans, threosium, saccharosium and oligosaccharides within the plant cells was observed: it is believed that this effect is due to an increase of osmotic pressure inside the cells and to an higher retention of water. Furthermore, with the aim of maximizing water retention, an increase in the production of abscissic acid takes place, with a consequent closure of stomes [5].

Thermal stress at high temperature is caused by the exposure of plants at a temperature higher than that of maximum growth for long or short periods. Most plants are not able to survive at temperatures higher than 45°C, with a few exceptions as some Mediterranean and tropical species (48-55°C) and some woody subtropical plants (50-60°C). Independently from the capacity of surviving at high temperature, the stress condition starts above 30°C. [6].

It is well known that the presence of high concentration of heavy metals can cause some of the most severe damages to plants. Most metal ions are essential in the maintenance and evolution of all living systems, and also mediate all stages of dissemination of genetic information carried in the genetic code. On the other hand, essential metals, when present in excess can produce errors in the genetic information system and/or cause serious damages to plants or even their death. In particular, these detrimental effects are exerted by metals which have no biological roles, such as Cd and Pb, when present above certain limits of concentration. The amount of a metal absorbed by plants from soil depends on its concentrations and speciation in the soil solutions, its translocation from the bulk

soil to the root surface, then into the root and finally into the shoot and leaves. The same uptake processes take place both for essential micronutrient cations and non-essential metal ions.

Excessive concentrations of metals result in phytotoxicity through: i) changes in the permeability of the cell membrane; ii) reactions of sulphhydryl (-SH) groups with cations; iii) affinity for reacting with phosphate groups and active groups of ADP or ATP; and iv) replacement of essential ions [4,7]. Several plant species developed tolerant variants that can survive on metal-contaminated environment through adaptation [8].

The knowledge of the possible response to different stresses can be helpful i) to understand the physiology and the biochemistry of plants, ii) to identify methodologies and techniques suitable to reduce the stress effect on their growth and production and iii) to select stress-resistant cultivars.

The effect of both biotic and abiotic stresses can be assessed by monitoring the variation of the in-cell levels of different phytohormones produced by the plants for their defense. Among these, phytohormones are the most extensively studied in literature, because they exert remarkable effects on the processes of growth and development of plants such as germination, foliar expansion and flower development [9-12]. Also some inorganic species represent important indicators of stress in vegetation. In particular the content of alkali metals, namely sodium and potassium, and inorganic anions, such as chlorides and nitrates, changes as a consequence of stress. For example, K^+ , Na^+ and Cl^- are nutrients that remain as ions in the plant tissues. K^+ is required as cofactor from more than forty enzymes and has an essential role in the maintenance of cellular turgidity and electroneutrality. Cl^- is required in the photosynthetic reactions for the generation of O_2 . Na^+ is involved in the generation of phosphoenolpyruvate and it can substitute K in some functions. N is a constituent nutrient of the organic vegetal components: amino acids, amides, proteins, nucleic acids, nucleotides, coenzymes and esamines. Plant available inorganic forms of N include nitrate (NO_3^-); and nitrite (NO_2^-); as well as ammonium (NH_4^+) [13]. The variation in their concentration as a consequence of abiotic stress is due to the fact that such ions have an important role in the transport of matter within plants and in osmoregulation. While a lot of papers are available in literature about phytohormones, very few papers pertain to these species [14-16].

In this paper, we have compared the changes of the inorganic markers pattern in a model plant system, namely *Nicotiana langsdorffii* wild type plants and transgenic for the rat glucocorticoid receptor gene (*GR*) or transgenic for the *Agrobacterium rhizogenes rolC* gene, after the Cr(VI) exposure, scarcity of water and high temperature. This system has been used in previous studies for the analysis of different stress-induced effects. Data on changes in hormonal endogenous content are reported in Fuoco et al., 2013 [17], while changes in sugars and phenolic compounds or heavy

metals accumulation are showed by Del Bubba et al. 2013 [18] and Ancillotti et al 2015 [19]. A deep analysis of metabolomic patterns [20] and changes in terpenoids and ionic profiles [21,22] were also already reported. In this paper we obtained new data both on the effects of different stresses and of the presence of a genetic modification on the plant response. In addition to the selected markers, namely Cl^- , NO_3^- , K^+ and Na^+ , the concentrations of Cr and of other elements (Al, Ca, Fe, Mn, Si, P and Zn) were determined. The data were treated with chemometric techniques to better value the effect of the different stresses on the considered plants. Moreover, this work provides a complete optimised protocol for plant treatment and analysis for the determination of these type of analytes.

2. Materials and methods

2.1 Plant material and sample pretreatment

Nicotiana langsdorffii wild-type and transgenic for rat gluco-corticoid receptor gene (GR) or *Agrobacterium rhizogenes rolC* gene were analyzed.

Plant modification, incubation and growth were carried out at the former Department of Evolutionary Biology “Leo Pardi” (now Department of Biology), University of Florence, by Prof. M. Buiatti and Dr. P. Bogani. The modified plants were obtained and characterized following the procedure described elsewhere [19,23,24].

The examined plants can be considered as a homogeneous material with the exception of the presence of a gene in the modified plants, that changes the hormonal equilibrium. The plant specimens were grown in controlled and reproducible conditions. Cultures were incubated in a growth chamber at $24\pm 1^\circ\text{C}$ with a photoperiod of 16 h of light (1500 lux) and 80% relative humidity. After germination, seedlings were grown on LS medium for one month.

2.2 Apparatus and reagents

Sample dissolution for the determination of total metal concentrations was performed in polytetrafluoroethylene PTFE bombs, with a Milestone MLS-1200 Mega (Milestone, Sorisole, Italy) microwave laboratory unit.

The cations were determined with a Perkin Elmer Optima 7000 (Perkin Elmer, Norwalk, Connecticut, USA) inductively coupled plasma-optical emission spectrometer (ICP-OES). The

anions were determined by a DX-500 (Dionex, Sunnyvale, CA, USA) ion chromatograph (IC) equipped with a conductimeter detector (ED40, Dionex), using an Ion Pack AS9-HC anion exchange column. The results were compared with those obtained analyzing the same extracts by an Metrohm potentiometer equipped with combined Ion Selective Electrodes, ISE (Thermo Fisher Scientific, the Netherlands).

Analytical grade reagents were used throughout. Standard metal solutions were prepared from concentrated Merck Titrisol stock solutions (Merck, Darmstadt, Germany).

High purity water (HPW) obtained from a Milli-Q apparatus (Millipore, Bedford, USA) was used throughout for the preparation of sample and standard solutions.

2.3 Procedures

2.3.1 Stress exposure

Chemical stress (CS). The concentration of Cr for the induction of metal stress in *Nicotiana* plants was selected on the basis of preliminary experiments aimed at testing the lethal-dosage 50 (LD50) on survival and callus formation capacity of leaf tissues of wild type *N. langsdorffii* plants grown *in vitro*. The standard working solutions were prepared from potassium dichromate ($K_2Cr_2O_7$) concentrated standard solutions (Merck Titrisol).

Working solutions were filter-sterilised by 0.22 μm Millipore membrane filters and a suitable volume was added to the LS (Linsmaier e Skoog) culture medium to have a concentration of 50 ppm. The complete procedure is described in the paper by Fuoco et al. (2013)[17].

Drought stress (DS). Polyethylene glycol (PEG) was added to dehydrate the LS medium (described in Bogani et al. 2015 [24]). PEG is widely used to modify the osmotic potential of nutrient solution cultures and to induce plant water deficit in a relatively controlled manner, appropriate to experimental protocols. It was assumed that PEG of large molecular weight does not penetrate the plant and thus is an ideal osmoticum for use in hydroponics root medium [25]. The concentration of PEG for the induction of drought stress was evaluated by a dose/effect curve experiment conducted on leaf discs of *N. langsdorffii* wild type and transgenic. In particular, to each Petri dish containing approximately 20-25 ml of solidified LS + BAP (6-benzylaminopurine) 1mg/l, NAA (1-naphthaleneacetic acid) 0.1 mg/l and 20 ml of 0% or 5% or 10% or 20% PEG 6000 were added 60 leaf discs per genotype. After 24 hours the PEG solution was removed in accordance with the method described in van der Weele et al., 2000 [26], and sample growth capacity (callus induction)

and regeneration (induction of shoots) was evaluated after. The results of the experiments indicate that the ability of growth of non-transformed *N. langsdorffii* is strongly inhibited by the concentration of PEG 20% and that the regeneration capacity is drastically reduced already at a PEG concentration of 5%. Surprisingly, the explants of transgenic plants, and particularly those obtained from the transgenic plants for the GR receptor are more resistant to treatment with PEG 20%. On the basis of these results, a PEG concentration of 20% was used for the assessment of the effects of response to water stress. For this purpose, 50 plants of *N. langsdorffii* wild type and transgenic for GR or *rolC* gene were maintained on LS culture medium containing 20% PEG for two weeks before being analyzed along with the corresponding not exposed controls.

Heat stress (HS). Since most plants are not able to survive at temperatures higher than 45°C and the stress condition starts above 30°C [6], the heat stress was induced on 50 plants of *N. langsdorffii* wild type and transgenic for GR or *rolC* gene in LS culture medium by maintaining the samples at 50°C for 2 h in an incubator. At the end of the period the samples were immediately frozen with liquid nitrogen and analyzed along with the corresponding not exposed controls.

2.3.2 Sample pretreatment

The samples were frozen with liquid nitrogen if not already frozen and immediately grinded in a mortar in order to obtain a powdered sample. Then, they were freeze-dried and a recovery of 5 ± 0.5 % was obtained for all samples. The results reported are referred to the freeze-dried samples.

2.3.3 Extraction of anions and cations

Three procedures were selected among the different ones available in literature regarding the extraction of anions from vegetables. These procedures were preliminarily tested on lettuce plants purchased from a local market to value the extraction efficiency for the considered inorganic markers. The tested procedures were:

Procedure A: the sample was cut into small pieces of about 1 cm. A 1.5 g aliquot was weighed, suspended in 500 ml of high purity water (HPW), and homogenized for 1 minute [27-29].

Procedure B: a 10 g aliquot of sample was added with 50 ml of 200 mM sodium borate solution heated at 80°C. The sample was brought to 100 ml with sodium borate solution, heated again at 80°C for 15 min, and finally cooled [30].

Procedure C: a 0.5 g aliquot of sample was shaken with 25 ml of 2 % (v/v) acetic acid for 30 minutes on an oscillating rolling-bed shaker [31].

In all cases, the suspensions obtained were centrifuged for 5 min, filtered through a 0.45 μm cellulose acetate filter and conveniently diluted for anion and cation determinations.

Taking into account the preliminary results on the lettuce plant, HPW was chosen as extractant for the analysis of *Nicotiana* samples (see section 3.3) following the conditions specified in procedure C. The solutions obtained were diluted to 1:20 and 1:200 with water for the determination of cations and anions respectively.

All experiments were performed in triplicate.

2.3.4 Total cation content

The total concentrations of cations were determined after acid digestion in microwave oven, treating aliquots of 0.2 g of each sample with 6 ml of a mixture of HNO_3 : H_2O_2 = 1:1. After the digestion, the solutions were filtered through Whatman 5 filters and diluted with HPW to 50 ml. The final solutions were analysed by ICP-OES.

Since the plant dissolution requires the use of nitric acid that would obviously add a great amount of NO_3^- ions to the sample solutions, the total concentration of anions was not determined.

2.3.5 Sample analysis

Water extracts were analysed using IC and potentiometry for the determination of anions. Analyte concentrations were quantified by the standard addition technique.

The eluent used during IC determination consisted of a mixture of 10 mM K_2CO_3 / 4 mM NaHCO_3 (90%) and HPW (10%) and was run through the system at a flow rate of 1 ml/min.

Potentiometric determinations of anions were conducted in the presence of total ionic strength adjustment buffer solutions, 0.06 M KNO_3 and 0.1 M $(\text{NH}_4)_2\text{SO}_4$ for Cl^- and NO_3^- determination respectively. To minimize the effect of the temperature, the solutions were stored at 25 $^\circ\text{C}$ until the measurement; the signal rapidly reached a stable value (10 s), and the electrode was carefully washed with HPW after each measurement.

The water extracts and the digested solution were analysed by ICP-OES for the determination of cations.

Three replicates were performed and blanks were simultaneously run. Standard solutions for instrument calibration were prepared in the corresponding blank, according to the matrix-matching technique.

The following wavelengths were selected for the determination of Na, K, Cr and other elements: $\lambda_{Al} = 396,153$, $\lambda_{Ca} = 317,933$, $\lambda_{Cr} = 267.716$, $\lambda_{Fe} = 259,939$, $\lambda_{K} = 769.896$, $\lambda_{Mg} = 285,213$, $\lambda_{Mn} = 257.61$, $\lambda_{Na} = 589.592$, $\lambda_{P} = 213,617$, $\lambda_{Si} = 251,611$, $\lambda_{Zn} = 213,857$ nm.

2.3.5 Chemometric data processing

A chemometric analysis of the experimental results was performed by principal component analysis (PCA), with the aid of an XLStat 7 software package, used as a Microsoft Excel plug-in.

Unscrambler X 10:2 was also employed for data standardization, obtained by mean-centering (for each variable) and dividing by the corresponding standard deviation, and for substituting values below LoQs with estimated values. Analytes with most values below the LoQ were not included. For the principles of the technique, the reader is referred to existing textbooks and papers on chemometrics [32,33].

3. Results and discussion

3.1. Plants appearance

The transformation of *N. langsdorffii* plants with both *GR* gene and *rolC* gene induced significant differences in both morphology and growth [24,34]. The exposure to the water stress due to the growth of plants for 15 days on LS medium containing 20% PEG, induced a clear wilting effect only on untransformed plants while *GR* and *rolC* transgenic plants did not show any stress symptom. Moreover, the wild type plants showed a significant decrease in both root and shoot biomass (dry weight data), only a decrease of root biomass in *GR* plants and any difference of biomass data in *rolC* plants. The exposure of plants to 50°C for two hours did not affect the plant appearance in the three types of plants. However a different re-growth capability was observed in transgenic plants, compared to the untransformed ones, when the plants were grown for three weeks at 25°C (the optimal growth temperature), after the heat treatment. These data suggested a higher tolerance to stress condition as due to the insertion of transgenes [19].

3.2 Total cation contents

Nicotiana langdorffii was not grown in soil, so the only sources of metals and of inorganic species is the growth medium and, at a much lesser extent, atmospheric deposition or impurities present in the water used for irrigation. A preliminary semiquantitative analysis of the plant extract showed the presence of quantifiable amounts of Al, Ca, Fe, Mg, Mn, P, Si and Zn: we determined the concentrations of these elements, in addition to those of the markers (Na and K) and of the added pollutant (Cr), to obtain further information on the changes of chemical composition of the plants after the applied treatments.

Table 1 reports the contents (expressed in g/kg) of the cations in the digested solutions. The most abundant element is K ($66\div 221$ g/kg), followed by the other macronutrients, namely P ($4.65\div 14.9$ g/kg), Ca ($4.38\div 7.39$ g/kg) and Mg ($1.10\div 2.48$ g/kg); as to micronutrients, Na ranges between 0.48 and 0.78 g/kg, whereas Al, Fe, Mn, Si and Zn contents are always lower than 0.4 g/kg.

We observed remarkable differences among the unstressed plants; in particular, WT samples contained the highest concentration values for K, Mg, Mn, P and Si, while *rolC*-modified specimens generally presented the lowest amounts of these cations. WT and GR contain 0.97 and 1.15 g/kg of Na respectively, while the concentration of this element in *rolC* is about half of this value. These results show how the genetic modifications are able to change the natural inorganic content in *Nicotiana langdorffii*. We can observe that Ca concentration does not present a particular trend, as confirmation of the fact that this element could not be considered as a good marker for this type of studies; in fact, Taiz et al. [6] underline that after any type of modification or variation it moves to different parts of the plant, but its total amount remains quite constant.

All the considered stresses cause a decreasing in K concentration both in the wild and the modified plants. In particular the heat stress has the strongest effect, causing a 65 ± 3 % loss of this element in all the plants. In dryness condition the plants naturally tend to increase the adsorption of K and other elements from the soil to maintain the osmoregulation [6]. It is a strategic response of the plant to survive, since storing solutes in the cells helps maintaining a positive turgidity pressure that is required to the cell expansion, even in the presence of a low hydric potential. The opposite trend in K concentration found in our study is probably due to the fact that the *Nicotiana* specimens were grown *in vitro*.

For Na we can observe as the presence of stress causes a decrease in its concentration in WT plants, while heat stress causes an increasing of Na content in *GR* and *rolC*-modified plants; in the case of *rolC*, also Cr-stress causes an increasing of Na concentration. Plants have evolved mechanisms to regulate Na accumulation and to select against it in favor of other nutrients as K^+ and NO_3^- . For example, in most plants Na^+ and Cl^- are effectively excluded by roots while water is taken up from highly salinity soils [34]. In the presence of physical stresses that can cause a metabolic disorder the stomes would not be able to expel it. For this reason, we expected higher levels of Na in the presence of stressful conditions as observed in the cases underlined for the modified plants.

The chemical stress induced by the presence of a pollutant, as Cr, causes phenomena of antagonism among the cations, giving rise to a decreasing of the concentration of the elements naturally present in the plants: this behavior was observed both for K and Na, with the only exception of Na in *rolC*. As expected, the concentration of Cr dramatically increased after exposure of the plants to this element. The genetic modification with GR permitted to reduce the plant uptake towards this pollutant (-21 %), in comparison with WT samples. This result confirms the observation reported by Fuoco et al.[33]: they found that chromium did not exert any significant leaf biomass variation in GR-modified plants, denoting the resistance of this transgenic form to high Cr concentration; they associated these effects with the minor metal uptake compared to WT plants. On the contrary, the modification with *rolC* caused a greater absorption of Cr by *Nicotiana* (+42 %) in comparison with WT according to results obtained by Del Bubba et al. (2013)[18].

Considering the total cation content, all the considered stresses cause its reduction in comparison with the unstressed plant. In particular the total cation content in the plants is more than halved in presence of heat stress.

3.3 Extractable contents

Three different extraction procedures were initially tested. Owing to the time required to obtain *Nicotiana* plants and the difficulties to produce a sufficient amount of plants, these preliminary experiments were conducted on lettuce bought from a local market. Table 2 shows the results obtained in these experiments. Following the original procedures, the C one appeared as the most effective. Unfortunately, the presence of acetic acid gave some problems of reproducibility during the analysis by IC. Obviously, the use of sodium borate does not permit the determination of Na. Then, the efficiency of the different extractants was compared using the same experimental conditions (amount of sample, volume of the extractant solution, contact time), as those reported in

procedure C. Since the extraction yield did not significantly increase using the other two reagents in comparison to water, we chose HPW as extractant.

Water is largely used in literature for the extraction of anions from plants, also because it is the most suitable matrix for chromatographic techniques [27-29,35].

Before treating the samples of *Nicotiana*, we also valued the effect of the contact-time between the plant and water on the efficiency of the extraction. During the stirring, every 20 minutes, an aliquot of the suspension was taken and analysed. The concentrations of all the markers in the extract increased until 120 minutes, then remained constant: so we chose 120 minutes as the best extraction time. We underline that the aim of this study is not to extract the maximum amount of elements chosen as markers, but to compare the behavior of the markers in the considered plants.

3.3.1 Extractable cation contents

The concentrations (g/kg) of Ca, Cr, K, Mg, Mn, Na, P and Zn found in the water extracts is reported in Table 3. Since the total element content greatly varies among the considered plants, for a better interpretation of the data we also report the percentages of extraction with respect to the total content.

Generally, the percentages of extraction for the same analyte from different plants and for different analytes from the same plant are quite variable (from 7 to 21 % for Ca, from 23.6 to 63.7 % for Mg, from 15.6 to 100 % for Mn, from 28.9 to 100 for Na, from 27.1 to 100 for P and from 19.8 to 66.5 % for Zn).

Al, Fe and Si concentrations were very close to the instrumental detection limit in all samples.

Comparing the unstressed plants, K, Na, Al, Ca Fe, Zn are released in higher extent from WT samples. A particular behavior is observed for Mn and P, quantitatively extracted from *rolC*.

Even if the total and the extracted amounts of K are very variable, the percentage of extraction of K is comprised in the range 12-19 for all the samples, with the exception of the stressed *rolC* plants, in which this element is little released. Heat stress caused a decrease in the release of this metal in all the plant types, in particular for the GR-modified one. We can observe that, while the presence of any type of stress causes the halving of the Na content in WT and GR, in *rolC*-modified plant the concentration of Na in unstressed specimens is similar and slightly lower than that found in the stressed plants.

Regarding the behavior of Cr, while *rolC* presented the highest total concentration of this pollutant, it shows the lowest Cr content in the water extracts, demonstrating that this genetic modification causes a higher and stronger adsorption of this metal by the plant. This result can be useful in soil remediation studies to obtain more heavy metal-resistant plant, able to adsorb greater amount of pollutants from the soil and not easily releasing them into the environment. It is well known that *rolC* is capable of inducing root development in different plants [36-38]; this effect could be used to improve the capability of selected vegetals to immobilize pollutants possibly present in contaminated soils. Also Nair et al. [39] cloned a secondary Ni–Co transporter (NiCoT, TC 2.A.52) from *Rhodospseudomonas palustris* vector into plant expression vector and introduced it into *Nicotiana tabacum*, demonstrating the potential of this transgenic plant to remediate cobalt from contaminated soils.

As to the other elements, no particular trends are observed. Only for WT, we can see that all applied stresses cause a decrease in the percentage of release, with the only exceptions of Mg and P in WT DS.

3.3.2 Anion contents

The content of NO_3^- and Cl^- in the water extracts was determined using IC and potentiometry. We chose to use IC results for data interpretation, since this technique is widely accepted as a suitable method for the determination of anions in environmental samples and is less prone to interferences. The results obtained with the two analytical techniques are compared in section 3.3.3.

Table 4 reports the concentrations found in the water extracts and Figures 1a and b show the trends of Cl^- and NO_3^- respectively, obtained with the two adopted techniques.

Among the unstressed plants no significant differences were present in the concentration of anions in WT and modified specimens.

Regarding the effect caused by the considered stresses, a common trend was observed: the concentrations of the anions decrease in presence of any type of stress. Chemical stress causes a very similar (-43÷-48%) reduction in the extractable anions in all the plants, with the exception of Cl^- extracted from GR-CS plant, which is extracted at a lower extent (-23%).

Heat stress causes a substantial decrease (> 50 %) of the extractable Cl^- from all the plants, while we observe low effect on NO_3^- . Water deficit gives rise to the most relevant reduction (> 80%) of

the extractable anions in the wild type-plants; both the genetic modifications, in particular that with GR, reduce this effect.

Usually, Na and Cl present a similar behavior, that is their concentration increase and decrease at the same time, both in unstressed and stressed conditions. For example, because NaCl is the most soluble and widespread salt, plants have evolved mechanisms to regulate its accumulation and to select against it in favor of other nutrients commonly present in the soil [34]. For this reason, we compared the results obtained for these two analytes (Figure B1: we obtained a good correlation between their concentration ($r^2 = 0.8661$), (after removing WT-DS, WT-HS and *rolC* from the data set).

3.3.3. IC vs ISE

The results obtained by IC and potentiometry were compared. The two techniques revealed the same trend in the concentrations of the considered anions with some differences in the quantification. The concentrations of Cl^- found with ISE were generally higher than those determined by IC, with the only exception of WT extract. The mean concentrations of Cl^- found in the unstressed plants (WT, GR and *rolC*) using IC and ISE were not significantly different for $p < 0.05$.

As to NO_3^- , the concentration observed with ISE is overestimated for the WT plants in comparison with IC, while for the genetically modified plants the concentration determined by ISE was always lower. For this anion the concentration obtained were not significantly different for $p < 0.05$ for all the samples, with the exception of GR-CS and *rolC*-CS.

We can make some considerations on the matrix effect observed using the two techniques. Table 5 reports the equations of the calibration curves obtained for each determination in HPW and in sample matrices using the two techniques.

The adopted conditions permitted us to obtain well defined chromatograms, with the peaks of interest at $t = 6.35$ min and $t = 11.05$ min for Cl^- and NO_3^- respectively. The repeatability obtained in three replicates of the sample extracts was good: the relative standard deviation was always lower than 2.5 %. To better value the effect of the sample matrix we sequentially added aliquots (three aliquots of 0.5 ml) of each plant extract to 100 ml HPW standard solutions (containing 4 mg/l of Cl^- and 1 mg/l of NO_3^-): a linear response was obtained. We can observe similar slope values for all the considered sample solutions both determining Cl^- and NO_3^- . The slopes of the calibration equation

obtained analyzing standard solutions prepared in water were higher than that obtained in the sample matrix in the case of Cl^- and lower in the case of NO_3^- .

Potentiometry is a cheaper and simpler method than IC, but the potentiometric response can be subject to interferences from other species possibly present in solution [40-42], which can bind to the ISE membrane, leading to an overestimation of the analyte concentration, or bind to the analyte in solution, hindering its interaction with the membrane and thus causing a negative interference. The theoretical slope of ISEs is $59,16$ working at room temperature, but the range 54 - 60 is considered satisfactory (user guide, Thermo scientific). As we can see from the table, performing the calibration with standard solutions prepared in HPW, the ISEs response is as expected. On the contrary the presence of the sample matrix drastically reduced the slope values: in the case of Cl^- it was -36.54 for all the types of plants, while in the case of NO_3^- it ranged from -32 to -23 decreasing (in absolute value) in the order $\text{WT} > \text{GR} > \text{rolC}$. This can be caused by the fact that the extracts contained high concentration of other ions in addition to the analytes: salts present in sample solutions may be extracted into the electrode membrane, causing deviation from the theoretical response. The slope variation follows the same order as the total ion content present in the water extract ($\text{WT} > \text{GR} > \text{rolC}$).

We valued the effect of the presence of other anions in the extracts, by analyzing synthetic solutions containing 4 mg/l of Cl^- and NO_3^- in the presence of 10 mg/l of potential interfering agents (Br^- , I^- , CH_3COO^- , NO_3^- or Cl^-). The results are reported in Table 6.

The presence of Br^- seems to have an acceptable effect on the determination of the considered analytes; also the presence of Cl^- or NO_3^- does not cause a variation in the measurement of the other marker. The presence of acetate ions has opposite effects on the determination of Cl^- and NO_3^- , and causes a strong positive interference on the determination of the latter. I^- gives rise to an overestimation of the Cl^- content and does not permit the quantification of NO_3^- , because its presence does not permit to reveal any variation in the instrumental potentiometric response after the additions of standard solutions.

In any case, even if repeatability with ISEs is limited by factors such as temperature fluctuations and noise, electrode measurements repeatable to $\pm 1,5 \%$ were obtained, which is satisfactory.

3.4 Chemometric treatment

The chemometric treatment of the experimental data was carried out through the application of a well-known statistical multivariate analysis technique: Principal Component Analysis (PCA).

The chemometric study was carried out i) on the total cation content in the digested solution (Figure 2a); ii) on the cation content in the water extracts (Figure 2b); iii) on the whole data set (all the cations and anions) found for the water extract solutions (Figure 2c); iv) considering only the concentrations of the markers in the water extract solutions (Figure 2d). We considered the concentration values and not the percentages of extraction, since it was not possible to determine the total content of the anions: in this manner we can compare the results obtained considering only the markers, only the cations and the whole data set. We took into account the first two PCs. Pearson's correlation coefficients for all data treatments are reported in Table A.1-A.4 (Appendix A).

As we can observe in the combined plot of scores and loadings in Figure 2a resulting from the treatment of data about the total cation content, the samples are grouped in different areas of the graph on the basis of the induced stress. The scores of unstressed and DS plants are in the same area of the biplot because these are generally characterized by the highest amount of K, Mn and Zn; the only exception is represented by GR-DS that contains a higher level of Na. HS-plants are grouped together in the opposite plot region in comparison with the direction of the loadings, since they generally contain the lowest concentrations of cations. Also the chemically stressed plants are grouped together and seem to be characterized by higher concentration of Al, Mg and Si. The genetic modifications seem to have no effect on the classification of the samples on the basis of the cation content.

Regarding the loadings, K is correlated with P ($r^2=0,820$), Mn ($r^2= 0,754$), Fe ($r^2=0,721$), Ca ($r^2=0,717$) and Zn ($r^2= 0,716$) and all these elements heavily load on F2; Na shows a particular behaviour in comparison to most of the measured cations. Good correlation is observed among Al, Mg and Si; they all load on F1, that seems to characterize Cr-stressed plants.

Considering the cation content in the water extracts (Figure 2b), the scores show the same behaviour as previously described for the total concentrations with the only exception of GR-DS that is grouped together with the other hydric-stressed plants. As observed for the total content, K shows correlation with Mn ($r^2=0.904$), P ($r^2=0.885$), Zn ($r^2=0.868$), Ca ($r^2= 0.781$) and Mg ($r^2=0,766$). No correlation is observed between Na and the other elements (the maximum value of correlation coefficient of Na is with K, $r^2 = 0.621$).

Introducing in the treatment the results obtained for the anions (Figure 2c), the resulting biplot confirms that the element content in the plants is especially influenced by the induced stress and not by the type of plant (wild or modified). In this case, the scores of the unstressed plants are in the first quarter of the biplot, well separated from all the stressed ones, the thermally-stressed plants are in the opposite region in comparison with most of the loadings, since they are generally

characterized by lower concentrations than the other ones; the plants affected by hydric and chemical stress form a group in the opposite region in comparison to the direction of the loadings of Cl^- and NO_3^- . While the cations mainly load on F1, the anions load on F2 (even if they are only partially correlated, $r^2=0,601$). Na shows a similar behavior to the other cations, but it is weakly correlated with the other analytes; in particular it shows the highest correlation coefficient with the other cationic marker, K ($r^2=0,621$). Remarkable correlations ($r^2 > 0,750$) are observed between K and Ca, Mg, P, Mn and Zn. The two anions are not strongly correlated with any cation.

Considering only K, Na, Cl^- and NO_3^- concentrations in the water extract (Figure 2d), we can see that all the loadings are oriented in the same direction along F1, with cations and anions having opposite loadings on F2. Their direction is toward the unstressed plants (in particular the anions), in opposite direction in comparison to thermally stressed plants; chemically and water-stressed plants are grouped together as just previously explained. The chemometric treatment of the whole data set and of the marker data alone is able to distinguish the same groups of samples, demonstrating how these analytes represent good markers to monitor the effect of stresses in plants. In particular, the plants are grouped on the basis of the stress treatment and not for the genetic modification.

Conclusions

The presence of genetic modification in *Nicotiana langsdorffii* causes a change of its capacity of absorption and of the facility of release of the considered elements, moreover genetic modifications cause a different plant response to stresses. In particular, the modification with *rolC* increases the capacity of the plant to adsorb a pollutant, Cr, and decreases its release in water. This behavior can be useful to develop modified plants able to be adopted for phytoremediation purposes. The presence of GR influences the inorganic composition of the plant, confirming its effect on the whole plant system, well known from a hormonal point of view.

Stress conditions have strong effects on the species content causing a reduction of their total concentrations, and of the extractable amount of cations and anions: the effect depends on the type of stress, and the considered analyte. For example, after applying a heat thermal stress, we observed a drastic reduction of cation content, while water deficit and the addition of Cr give rise to a less extensive decrease. A different effect was observed for the anions, which showed the greatest reduction in the case of water deficit.

Chemometrics helped us to value the whole dataset: differences among the considered plants derive in particular from the presence and type of stress. The selected markers seem to be a good choice, since their loadings are able to give the same information as that obtained considering the whole dataset.

The use of potentiometry for the determination of anions in plant extracts offers some advantages, as such as low costs and portability, over other techniques, but it is undoubtedly influenced by the presence of other species in solution. It might be used as for a first screening or to compare the effects of different parameters in a relative mode without absolute values.

The study of the behavior of plants and the effect of genetic modifications in different environmental conditions can be very useful i) to predict the variations induced in the normal development of the plants caused by environmental pollution and climatic variations, also taking into account the global warming that is affecting our Planet; and ii) to develop new plant species with higher resistance to particular conditions. The procedure adopted for this study can be followed in other investigations with different plants and different causes of stress.

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Table 1. Total cation contents (g/kg) in *Nicotiana langdorffii* and concentration percentages in stressed plants with respect to those present in the unstressed ones (in brackets).

| Analyte | WT | WT | WT | WT | GR | GR | GR | GR | rolC | rolC | rolC | rolC | Min | Max |
|---------|------|-----------------|----------------|----------------|--------------------|----------------|------------------------------|----------------------------|--------------------|-----------------|----------------|-----------------|--------------------|------|
| | | CS | DS | HS | | CS | DS | HS | | CS | DS | HS | | |
| K | 221 | 170 (-23%) | 202 (-9%) | 82.8 (-63%) | 213 | 181 (-15%) | 153 (-28%) | 67.8 (-68%) | 184 | 46.0 (-75%) | 172 (-7%) | 63.0 (-66%) | 63.0 | 221 |
| Na | 0.97 | 0.71 (-27%) | 0.62 (-36%) | 0.62 (-36%) | 1.15 | 0.69 (-40%) | 0.95 (-17%) | 1.78 (+54%) | 0.56 | 0.97 (+73%) | 0.48 (-14%) | 1.13 (+102%) | 0.48 | 1.78 |
| Cr | - | 0.19 | - | - | - | 0.15 | - | - | - | 0.27 | - | - | 0.15 | 0.27 |
| Al | 0.01 | 0.02 (+100%) | 0.01 (0%) | 0.01 (0%) | 2x10 ⁻³ | 0.01 (400%) | 4x10 ⁻³ (100%) | 2x10 ⁻³ (0%) | 5x10 ⁻³ | 0.01 (+100%) | 0.02 (300%) | 0.02 (300%) | 2x10 ⁻³ | 0.02 |
| Ca | 6.87 | 5.93 (-14%) | 7.39 (+8%) | 6.00 (-13%) | 5.94 | 4.40 (-26%) | 5.49 (-8%) | 4.48 (-2%) | 5.00 | 6.72 (+34%) | 5.64 (+13%) | 4.96 (-1%) | 4.38 | 7.39 |
| Fe | 0.40 | 0.31 (+22%) | 0.27 (-32%) | 0.16 (-60%) | 0.22 | 0.14 (-36%) | 0.17 (-23%) | 0.16 (-27%) | 0.28 | 0.04 (-86%) | 0.21 (-25%) | 0.18 (-36%) | 0.16 | 0.40 |
| Mg | 1.87 | 1.70 (-9%) | 1.94 (+4%) | 2.48 (+33%) | 1.28 | 1.20 (-6%) | 1.10 (-14%) | 1.87 (+46%) | 1.76 | 1.36 (-23%) | 1.63 (-7%) | 2.02 (+15%) | 1.10 | 2.48 |
| Mn | 0.36 | 0.24 (-33%) | 0.37 (+3%) | 0.23 (-36%) | 0.32 | 0.20 (-37%) | 0.22 (-31%) | 0.22 (-31%) | 0.29 | 0.21 (-28%) | 0.32 (+10%) | 0.22 (-24%) | 0.21 | 0.37 |

| Analyte | WT | WT | WT | WT | GR | GR | GR | GR | roIC | roIC | roIC | roIC | Min | Max |
|---------|------|-----------------|----------------|----------------|------|----------------|----------------|----------------|------|-----------------|----------------|-----------------|------|------|
| | | CS | DS | HS | | CS | DS | HS | | CS | DS | HS | | |
| P | 12.6 | 11.1 (-12%) | 12.8 (+2%) | 5.70 (-55%) | 10.9 | 11.6 (+6%) | 8.96 (-18%) | 5.73 (-55%) | 14.9 | 13.6 (-9%) | 14.9 (0%) | 4.65 (-69%) | 4.65 | 14.9 |
| Si | 0.09 | 0.31 (+244%) | 0.05 (-44%) | 0.35 (289%) | 0.06 | 0.07 (+17%) | 0.07 (+17%) | 0.11 (+83%) | 0.05 | 0.18 (+260%) | 0.06 (+20%) | 0.10 (+100%) | 0.05 | 0.35 |
| Zn | 0.15 | 0.10 (-33%) | 0.15 (0%) | 0.08 (-47%) | 0.13 | 0.10 (-23%) | 0.07 (-46%) | 0.10 (-23%) | 0.13 | 0.06 (-54%) | 0.12 (-8%) | 0.08 (-38%) | 0.07 | 0.15 |
| Total | 244 | 190 (-22%) | 226 (-7.4%) | 98.4 (-60%) | 233 | 203 (-13%) | 170 (-27%) | 82.3 (-65%) | 206 | 164 (-20%) | 195 (-51%) | 76.4 (-63%) | | |

Table 2. Marker concentrations (g/kg) found in lettuce extracts following procedures A, B and C and following the same procedure for all the extractants.

| Marker | Concentration (g/kg) obtained following the specific procedure (A, B or C) | Concentration (g/kg) obtained following the same procedure* |
|-----------------------------------|---|--|
| | HPW (procedure A) | HPW |
| Cl⁻ | 13.7 | 7.23 |
| NO₃⁻ | 53.2 | 35.7 |
| Na | 1.14 | 1.19 |
| K | 48.4 | 39.8 |
| | Sodium borate (procedure B) | Sodium borate |
| Cl⁻ | 0.98 | 8.25 |
| NO₃⁻ | 3.50 | 40.6 |
| Na | - | - |
| K | 3.47 | 41.2 |
| | Acetic Acid (procedure C) | Acetic Acid |
| Cl⁻ | 8.47 | 7.01 |
| NO₃⁻ | 40 | 35.2 |
| Na | 7.50 | 4.83 |
| K | 53.70 | 35.0 |

*0.5 g of sample + 25 ml of extractant solution. Contact time: 2h.

Table 3. Cation contents (g/kg) found in *Nicotiana langsdorfii* water extracts. Concentration percentages in stressed plants with respect to those present in the unstressed ones (in round brackets) and percentages of extraction with respect to the total concentrations (in square brackets).

| Analyte | WT | WT CS | WT DS | WT HS | GR | GR CS | GR DS | GR HS | rolC | rolC CS | rolC DS | rolC HS | Min | Max |
|---------|-------------------------------|--|--------------------------------------|---------------------------------------|------------------------------|---------------------------------------|---------------------------------------|--|----------------------------|--|--|---------------------------------------|----------------------------|-----------------------------|
| K | 36.9 [17%] ^T | 23.7 (-36%) [15%] ^T | 32.9 (-11%) [12%] | 25.4 (-31%) [14%] | 32.2 [12%] | 21.7 (-33%) [12%] | 29.2 (-9%) [16%] | 5.88 (-82%) [19%] | 22.4 [18%] | 17.8 (-21%) [7%] | 30.9 (+38%) [7%] | 4.77 (-79%) [8%] | 4.77 [7%] | 36.9 [19%] |
| Na | 0.96 [98.4%] | 0.48 (-50%) [68.2%] | 0.51 (-47%) [82.8%] | 0.51 (-47%) [49.0%] | 1.04 [90.5%] | 0.68 (-35%) [91.7%] | 0.58 (-44%) [60.9%] | 0.30 (-71%) [28.9%] | 0.32 [56.3%] | 0.38 (+19%) [59.8%] | 0.48 (+50%) [100%] | 0.40 (+25%) [35.7%] | 0.30 [28.9%] | 1.04 [100%] |
| Cr | - | 0.19 [34.8%] | - | - | - | 0.15 [26.7%] | - | - | - | 0.11 [19.5%] | - | - | 1.15 [26.7%] | 1.19 [34.8%] |
| Al | 1x10 ⁻³ [19%] | 1x10 ⁻⁴ (-90%) [15%] | 1x10 ⁻³ (=) [13%] | 3x10 ⁻³ (+200%) [2%] | 1x10 ⁻⁴ [3%] | 1x10 ⁻³ (+900%) [5%] | 1x10 ⁻³ (+900%) [5%] | 1x10 ⁻³ (+900%) [19%] | 1x10 ⁻³ [5%] | 1x10 ⁻³ (=) [38%] | 1x10 ⁻³ (=) [34%] | 2x10 ⁻³ (+100%) [1%] | 4x10 ⁻⁴ [1%] | 3x10 ⁻³ [38%] |
| Ca | 1.46 [21%] | 1.22 (-16%) [16%] | 1.56 (+7%) [15%] | 0.01 (-99%) [21%] | 0.94 [12%] | 0.72 (-23%) [15%] | 0.63 (-33%) [21%] | 0.67 (-29%) [11%] | 0.85 [17%] | 0.74 (-13%) [11%] | 0.94 (+11%) [7%] | 0.41 (-52%) [8%] | 0.31 [7%] | 1.56 [21%] |
| Fe | 3.9x10 ⁻² [10%] | 2.1x10 ⁻² (-46%) [5%] | 7x10 ⁻² (+79%) [2%] | 6x10 ⁻² (+54%) [7%] | 1.1x10 ⁻² [4%] | 1.3x10 ⁻² (18%) [6%] | 4x10 ⁻² (+264%) [3%] | 5x10 ⁻² (+355%) [2%] | 6x10 ⁻² [7%] | 1.4x10 ⁻² (-77%) [4%] | 1.4x10 ⁻² (-77%) [3%] | 7x10 ⁻² (+17%) [4%] | 4x10 ⁻³ [2%] | 0.04 [10%] |

| Analyte | WT | WT | | | GR | GR | | | rolC | rolC | | | Min | Max |
|---------|---------------------------------|----------------------|----------------------|----------------------|---------------------------------|----------------------|----------------------|----------------------|---------------------------------|----------------------|----------------------|----------------------|-------------------------------|-----------------|
| | | CS | DS | HS | | CS | DS | HS | | CS | DS | HS | | |
| Mg | 0.96 [51.2%] | 0.82 | 1.13 | 6.10 | 0.76 [59.0%] | 0.50 | 0.70 | 0.66 | 0.81 [45.9%] | 0.52 | 1.00 | 0.51 | 0.44 [23.6%] | 1.13 [63.7%] |
| | | (-15%) [48.4%] | (+18%) [58.1%] | (+535%) [26.6%] | | (-34%) [28.1%] | (-8%) [63.7%] | (-13%) [23.6%] | | (-36%) [36.7%] | (+23%) [61.1%] | (-37%) [25.2%] | | |
| Mn | 0.18 [49.1%] | 0.10 | 0.17 | 0.23 | 0.13 [41.2%] | 0.07 | 0.09 | 0.13 | 0.11 [100%] | 0.07 | 0.13 | 0.03 | 0.03 [15.6%] | 0.18 [100%] |
| | | (-44%) [42.8%] | (-6%) [47.2%] | (+28%) [26.1%] | | (-46%) [29.8%] | (-31%) [41.5%] | (=) [15.6%] | | (-36%) [30.5%] | (+18%) [40.5%] | (-73%) [17.8%] | | |
| P | 8.64 [68.6%] | 6.53 | 9.89 | 1.84 | 8.35 [76.4%] | 5.71 | 12.7 | 1.55 | 8.59 [100%] | 7.32 | 11.6 | 1.36 | 1.36 [27.1%] | 12.7 [100%] |
| | | (-24%) [58.9%] | (+14%) [77.0%] | (-79%) [27.1%] | | (-32%) [49.3%] | (+52%) [100%] | (-81%) [27.1%] | | (-15%) [53.8%] | (+35%) [77.8%] | (-84%) [29.2%] | | |
| Si | 3.0x10 ⁻² [34.2%] | 3.8x10 ⁻² | 1.6x10 ⁻² | 1.7x10 ⁻² | 2.4x10 ⁻² [38.5%] | 2.6x10 ⁻² | 1.5x10 ⁻² | 0.6x10 ⁻² | 1.5x10 ⁻² [29.3%] | 1.5x10 ⁻² | 2.2x10 ⁻² | 0.6x10 ⁻² | 6x10 ⁻³ [5.00%] | 0.04 [38.4%] |
| | | (+27%) [12.2%] | (-47%) [34.1%] | (-43%) [5.00%] | | (+8%) [7.84%] | (-38%) [20.8%] | (-75%) [5.26%] | | (=) [8.31%] | (+47%) [34.0%] | (-60%) [5.44%] | | |
| Zn | 0.08 [50.7%] | 0.05 | 0.05 | 0.06 | 0.05 [40.4%] | 0.03 | 0.05 | 0.06 | 0.06 [19.8%] | 0.02 | 0.06 | 0.02 | 0.02 [19.8%] | 0.08 [66.5%] |
| | | (-38%) [52.8%] | (-38%) [36.4%] | (-25%) [31.6%] | | (-40%) [30.6%] | (=) [66.5%] | (+20%) [20.5%] | | (-67%) [33.1%] | (=) [47.2%] | (-67%) [30.5%] | | |
| Total | 49.2 | 33.0 | 46.3 | 9.17 | 43.5 | 29.5 | 44.0 | 7.68 | 33.1 | 26.8 | 45.2 | 7.58 | | |

Table 4. Anion contents (g/kg) found in *Nicotiana langsdorffii* water extracts using IC and ISE and concentration percentages in stressed plants with respect to those present in the unstressed ones (in brackets).

| Analyte | WT | WT | WT | WT | GR | GR | GR | GR | rolC | rolC | rolC | rolC |
|------------------------------|------|---------------------|---------------------|---------------------|------|---------------------|---------------------|---------------------|------|---------------------|---------------------|---------------------|
| | | CS | DS | HS | | CS | DS | HS | | CS | DS | HS |
| IC | | | | | | | | | | | | |
| Cl ⁻ | 16.3 | 8.51 | 2.40 | 3.35 | 14.6 | 10.8 | 12.0 | 6.11 | 15.8 | 8.73 | 10.1 | 7.2 |
| | | (-48%) ^U | (-85%) ^U | (-67%) ^U | | (-23%) ^U | (-26%) ^U | (-56%) ^U | | (-44%) ^U | (-36%) ^U | (-54%) ^U |
| NO ₃ ⁻ | 63.8 | 36.0 | 11.6 | 55.2 | 54.0 | 30.9 | 45.3 | 43.0 | 55.8 | 30.3 | 36.2 | 55.3 |
| | | (-44%) ^U | (-81%) ^U | (-13%) ^U | | (-43%) ^U | (-29%) ^U | (-20%) ^U | | (-46%) ^U | (-35%) ^U | (-1%) ^U |
| ISE | | | | | | | | | | | | |
| Cl ⁻ | 15.6 | 15.2 | 12.2 | 13.3 | 16.4 | 13.1 | 16.6 | 12.8 | 17.7 | 17.6 | 17.7 | 16.8 |
| | | (-3%) | (-22%) | (-13%) | | (-20%) | (+6) | (-22%) | | (-1%) | (0%) | (-5%) |
| NO ₃ ⁻ | 77.4 | 45.0 | 12.5 | 68.6 | 39.0 | 12.6 | 37.9 | 35.7 | 39.3 | 9.09 | 29.4 | 40.8 |
| | | (-42%) | (-83%) | (-13%) | | (-68%) | (-51%) | (-8%) | | (-77%) | (-25%) | (+4) |

Table 5. Equations of the calibration curves and R^2 values obtained for the determination of Cl^- and NO_3^- in HPW and in sample water extracts (two standard additions of each anion) using IC and potentiometry with ISE.

| | | Cl^- | NO_3^- | | |
|-----------|---------------------|--------------------------|-----------------|--------------------------|----------|
| Technique | Matrix | Equation | R^2 | Equation | R^2 |
| IC | HPW | $y = 221.824x + 21.778$ | 0.9999 | $y = 116.543x + 13.071$ | > 0.9999 |
| IC | WT extract | $y = 160.210x + 136.387$ | 0.9968 | $y = 148.971x + 355.313$ | > 0.9999 |
| IC | GR extract | $y = 150.022x + 162.058$ | 0.9984 | $y = 145.094x + 657.251$ | 0.9993 |
| IC | <i>rolC</i> extract | $y = 153.036x + 154.193$ | 0.9998 | $y = 141.568x + 429.300$ | 0.9937 |
| ISE | HPW extract | $y = -56.584x + 189.1$ | 0.9967 | $y = -60.33x + 93.68$ | 0.9837 |
| ISE | WT extract | $y = -36.54x + 148.5$ | > | $y = -31.71x + 52.5$ | > 0.9999 |

| | | | | | |
|-----|--------------|-----------------------|--------|----------------------|----------|
| | | | 0.9999 | | |
| ISE | GR extract | $y = -36.54x + 147.5$ | > | $y = -28.08x + 58.5$ | 0.985 |
| | | | 0.9999 | | |
| ISE | rolC extract | $y = -36.54x + 146.5$ | > | $y = -23.25x + 31.0$ | > 0.9999 |
| | | | 0.9999 | | |

Table 6. Contents (g/kg) of Cl^- and NO_3^- measured in water by ISE in the presence of potential interfering agents. Expected concentration: 4,00 g/kg

| Analyte | HPW | Br^- | I^- | CH_3COO^- | NO_3^- | Cl^- |
|-----------------|------|---------------|--------------|---------------------------|-----------------|---------------|
| Cl^- | 4.00 | 4.09 | 6.08 | 3.90 | 3.97 | - |
| NO_3^- | 4.05 | 4.07 | - | 6.40 | - | 4.08 |