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Rare Earth Ions doped ZnO: Synthesis, Characterization and Preliminary Photoactivity Assessment

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Abstract

This work reports the effect of doping zinc oxide with lanthanide ions on structural, EPR and UV visible properties. Bare and doped samples were synthesized using the simple and green hydrothermal process. Different rare earth ions (RE= La, Ce, Pr, Er and Yb) with 1% molar ratio RE/Zn were used. The samples have been studied using X Ray Diffraction, Scanning Electron Microscopy (SEM), Transmission Electron Microscopy (TEM) and UV visible diffuse reflectance spectroscopy. Finally, electron paramagnetic resonance (EPR) spectroscopy, was used to assess the materials photoactivity under UV irradiation, both in solid state, to see the charge carriers' generation and in solution, evaluating the OH• radical formation using the DMPO (5,5-Dimethyl-1-Pyrroline-N-Oxide) spin trapping technique. The results suggest that the synthesized materials could be interesting systems for the photocatalytic abatement of emerging organic persistent pollutants in wastewater treatment plants.

1. Introduction

Recently, implementation of water remediation and re-use is gaining attention rapidly world-wide due to the water scarcity occurred as a result of climate change and poor water resource management (i.e. limited access to clean water resources and water demands exceed the available resources). Access to clean water is becoming an increasing problem in an expanding global economy and population countries. [1] Persistent organic pollutants and the new emerging pollutants are considered dangerous for living organisms due to their acute toxicity and bio-recalcitrant nature, in fact they are resistant to environment degradation and have been continuously released into the environment.[2, 3] In particular,

the biological process is usually ineffective because of a self-inhibitory effect on the microorganisms in the presence of these compounds. [4] For this reason, photocatalytic treatment represents an advanced oxidation technology able to remove organic pollutants thanks to the efficient degradation rate and high mineralization efficiency that leads to CO_2 , H_2O , and other minerals as final products. Among various semiconductor materials, ZnO has received considerable scientific interest as an alternative to TiO_2 [5, 6]. It was also reported by some of us that ZnO was more efficient than TiO_2 in photocatalytic degradation of some emerging pollutants in aqueous solution[7]

ZnO is considered as a benchmark catalyst due to its non-toxic nature, physical and chemical stability coupled with its low cost. [8] One major drawback of existing technology is the low quantum efficiency of ZnO, which is due to very fast recombination of photogenerated electron–hole pairs. In order to improve the photocatalytic activity of ZnO, several modification methods have been developed. It has been reported that photocatalytic performance of semiconductors can be improved by doping with certain cations, which can act as electron trapping agent to decrease the electron-hole recombination rate. [9, 10, 11] The photocatalytic activity of ZnO can be significantly enhanced by doping with the lanthanide ions having 4f configuration, among them, cerium-doped ZnO nanomaterials exhibited versatile properties and much attention has been focused to explore them for possible applications in various technologies. [12, 13] Doping RE ions into ZnO produces a significant improvement in the charge separation between photogenerated electrons and holes. [14-16] The efficiency of cerium-doped ZnO towards degradation of water pollutants and emerging contaminants is already well known: phenol, iodinated X ray contrast agents, namely iopamidol, iopromide and diatrizoate, emerging contaminants known to be scarcely abated in the wastewater treatment plants, and acesulfame K, an artificial sweetener, were efficiently abated. [7, 17] The role in the photoactivity of the f orbitals, which in the case of cerium and lanthanum are empty and for the others lanthanides are partly filled, is still under debate. Moreover, the presence of lanthanide ions into the ZnO structure has never been demonstrated, thus the understanding of the role of the dopants in the host oxide is crucial.

In this work different materials based on ZnO doped with a low amount of RE elements (RE = La, Ce, Pr, Er and Yb), aimed to develop new photocatalysts, were synthesized. Pure and doped ZnO with low RE loading (1% molar) were prepared via hydrothermal process, a simple and green way to obtain controlled nanostructures at relatively low temperatures. Compared with other techniques this method is fast, cheap and highly efficient. The precursors used are easily available on the market and the amount

of rare earth compound employed is very low and does not affect dramatically the final price of the catalyst. The morphological and textural features, the crystal structure and optical properties of the synthesized materials were investigated and their photoactivity was assessed by EPR spectroscopy. The use of a spin probe molecule like DMPO (5,5-Dimethyl-1-Pyrroline-N-Oxide) allowed the evaluation of oxidant species present in solution.

2. Experimental

2.1 Sample preparation

All reactants employed in this work were purchased from Sigma-Aldrich (Milan Italy) and used without any further purification treatment. All the samples have been synthesized following hydrothermal synthesis. Bare ZnO sample was synthesized starting from a 1 M water solution of zinc nitrate. Then a 4 M NaOH solution was added dropwise until the pH was 11. Finally, the solution was transferred into a PTFE-lined stainless steel 100 mL autoclave (filling 70%), and treated at 175°C for 15 hours. The product was centrifuged, washed three times and dried at 70°C overnight. The doped samples (RE=Ce, La, Er, Yb, Pr, molar concentration 1 %) were prepared adding to the starting solution the RE precursor salts in stoichiometric amount ($\text{RE}(\text{NO}_3)_3 \cdot x\text{H}_2\text{O}$, $x=5$ or 6 depending on the RE). Hereafter the doped samples will be labelled as ZnO-RE.

2.2 Characterization techniques

X-rays powder diffraction (XRPD) patterns were recorded using a PANalytical PW3040/60 X'Pert PRO MPD, Lissone (MI) Italy (45kV, 40mA) with a copper $K\alpha$ radiation source (0.15418 nm). Samples were scanned continuously in the 2θ range between 10° and 100°. The X'Pert High-Score software was used to identify the mineral phases present in the samples. A Rietveld refinement was performed on the acquired patterns using the MAUD 2.2 software[18].

The morphology of the materials was studied by scanning electron microscopy (SEM) using a JEOL microscope (JSM-6510-LV, JEOL, Tokyo, Japan) and by a transmission electron microscopy (TEM) using a Jeol JEM 3010 (300 kV) microscope.

The UV-Vis absorption spectra were recorded using a Varian Cary 5000 spectrophotometer, coupled with an integration sphere for diffuse reflectance studies (DRS), using a Carywin-UV/scan software. A sample of PTFE with 100% reflectance was used as reference.

Electron Paramagnetic Resonance Spectroscopy (EPR) was employed, using an X-band CW-EPR Bruker EMX spectrometer equipped with a cylindrical cavity operating at 100 KHz field modulation. The effect of UV light on EPR spectra was investigated using a 1600 W Xenon lamp (Oriol Instruments) equipped with an IR water filter. The samples were treated before the measurement as follows: first, the sample was heated under vacuum at 300°C for 30 minutes, then heated under oxygen pressure (50 mbar) at 500°C for 1 hour. The formation of OH[•] radical species upon irradiation (500 W, Xenon lamp, Oriol Instruments) was monitored by EPR spin trapping technique using a Miniscope 100 spectrometer from Magnetech and using DMPO (5,5-dimethyl-1-pyrroline-N-oxide) as spin trapping agent.

3. Results and Discussion

3.1. Structural, Morphological and Spectroscopic characterization

XRPD patterns of the samples investigated in this work are shown in Fig.1. The pattern of bare ZnO shows peaks corresponding to (1 0 0), (0 0 2), (1 0 1), (1 0 2), (1 1 0), (1 0 3), (2 0 0), (1 1 2) and (2 0 1) planes of the wurtzite hexagonal phase of ZnO [19].

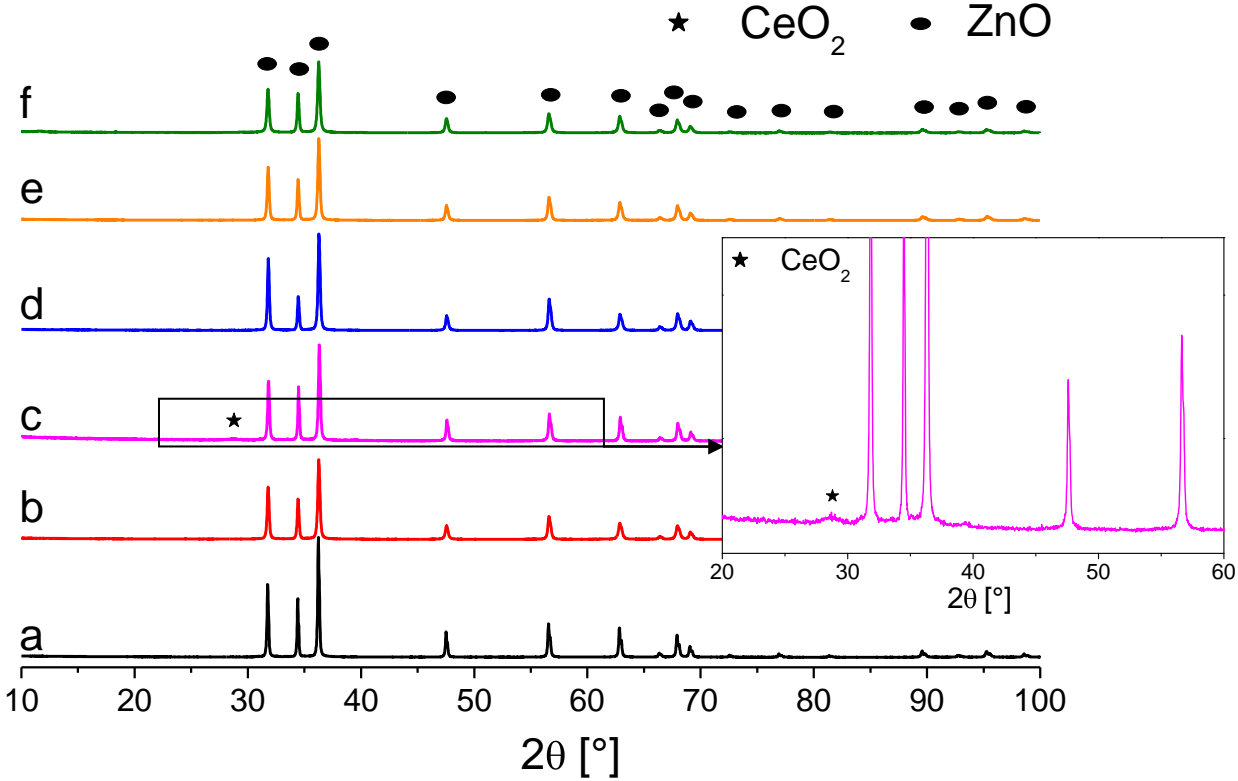


Figure 1. XRD patterns of a) ZnO, b) ZnO-La, c) ZnO-Ce, d) ZnO-Pr, e) ZnO-Er, f) ZnO-Yb. Inset: enlargement of the ZnO-Ce pattern.

The introduction of dopants does not bring relevant modifications in the diffraction patterns of ZnO with the exception of the sample ZnO-Ce in which we observed the formation of a tiny amount of the CeO₂ crystalline phase (see Fig. 1, inset). On the XRD patterns we performed a Rietveld refinement using the MAUD software[18]. Table 1 lists the obtained average crystallite size and the lattice parameters for all the detected phases. Moreover, the difference in the lattice parameters (*a* and *c*) due to the addition of lanthanides ions inside the ZnO structure is also evidenced ($\Delta a/c = a/c(\text{ZnO-RE}) - a/c(\text{ZnO})$). The lattice parameters of the ZnO phase are slightly influenced by the introduction of lanthanides, bringing to a limited contraction of the unit cell for all the doped samples respect to that of pure ZnO. The average crystallite size of bare zinc oxide is quite large (*d*=256nm), the insertion of RE ions in the structure causes in general a decrease in the average crystallite size, the sample ZnO-Yb being the mostly affected, while the sample Ce-ZnO is the less affected, probably because of the formation of the cerium oxide phase. This effect is due to the insertion of the foreign atoms in the structure, that limits the crystallite growth.

Table 1. Cell parameters and crystallite size obtained from Rietveld refinement of the XRD patterns of bare and doped ZnO. The difference between the parameters of the doped samples and the bare ZnO are also evidenced as Δ .

Sample	Phase	<i>a</i> (Å)	Δa	<i>c</i> (Å)	Δc	<i>d</i> (nm)	Δd
ZnO	ZnO	3,2554	0,0000	5,2143	0,0000	256	0
ZnO-La	ZnO	3,2528	-0,0026	5,2125	-0,0018	124	-132
ZnO-Ce	ZnO	3,2528	-0,0026	5,2111	-0,0032	163	-93
	CeO ₂	5,4111				10	
ZnO-Pr	ZnO	3,2529	-0,0026	5,2125	-0,0018	139	-117
ZnO-Er	ZnO	3,2529	-0,0025	5,2123	-0,0020	127	-129
ZnO-Yb	ZnO	3,2536	-0,0019	5,2122	-0,0021	93	-163

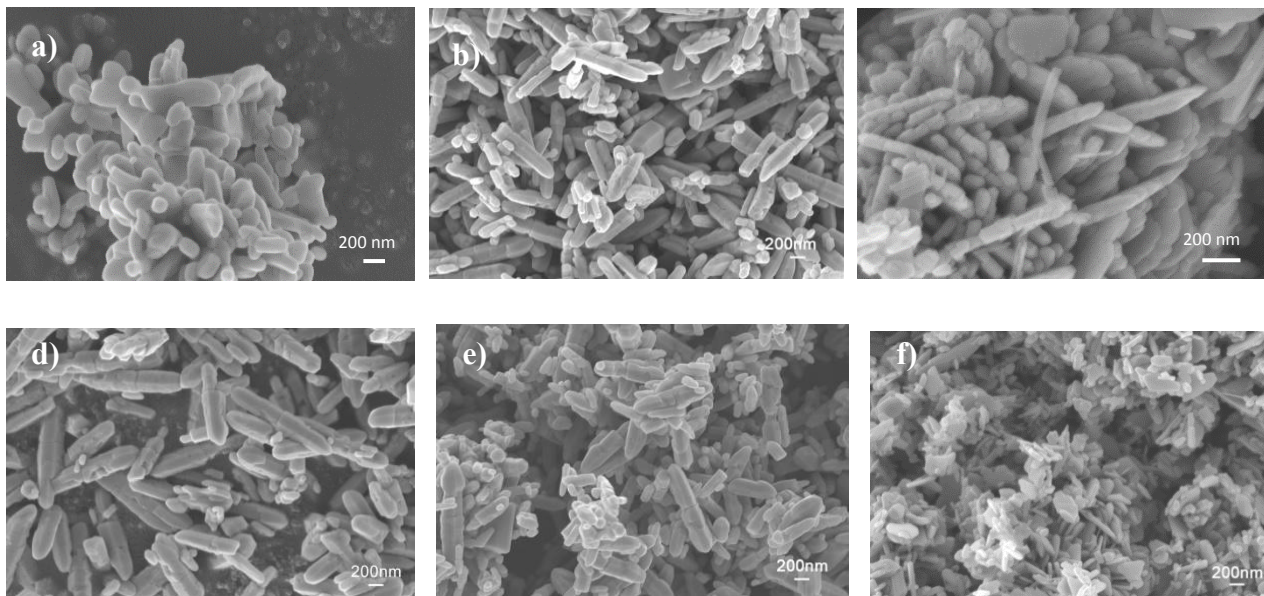
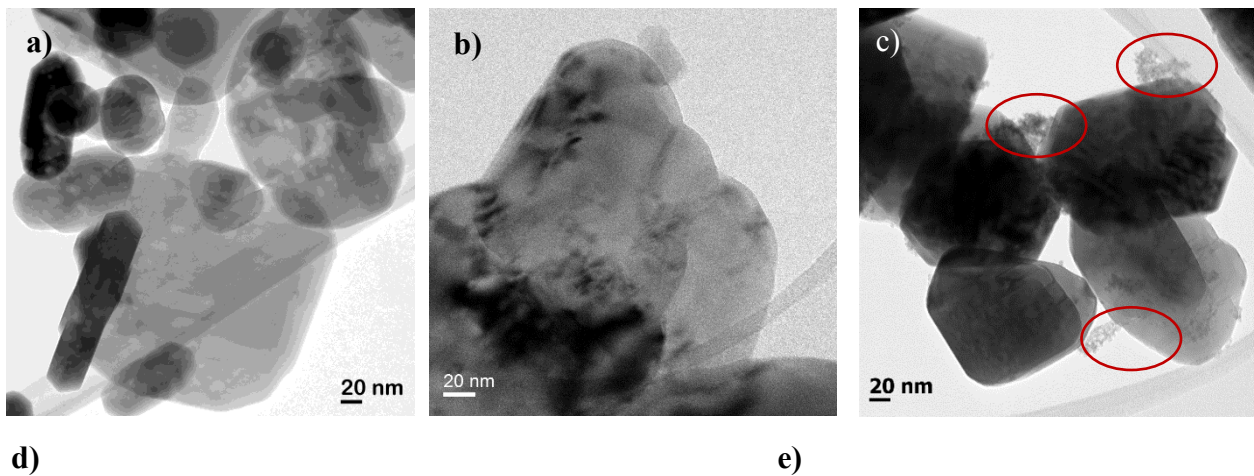


Figure 2: SEM micrographs of a) ZnO, b) ZnO-La, c) ZnO-Ce, d) ZnO-Pr, e) ZnO-Er, f) ZnO-Yb.

Figure 2 depicts the surface morphology of pure (a) and doped (b-f) ZnO catalysts. The surface of all the samples, with the only exception for ZnO-Yb, exhibits similar morphology: rods with 1.0-1.5 μm of length are randomly oriented and aggregated. The sample doped with Yb (Fig. 2f) exhibits an intermediate shape between platelets and rods with smaller particles as already revealed by the Rietveld refinement.



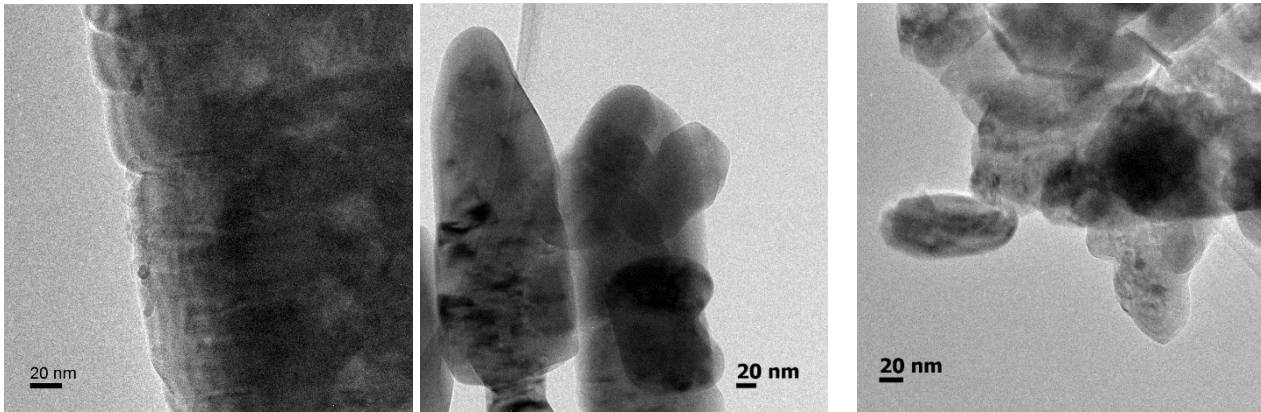


Figure 3. TEM micrographs of a) ZnO, b) ZnO-La, c) ZnO-Ce, d) ZnO-Pr, e) ZnO-Er, f) ZnO-Yb

Figure 3 reports TEM micrographs of ZnO and doped materials. Only in the case of cerium doped materials (Fig. 3b) is possible to observe clearly small particles anchored at the surface of ZnO platelet (highlighted with red circles). This observation is in agreement with the presence of CeO_2 (Ce oxidation state: +4) diffraction peaks in the XRPD analysis and it is further confirmed by the previously reported elemental EDS mapping that shows the Ce confinement [17]. Lanthanum, praseodymium and erbium are all stable as Ln^{3+} ions, and the ionic radii of the first two are quite larger than that of Zn^{2+} . The micrographs related to these dopants reveal a morphology that is close to that of pure ZnO, and it's not possible to see evidence of other phases except to some shadows that can be related to both ZnO defects and/or XRD silent amorphous phases. ZnO-Yb micrographs point out the presence of smaller particles, that could be correlated to the large decrease of the crystallite size (see table 1). Yb ions can be stable also with oxidation number +2 (isovalent with Zn^{2+} ions), moreover they are the smallest ions among all the dopants series, thus their insertion in the ZnO matrix is possible and likely.

The influence of RE doping on the optical absorption of the various samples has been investigated with Diffuse Reflectance (DR) UV-Vis spectroscopy. Figure 4 compares the spectra recorded for pure zinc with those of the variously doped samples. The spectrum of bare ZnO is well known and it is dominated by the band gap transition occurring at about 400 nm (3.3 eV) that is due to the electrons excitation from the valence band (VB) to the conduction band (CB) [20]. The use of different types of dopant does not dramatically affects the band gap of the materials. Indeed we calculated the energy gap (E_g) value for all the samples using the Tauc plot, *i.e.* by linearization of the plot reporting $(\alpha h\nu)^2$ vs $h\nu$, typical of direct band gap transitions [21]. The values obtained are very similar for all samples, indicating that the RE insertion in the ZnO matrix does not affect directly the VB \rightarrow CB transition of zinc oxide.

Nevertheless, in the case of samples containing Pr, Er and Yb it is possible to detect the presence of a multitude of defined absorption bands, related to the f-f transitions of Pr, Er and Yb ion states. The presence of these bands indicates that the RE ions are actually present in the samples (despite the other characterizations results) and well diluted in the matrix, preserving their optical properties [22-24]. Finally, for all the doped samples except for ZnO-Yb, we observed a slight tail of absorption extending to the visible region (see Fig. 4, panel B), this effect is more evident for the ZnO-Ce sample. ZnO-Yb sample presents a deep absorption in the NIR region. The intensity of this signal respect to the others suggests that for the other samples the nominal concentration was not achieved.

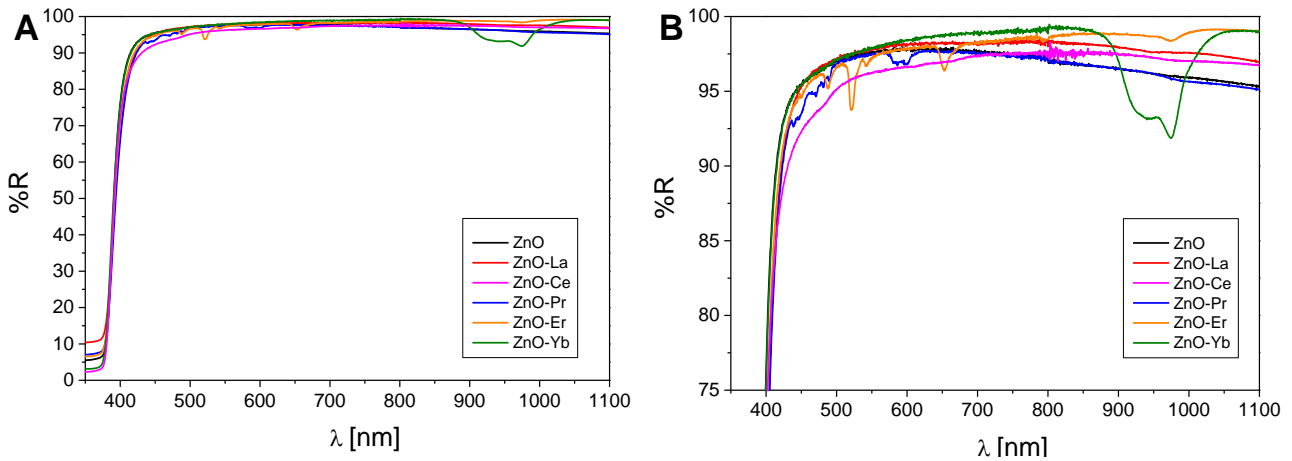


Figure 4. Percentage of Reflectance spectra of ZnO and RE doped ZnO, Panel B is the enlargement of panel A.

3.2 EPR spectroscopy

In this work EPR spectroscopy was used to investigate the effect of UV light on the synthesized samples. This technique allows the assessment of the photoactivity of the samples, monitoring the charge carriers' separation under irradiation. All the spectra were recorded after activation of the samples as described in section 2.1.

Fig. 5 reports the EPR spectra before (Panel A) and during UV irradiation (Panel B), respectively. Panel A reveals that before irradiation, all samples are characterized by an isotropic signal at $g=2.014$, due to the synthesis procedure, and a broader and slightly asymmetric signal around $g=1.96$. The assignment for the signal at $g=2.014$ is not straightforward, it was already reported in the past, without a clear identification. [25] The signal around $g=1.96$ has been extensively studied, and its assignment is still

under debate: oxygen vacancies [26-30], zinc interstitial [27, 31] and so interstitial hydrogen atoms acting as shallow donors. [32] Computational studies corroborate the hypothesis that this signal is generated by shallow donor species. [33, 34] The signal broadening that we observed in our spectra it is probably due to different environments felt by the unpaired electron [34, 35] indicating a certain degree of heterogeneity in the samples.

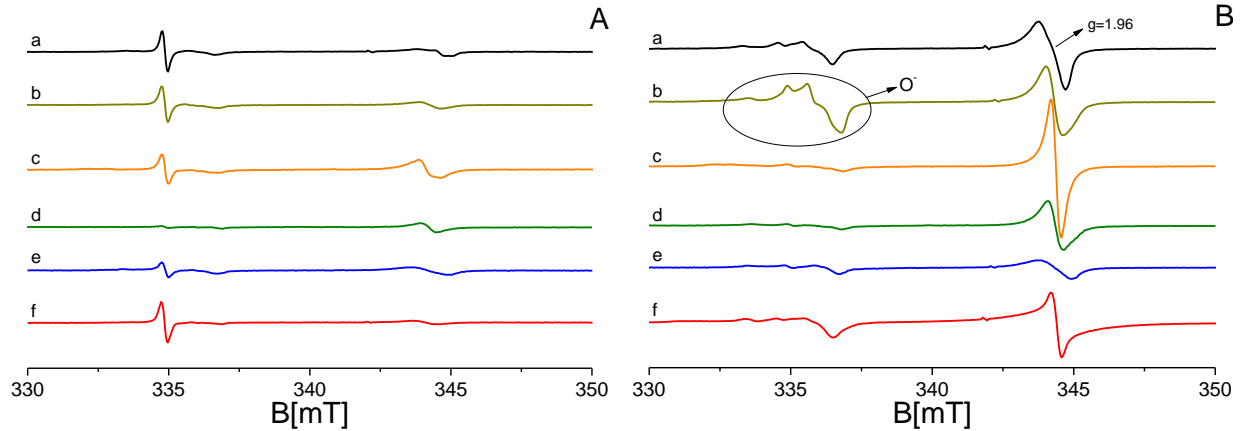


Figure 5. EPR spectra recorded at 77K of activated samples before (Panel A) and during (Panel B) UV irradiation: a) ZnO, b) ZnO-La, c) ZnO-Ce, d) ZnO-Pr, e) ZnO-Er, f) ZnO-Yb.

Figure 5 (panel B) reports the spectra obtained during UV irradiation. The absorption of photons allows the promotion of electrons from the valence band to the conduction band, with the consequent formation of holes in the valence band. [26] The photogenerated charge carriers can be stabilized at defect sites, forming paramagnetic species, thus EPR active.

For all samples, with different extent for the various dopants, the intensity of the signal at $g=1.96$ increases, indicating that irradiation probably formed trapped electronic species. The photogenerated holes are stabilized at O^{2-} sites, forming the well-known $O^{\bullet-}$ ions (as highlighted with a circle in Figure 5, panel B). [36, 37] This signal with $g_{//}=2.023$ and $g_{\perp}=2.003$ is formed mainly by the samples ZnO, ZnO-La and ZnO-Yb. These results indicate that the synthesized samples are active under irradiation, being capable of forming electron-hole pairs. Among all the synthesized materials ZnO-La showed the formation of higher signal due to the formation of holes and for this reason we decided to test its potential photocatalytic activity irradiating an aqueous suspension of the solid with the polychromatic light emitted by a Xe lamp and using DMPO (5,5-dimethyl-1-pyrrolineN-oxide) as spin-trapping agent to detect the formation of hydroxyl radicals (OH^{\bullet}) in solution. For comparison the same experiment has been carried out also for the bare ZnO. The DMPO molecule (diamagnetic) forms a stable paramagnet-

ic adduct with OH radicals (DMPO/OH[•]), detectable by the EPR technique [38]. The hydroxyl radicals are responsible for the oxidative degradation of organic molecules in photocatalytic reactions, and thus their detection is a good preliminary test for the potential photocatalytic activity. Hydroxyl radicals are formed under irradiation according to the following process:

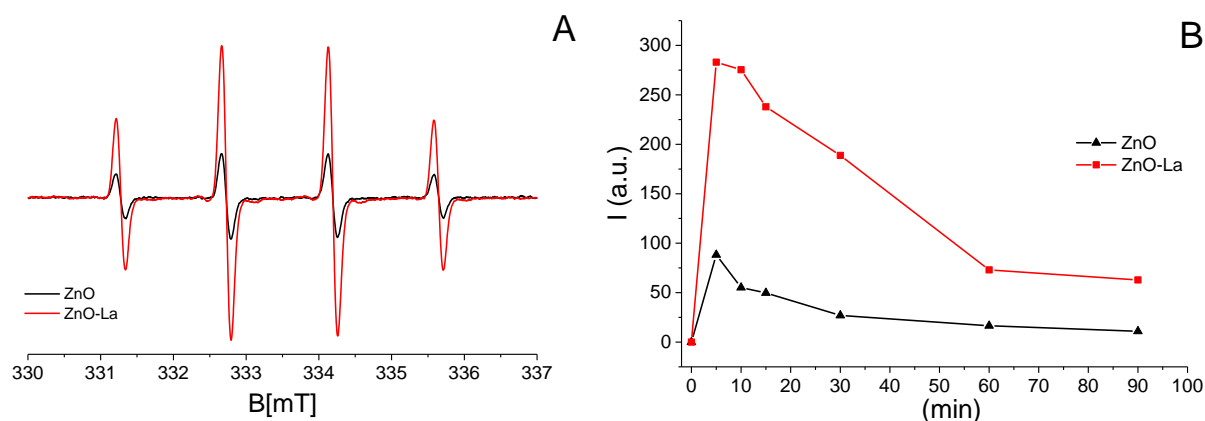
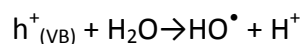


Figure 6. Panel (A): EPR spectra of the DMPO/OH[•] adduct produced by irradiation of an aqueous suspension of the samples ZnO and ZnO-La with UV light for 5 min; panel (B): integrated intensities of the EPR spectra as a function of irradiation time.

The aqueous suspension of bare ZnO forms OH radicals under UV irradiation, but the suspension with the doped sample exhibits signals that are three times more intense than the pure sample. Figure 6 (panel (A)) reports the classic four-lines EPR spectrum of the DMPO/OH[•] adduct, as observed after 5 min of irradiation for both samples. This irradiation time corresponds to the maximum intensity reached by the EPR spectra, meaning the higher amount of hydroxyl radicals formed. Panel (B) shows the integrated area values of the EPR spectra (which are proportional to the radical formation in solution) as a function of the irradiation time. The results show that the La-doped sample can generate hydroxyl radicals in solution when irradiated, with a higher efficiency compared with the bare oxide. The doped samples are thus good potential photocatalysts for the abatement of persistent emerging organic pollutants and will be the focus of a forthcoming paper.

4. Conclusions

Different photocatalysts based on ZnO and doped ZnO have been synthesized. The dopant used were all lanthanides: La, Ce, Pr, Er and Yb with 1% molar of nominal amount. All the prepared materials were fully characterized via structural, morphological and spectroscopic techniques. Dopants behave differently according to their oxidation number and to the number of electrons in their f orbitals. Only cerium creates a new oxide phase, all the others are probably inserted in the ZnO matrix. Among all the prepared materials ZnO-La, that showed the highest interaction with light at the solid state, has been chosen for the evaluation of OH radicals production in water suspension. The concentration of OH radicals has been evaluated using the spin trap molecule DMPO able to trap the OH^{*} generated by the UV irradiation. The doped sample is three times more active under UV light than the bare ZnO, demonstrating that lanthanide-doped zinc oxide can be a promising system to consider in the photocatalytic abatement of organic pollutants.

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