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(Article begins on next page)



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High Zn/Al ratios enhance dehydrogenation vs hydrogen transfer reactions of Zn-ZSM-5 catalytic systems in methanol conversion to aromatics.

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Abstract

Two series of Zn-ZSM-5 catalysts were prepared by ion exchanging two commercial zeolites with different Si/Al ratios (40 and 15) with increasing Zn loadings. The nature of the Zn sites in the zeolite was studied by spectroscopy using laboratory and synchrotron techniques. All the evidences suggest that catalytic activity is associated with $[\text{Zn}(\text{H}_2\text{O})_n(\text{OH})]^+$ species located in the exchange positions of the materials with little or no contribution of ZnO or metallic Zn. The effect of Zn/Al ratio on their catalytic performance in methanol conversion to aromatics has been investigated. In all cases, higher Zn content causes an increase in the yield of aromatics while keeping the production of alkanes low. For similar Zn contents, high densities of Al sites favour the hydrogen transfer reactions and alkane formation whereas in samples with low Al contents, and thus higher Zn/Al ratio, the dehydrogenation reactions in which molecular hydrogen is released are favoured.

Keywords: Zn-ZSM-5, Methanol to aromatics, Acidity, Zn sites, Ion exchange

1. Introduction

Light aromatic hydrocarbons, such as benzene, toluene and xylenes (BTX) are important intermediates in the petrochemical industry. They are commonly used as solvents and for the production of high-value secondary products such as plastics, paints, cosmetics or pharmaceuticals. BTX are traditionally obtained from naphtha via catalytic reforming or steam cracking with a global production in the range of 100 million metric tons per year that is projected to grow at a rate of 3.5-4 % in the next five years [1]. The increase in the global demand of aromatics together with the notion of finite crude oil reserves and the consequential necessity to implement an oil-independent economy makes it crucial to develop alternative processes for the production of BTX from non-petroleum sources. In this context, the possibility to directly convert methanol into aromatics (MTA) over zeolitic catalysts, that was described already 25 years ago [2, 3], is attracting a renewed interest from both industrial and academic researchers.

The methanol to aromatics reaction (MTA) falls into the general group of the so-called methanol to hydrocarbons (MTH) transformations, which allow obtaining fuels and chemicals avoiding the use of petroleum. MTH reactions have been widely investigated since their discovery by Mobil in 1976 [4]. The discovery attracted great attention due to the versatility of the process and the availability of methanol from rather abundant and diverse sources, such as coal and natural gas [5, 6], biomass or even CO₂ [7, 8]. The MTH process involves multiple reaction steps within a quite complex mechanism, which has been intensively studied for decades [9-14]. Today, the dual cycle hydrocarbon pool mechanism is generally accepted [11]. According to this mechanism, the process is an autocatalytic reaction in which both alkene and aromatic molecules act as intermediates in two inter-related cycles involving many reactions such as methylation, cracking,

hydride transfer, cyclization and dealkylation. Eventually, the reaction between the hydrocarbon intermediates and methanol molecules or the combination of hydrocarbons can lead to the formation of molecules that are too large to be able to diffuse through the zeolite pores and remain trapped within the framework, evolving to form carbonaceous deposits that are the responsible for the deactivation of the catalyst [15-20].

Initially, MTH technology was considered as a powerful alternative to oil-based routes for the production of gasoline compatible with the existing infrastructure. However, further investigations revealed that it is possible to selectively obtain practically any kind of hydrocarbons, including gasoline (MTG) light olefins (MTO), propylene (MTP) or aromatics (MTA), by modifying the catalysts and adjusting the operation conditions [21, 22]. Aromatization involves cyclization and hydrogen transfer reactions catalyzed by the Brønsted acid sites of the zeolite [11, 22, 23]. Olefins act as hydrogen acceptors in the bimolecular hydrogen transfer reaction, forming alkanes as hydrogen rich byproducts. In the case of MTA, the zeolite used as catalyst is partially doped with transition metals, forming a bifunctional system with Brønsted and Lewis acid sites from the zeolite and new metal Lewis sites. The goal of the incorporation of transition metals is to favour the aromatization reactions via additional dehydrogenation functionality. Recently, it has been reported that the hydrogen transfer may also be induced by methanol rather than by olefins, and not just Brønsted but also Lewis sites of the zeolite are proposed to be involved in the hydrogen transfer reactions for the formation of aromatics (and alkanes) [24]. This mechanism involves the presence of oxygenate intermediates such as formaldehyde. When transition metal ions are incorporated in the zeolitic catalyst, a new possibility appears, and the aromatization can take place at the metal Lewis acid sites by dehydrogenation instead of hydrogen transfer, thus forming molecular hydrogen as

byproduct and keeping the production of alkanes low. This is interesting, first, because it means that a higher proportion of the carbon atoms from the reactants would be tuned into highly demanded products and, second, because of the intrinsic value of H₂ as byproduct.

Several combinations of zeolites and transition metals have been investigated as MTA catalysts, with ZSM-5 being one of the most interesting topology in terms of stability and selectivity for the process, and Zn the metal that provides the highest yield of BTX products [3, 25-28]. A full picture of the interplay between acid sites and metal sites is still missing, but it is known that the location of the metal sites (on the external crystal surface or within the micropores) and the nature of the sites formed have a great influence in the catalytic performance of the final material [28-32]. While only few contributions report the stabilization of unusual Zn⁺ species in the zeolitic system [33], metallic Zn, ZnO agglomerates or other Zn²⁺ species are generally proposed as the possible Zn sites formed in the zeolite. According to this proposal, two relatively close Al sites, and hence a quite low Si/Al ratio, would be needed to stabilize the two positive charges of the exchanged cation. However, to compensate the negative charge generated by the presence of an isolated Al³⁺ in the inorganic framework, Zn²⁺ cations can also be incorporated as **Z**[Zn²⁺(OH)]⁺ species [34], where **Z** indicates coordination to framework oxygens (O_{fw}) in the proximity of one framework aluminium (Al_{fw}) site. Again, depending on the elemental composition and the spatial distribution in the zeolite, such hydroxylated species are expected to undergo dehydration reactions at high temperature, and evolve into new Zn²⁺ species [31]. In samples with low Zn content and low Si/Al ratio, monomeric **Z**₂Zn²⁺ species are expected. On the other hand, at higher Zn loadings and

intermediate Si/Al ratio, two $Z[Zn(OH)]$ species (with each Zn cation interacting with one Al site) might be able to form oxygen bridged Zn-O-Zn species.

Not surprisingly, the nature of the Zn sites strongly depends on the distribution of Al sites in the zeolite and hence on the exact preparation method [28, 35, 36]. Among the different proposed methods to incorporate Zn in the parent zeolite, ion exchange generally provides higher ratio of active cationic species promoting aromatization and dehydrogenation reactions [28]. It is important to note that the incorporation of Zn ions in the exchange positions effectively reduces the amount of Brønsted acid centers, which are also required in order to form the olefins that can undergo cyclization and aromatization reactions. An optimum elemental composition for high aromatization yields should therefore balance the amounts and nature of Zn species with respect to the Brønsted acid centers in the bifunctional material.

In this contribution, we aim to study exactly this balance between acid site distribution and Zn content and its effect on the aromatization activity of HZSM-5 zeolites. For this purpose, samples with increasing Zn loadings have been produced by ion exchange of two parent HZSM-5 zeolites with different Si/Al ratios. In addition to detailed catalytic testing, the samples were characterized by X-ray diffraction (XRD) and N_2 -sorption in order to confirm their structural integrity and the accessibility of the micropores. Pyridine adsorption followed by transmission Fourier transform infrared (FTIR) spectroscopy [37] was applied to obtain information on the amount and distribution of acid sites and a combination of *in situ* and *operando* X-ray absorption spectroscopy (XAS) techniques [38] were used to characterize the nature of the Zn species formed during activation and MTA reaction. We show that, even though high yields of aromatics can be achieved over pure zeolites with a high density of Brønsted acid sites, the presence of Zn at relatively

high Si/Al ratio allows to force the aromatics formation mainly via the dehydrogenation pathway.

2. Experimental

2.1 Preparation of the catalysts

Two commercial ZSM-5 samples were used as raw materials: CBV-3024E (Si/Al=15) and CBV-8014 (Si/Al=40) obtained from Zeolyst International. Both samples were ion-exchanged, first with a NH_4NO_3 solution (98 %) to have them in their ammonium form and, subsequently with excess of $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ (36.6 g in 1000 mL demineralised- H_2O), one, two or three times, respectively, at 80 °C for one hour in order to produce a set of 8 samples with different acidity and Zn loadings. After each exchange, the samples were filtered and washed with water at 80 °C and dried at 110 °C overnight. In a last step, the materials were calcined at 450 °C for 2 hours. Details on the composition of the samples are collected in Table 1.

Table 1. Experimental details of the Zn incorporation in two parent ZSM-5 zeolites with different Si/Al ratios.

Sample	Parent zeolite	Si/Al	Number of exchanges	Zn content (%)
Si/Al15-P			0	0.0
Si/Al15-1Zn	CBV-3024E	15	1	1.3
Si/Al15-2Zn			2	1.7
Si/Al15-3Zn			3	1.9
Si/Al40-P			0	0.0
Si/Al40-1Zn	CBV-8014	40	1	0.7
Si/Al40-2Zn			2	1.0
Si/Al40-3Zn			3	1.0

Parent samples are named as Si/AlX-P, where “X” indicates the Si/Al ratio of each material, and exchanged materials are named as: Si/AlX-YZn, with “Y” denoting the number of Zn-exchanges.

2.2 Catalysts characterization

2.2.1 Laboratory characterization

Powder X-ray diffraction (XRD) patterns of the catalysts were recorded using a Siemens Bruker D5000 instrument with Bragg-Brentano geometry and CuK α radiation ($\lambda=1.5406$). Samples were measured using flat sample holders loaded with 25 mg of powder. The chemical composition of the samples (Si, Al and Zn content) was analyzed by inductively coupled plasma optical emission spectrometry (ICP-OES) using an Agilent 720 instrument.

Adsorption capacity and textural properties of the catalysts were evaluated by N₂ adsorption-desorption measurements carried out at -196 °C on a Bel Belsorp-mini II apparatus. Prior to the measurements, the samples were degassed under vacuum for 4 hours, at 80 °C the first hour and then at 300 °C. The BET equation and t-plot method were used to calculate the surface area, external area and micropore volume respectively [39, 40]. Crystal sizes and morphology were studied by scanning electron microscopy (SEM), recording the images on a Hitachi Su 8230 FE-SERM working at 5-10 kV with beam acceleration mode at distances around 10 mm.

Fourier-transform infrared spectroscopy (FTIR) experiments were performed in transmittance on a Bruker Vertex 80 instrument with a Mercury Cadmium Telluride (MCT) detector using a resolution of 2 cm⁻¹. Before dosing probe molecules, (CO at -196 °C and pyridine at room temperature) the samples were pretreated under vacuum

(10^{-5} mbar) following a temperature ramp (3 °C/min) till the set point of 450 °C, which was maintained for 1 hour. Pyridine saturated samples were outgassed at 200 °C, to remove excess and physisorbed pyridine. Brønsted and Lewis sites were quantified based on the integrated intensity of the bands at 1544 and 1455 cm^{-1} , respectively using the integrated molar extinction coefficient reported in previous literature [41].

2.2.2 Solid state NMR spectroscopy

Single-pulse ^{27}Al MAS spectra have been acquired at ambient temperature on a Bruker Avance III 950/54 μs^2 spectrometer (22.3 T), employing a Bruker ^1H -X-Y 2.5 mm MAS probe and a spinning speed of $\nu_{\text{R}} = 30.0$ kHz. The single-pulse experiments employed short rf pulses of 0.5 μs for rf field strengths of $\gamma B_1/2\pi = 100$ kHz and relaxation delays of 0.5 s. The ^{27}Al chemical shifts are referenced to a 1.0 M aqueous solution of $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$. A spectrum of an empty rotor recorded under the same experimental conditions has been subtracted from the ^{27}Al MAS NMR spectra since the rotor and probe contribute with broad resonances of very low intensity.

Single-pulse ^{29}Si MAS NMR spectra have been recorded on a Varian INOVA-300 (7.05 T), using a homebuilt CP/MAS probe for 7 mm o.d. PSZ rotors and a spinning speed of $\nu_{\text{R}} = 7.0$ kHz. The single-pulse spectra employed a $\pi/4$ excitation pulse ($\tau_{\text{p}} = 3.0$ μs for $\gamma B_1/2\pi = 40$ kHz), and a relaxation delay of 30 s. The ^{29}Si chemical shifts are referenced to neat tetramethylsilane (TMS) using a sample of $\beta\text{-Ca}_2\text{SiO}_4$ ($\delta = -71.33$ ppm) as a secondary reference. For both the ^{27}Al and ^{29}Si NMR studies, all samples have been stored at ambient conditions and thus, they are present in a hydrated state.

^1H MAS NMR spectra have been acquired on a Varian Direct Drive VNMR-600 spectrometer (14.1 T), using an Agilent FastMAS probe for 1.6 mm o.d. PSZ rotors and a spinning speed of $\nu_{\text{R}} = 40.0$ kHz. The experiments employed the DEPTH pulse sequence

[42], to minimize background signals from the rotor and the probe, π and $\pi/2$ excitation pulses for an rf field strength of $\gamma B_1/2\pi = 118$ kHz and a relaxation delay of 7 s. The ^1H chemical shifts are referenced to TMS using solid adamantane ($\delta = 1.74$ ppm) as a secondary reference. For the ^1H NMR studies all samples were dehydrated (vacuum at 10^{-4} bar, 450°C , 20 h) and the rotor was packed in a glovebox under argon atmosphere. Despite the minimization of background resonances by the DEPTH pulse sequence, a spectrum of an empty rotor recorded under the same experimental conditions was subtracted in order to remove some low-intensity resonances for the rotor/probe. In addition, low-intensity signals from residual crystal water (6 – 10 ppm) [43] were removed by subtracting simulated peaks at these chemical shifts for a better illustration of the spectra.

2.2.3 Operando XAS

operando XAS data were collected at the BM23 beamline [44] of the European Synchrotron Radiation Facility (ESRF, Grenoble, France). To keep the Zn-ZSM-5 catalysts under controlled conditions in terms of temperature and gas feed during XAS acquisitions, a setup based on the Microtomo reactor cell [45] designed by the ESRF Sample Environment Group was employed. The gas feed and reaction products formed during the XAS measurements were qualitatively monitored by a mass spectrometer (MS) connected to the outlet of the cell.

XAS spectra were initially collected for the as-prepared hydrated catalysts at room temperature (RT) in air. Subsequently, the samples, held inside the Microtomo reactor cell, were pretreated heating in technical air (20 ml/min) from RT to 500°C (heat ramp: $5^\circ\text{C}/\text{min}$) and kept at 500°C for 30 min. The temperature was subsequently lowered to 420°C , and the XAS spectra of the activated catalysts were collected. Finally, while

keeping the catalysts at 420 °C, the gas feed was switched to methanol (MeOH)/He, by passing a 20 ml He flow through a saturator filled with MeOH at 17 °C. The XAS spectra in MTA conditions were collected after a stable MS signal from toluene was detected.

Zn K-edge XAS data were collected in transmission mode, using a double-crystal Si(111) monochromator for the incident energy scan, two Rh-coated mirrors at 5 mrad angle for harmonics rejection, and ionization chambers for the detection of incident (I_0) and transmitted (I_1) photons. For energy calibration, Zn metal foil was measured simultaneously with all the acquired spectra using a third ionization chamber I_2 [46]. The Zn-ZSM-5 catalysts were measured in the form of a self-supporting pellets, fixed inside the reactor cell. The optimized pellet weights were in the 150–170 mg range for 1.3 cm² area pellets, resulting in edge jumps $\Delta\mu_x$ in the 0.2–0.4 range (depending on Zn-loading) for a total absorption after the edge of $\mu_x = 2.55$.

For each catalyst state (hydrated, pretreated, and during MTA reaction), we acquired a series of consecutive XAS scans of ~ 30 min each, collected from 9470 to 10635 eV (up to ~ 16 Å⁻¹) with k-space sampling $\Delta k = 0.035$ Å⁻¹ and acquisition time in the EXAFS region quadratically increasing with k from 1 to 4 s/point to improve the signal-to-noise ratio at higher k-values. The final XAS spectra were obtained by averaging the $\mu(E)$ curves corresponding to the consecutive scans (two or four for high-loading and low-loading samples, respectively), after checking reproducibility among the acquisitions [47]. A faster XANES acquisition mode (ca. 6 min/scan) collected in a shorter energy range (9470-10407 eV) was employed to check the evolution of the XAS features during heating/cooling ramps and stabilization at steady conditions.

Three relevant reference samples; a 50 mM aqueous solution of $\text{Zn}(\text{NO}_3)_2$; a Zn metal foil; and a ZnO powder were also measured at RT in air, using the same acquisition mode as described above.

XAS spectra of the Zn-ZSM-5 and the reference materials were aligned in energy employing the corresponding Zn metal foil spectra detected by the I_2 ionization chamber and normalized to unity edge jump using the Athena software from the Demeter package [48]. The $\chi(k)$ functions extracted using the Athena program and R-space EXAFS spectra were obtained by calculating the Fourier transform of the $k^2\chi(k)$ functions in the (2.5 – 12.5) \AA^{-1} k-range.

2.2.4 DFT-assisted EXAFS fitting

DFT-assisted EXAFS fitting is an approach that has shown its potential for cases of a complex local environments around the adsorbing atom [49, 50]. The structure of $\text{Z}[\text{Zn}(\text{OH})]$ species used as starting point for EXAFS fitting was obtained by means of DFT geometry optimization. As a starting point, a cluster containing a complex of two 10Rs with the corresponding eight 5Rs was cut out from the bulk structure of ZSM-5 taken from IZA-SC Database of Zeolite Structures [51]. To eliminate dangling bonds, outer oxygen atoms were terminated by hydrogens.

Geometry optimization was carried out with the ADF2014 software [52]. Slater-type basis sets were employed for all atoms: TZ2P for the Zn(OH) group, TZP for Si, Al and O atoms, and DZP for terminal OH groups. To speed up the calculations, the frozen core approximation (up to 2p frozen for Zn, Si and Al, 1s frozen for O) was used [53]. In order to account for exchange correlation effects, the dispersion corrected PBE-D3-BJ functional was chosen [54]. Relativistic effects were dealt with by means of the scalar

ZORA approximation [55]. Convergence criteria were set to 0.001 Hartree, 0.001 Hartree/Å and 0.01 Å for energy, gradients and bond distances respectively.

Geometry optimization was performed in three steps, the first two involving a purely siliceous cluster. First, only terminal hydrogens were relaxed, while Si and O atoms were kept fixed. Then, the hydrogens together with the adjacent O atoms were frozen, while the rest of the cluster was optimized. Finally, one Si atom at the T12 position [56] was substituted by Al, and the ZnOH group was placed in its vicinity above the 5R. The optimization was performed keeping the terminal OH groups fixed, while all the other atoms could move without any symmetry restrictions.

EXAFS fitting was performed in R-space in the $\Delta R = 1.0 - 3.3$ Å range, on the FT of the k^2 -weighed $\chi(k)$ EXAFS spectra performed in the $2.5 - 11.0$ Å⁻¹ range, resulting in 17 independent points ($2\Delta k\Delta R/\pi > 17$). Phases and amplitudes have been calculated by FEFF6 code [57, 58] using the Arthemis software from the Demeter package [48], employing the DFT-optimized structure of **Z**[Zn(OH)] as input.

The fitting model included all the single-scattering (SS) paths present in the R-range of interest, representing the dominant contribution to the experimental EXAFS signal. To limit the number of optimized variables, all the SS paths have been optimized with the same passive amplitude reduction factor (S_0^2) and with the same energy shift parameter (ΔE). In the EXAFS refinement, coordination numbers N_i for each of the considered coordination shells were fixed to the DFT values. In the first shell region, the adopted fitting model included the SS path including an extra-framework oxygen atom from the OH group (O_{ef} , $N_{O(ef)} = 1$, parametrized with independent radial shift $\Delta R_{O(ef)}$ and Debye-Waller parameter $\sigma_{O(ef)}^2$) and two subshells of framework oxygen atoms (O_{fw} , $N_{O(fw)} = 1$

and O'_{fw} , $N_{O'(fw)} = 2$, parametrized each with an independent radial shift parameter, namely $\Delta R_{O'(fw)}$ and $\Delta R_{O'(fw)}$, but with a common DW $\sigma^2_{O'(fw)}$. In the second shell region, the fitting model includes the SS paths involving the nearest Al neighbor (at 2.97 Å from Zn according to DFT), parameterized with an independent radial shift (ΔR_T) and DW factor (σ^2_T). Finally, a lower-level parametrization strategy has been adopted for the SS paths which involve the farthestmost Si and O atoms of the ZSM-5 framework (3.1 –3.3 Å distance range from the Zn center), providing a minor but not negligible contribution to EXAFS signal. These SS paths were modelled considering a common contraction/expansion factor α_{fw} and DW factor σ^2_{fw} increasing as the square root of the distance $R_{eff,i}$ of the i^{th} scattering atom from the absorber ($\Delta R_{fw,i} = \alpha_{fw} R_{eff,i}$, $\sigma^2_{fw,i} = SS_{fw} (R_{eff,i}/R_0)^{1/2}$, where $R_0 = 3.14$ Å denotes the DFT distance of the shortest SS path of the group). This approach is effective in considering a larger number of scattering paths, while limiting both the number of optimized parameters and the correlation among them, as already demonstrated in several cases [59, 60].

2.3 Catalytic testing

The catalytic tests were performed with 50 mg of catalyst (sieve fraction 250-420 µm) in a fixed bed reactor operated at atmospheric pressure and at 400 °C, using a straight tubular quartz reactor with 4mm inner diameter. A temperature couple enclosed within a quartz sheath was located inside the reactor, in contact with the catalyst bed to monitor the reaction temperature. Prior to testing, all samples were pretreated in situ at 550 °C under oxygen flow for an hour to remove traces of adsorbed hydrocarbons and water. Methanol was fed by flowing He (15 ml/min) through a saturator at 20 °C to obtain a partial pressure of methanol of 130 mbar, which results in a weight hourly space velocity (WHSV) of 3

h^{-1} . The reactor effluent stream was analyzed on-line by gas chromatography using an Agilent 6890 A instrument equipped with a flame ionization detector (FID) and a Supelco SPB-5 capillary column with dimensions of 60 m x 0.53 mm x 3 μm .

3. Results

3.1 Physico-chemical characterization: composition, morphology, porosity

The X-ray diffraction confirms that both the parent and exchanged samples exhibit only the characteristic diffraction peaks of **MFI** topology (Fig. S1). No substantial loss in the intensity of the diffraction peaks was detected in the series of samples with Si/Al=40 after the ion exchange. However, a slight decrease is observed in the exchanged samples produced from the ZSM-5 with higher density of acid sites (Si/Al=15) when Zn is incorporated.

The amount of Zn in the ZSM-5 zeolites and the overall chemical composition of the calcined samples were determined by ICP-OES analysis (Table 2). We observe that the incorporation of Zn is favoured in the parent zeolite with Si/Al=15, which has a higher density of Brønsted acid sites. In addition, for this zeolite, the incorporation of Zn increases progressively with further ion exchange treatments, whereas the Zn content remains rather constant in the Si/Al=40 series. This observation is in good agreement with the previous literature describing that it is more likely to stabilize higher amounts of metal in exchange positions in zeolites with higher density of acid sites because of the higher probability of having two Al sites close enough to stabilize isolated Zn^{2+} ions [34]. However, considering the lower Al content in the Si/Al=40 material, the Zn/Al ratio is higher in that series of samples.

Table 2. Elemental composition of the samples measured by ICP and density of acid sites detected from the FTIR experiments using pyridine as probe molecule.

Sample	Elemental composition (mmol/g)				Density of acid sites (mmol/g)		
	Al	Si	Zn	Zn/Al	Brønsted (B)	Lewis (L)	B/L
Si/Al15-P	0.83	13.64	0.00	0.00	0.91	0.34	2.68
Si/Al15-1Zn	0.95	15.28	0.19	0.20	0.27	1.10	0.25
Si/Al15-2Zn	0.94	15.28	0.26	0.28	0.18	1.12	0.16
Si/Al15-3Zn	0.92	15.13	0.29	0.32	0.20	1.50	0.13
Si/Al40-P	0.35	13.96	0.00	0.00	0.31	0.12	2.58
Si/Al40-1Zn	0.46	18.52	0.11	0.24	0.24	0.45	0.53
Si/Al40-2Zn	0.37	15.75	0.15	0.41	0.18	0.62	0.29
Si/Al40-3Zn	0.36	15.52	0.16	0.44	0.17	0.70	0.24

Scanning electron microscopy (SEM) reveals significant differences in crystal size and morphology for the two parent zeolites (Fig. S2). The materials with Si/Al=40 are formed by agglomerated small crystals (agglomerate sizes between 1 to 12 microns), whereas smaller agglomerates (0.3-3 microns) are observed in the samples with Si/Al=15. This might influence the diffusion capacity of the samples and should be kept in mind when analysing the adsorption capacity of the different materials. No significant variations in either agglomeration and crystal size or shape were seen for the exchanged samples.

The presence of metal ions in the channels of the zeolites could cause some spatial restrictions. To identify if the presence of Zn in the micropores of the zeolites is altering the adsorption capacity or the textural properties of the samples, all the parent and exchanged materials were studied by N₂ physisorption measurements. Surface areas and pore volumes are collected in Table 3. These characteristics remain constant with the incorporation of Zn. Thus, there is no indication of diffusion restrictions for the zeolitic materials due to the presence of Zn ions (Table 3). All the isotherms are close to type I (according to the IUPAC classification), which is associated to microporous materials

[39]. However, small non-microporous contributions are also revealed, as shown by the increasing nitrogen uptake at relatively high partial pressures (Fig. S3). This contribution is higher in all the samples generated from the parent with Si/Al=15, and likely due to filling of interparticle spaces formed in the agglomerates which are in the meso-/macropore range (Table 3).

Table 3. Adsorption capacity and textural properties of the parent and Zn-exchanged ZSM-5 samples*.

Sample	Surface area (m ² /g)			Pore volume (cm ³ /g)		
	BET	NonMicrop	Microp	Total	NonMicrop	Microp
Si/Al15-P	439	34	405	0.37	0.19	0.18
Si/Al15-1Zn	403	36	367	0.36	0.20	0.16
Si/Al15-2Zn	401	36	365	0.37	0.21	0.16
Si/Al15-3Zn	409	37	372	0.37	0.20	0.17
Si/Al40-P	434	29	405	0.33	0.14	0.19
Si/Al40-1Zn	445	21	424	0.29	0.08	0.21
Si/Al40-2Zn	449	17	432	0.29	0.08	0.21
Si/Al40-3Zn	415	16	399	0.27	0.08	0.19

* The standard error of the calculated surface areas and pore volumes of the materials are estimated to be 4 and 5% respectively.

3.2 Identification of Zn species in Zn-ZSM-5

3.2.1 FTIR analysis of acidic properties

To investigate how the incorporation of Zn in the zeolite alters the distribution and nature of the reactive sites in the material, samples were characterized by FTIR using pyridine and CO as probe molecules.

Fig. 1.a shows the FTIR spectra of the activated samples in the range of the O-H stretching vibrations before pyridine absorption. The band at 3610 cm⁻¹ is assigned to Brønsted acid sites [Si-(OH)-Al], while the band at 3745 cm⁻¹ is related to isolated silanols [Si-OH], typically located at the external surface. The shoulder towards lower frequencies (~3728

cm^{-1}) has been assigned to weakly interacting internal silanols. This contribution becomes slightly more important in samples with Zn of the Si/Al=40 series. Finally, for all samples of the Si/Al=15 series an additional weak band located at $\sim 3664 \text{ cm}^{-1}$ is observable, which is assigned to extra-framework aluminium (EFAI) species [61, 62]. This band is also more intense in samples with Zn.

After chemisorption of pyridine, the corresponding ring vibrational modes of the molecule appear in the spectra (Fig. 1b). All the Brønsted sites in the materials are accessible to pyridine, as the band associated to the OH stretching vibrations of Brønsted sites (3610 cm^{-1}) is completely extinguished after pyridine is dosed (Fig. S4). The total amount of adsorbed pyridine molecules on Brønsted and Lewis acid sites is calculated from the integral areas of the two characteristic bands (1544 and 1455 cm^{-1} respectively), and the values are collected in Table 2.

We observe that the incorporation of Zn in the parent zeolites causes an increase in the density of Lewis acid sites together with a decrease in the density of Brønsted sites. This is expected since the incorporation of Zn in a zeolite framework replaces the charge balancing protons, effectively reducing the concentration of Brønsted acid sites and giving rise to new Lewis sites [28]. Importantly, from the calculated densities of Brønsted and Lewis sites it is clear that more Lewis sites are formed than the number of Brønsted sites consumed. This point will be elaborated in the discussion.

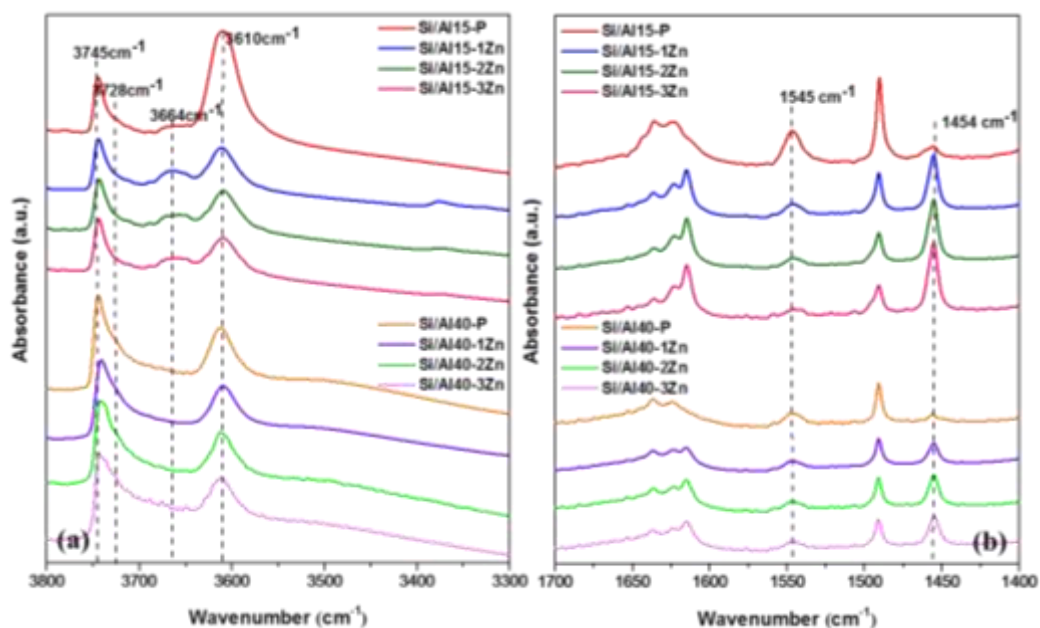


Fig. 1. FTIR spectra of the parent H-ZSM-5 and the Zn-containing HZSM-5 zeolites. **a)** FTIR spectra of the pre-treated samples in the OH stretching region. **b)** FTIR spectra of pyridine adsorbed on the zeolites.

Further insights about this result can be obtained by using CO as a probe molecule. Fig. 2 reports background subtracted spectra at low CO coverage. In both parent zeolites, the typical band related to CO interacting with [Si-(OH)-Al] at 2174 cm⁻¹ are present together with weak components at 2230 and 2214 cm⁻¹, assigned to CO on EFAl sites with different Lewis character [63]. After Zn insertion, bands in the Lewis sites region (2250 – 2200 cm⁻¹) grow substantially, and at least one new component is observed at 2224 cm⁻¹. This is more evident on the samples with Si/Al =15, in agreement with the almost double Lewis site concentration measured by pyridine adsorption (Table 2). The new band can be safely assigned to Zn²⁺ ions, in agreement with the reports by Kazansky et al. [64]. These results indicate that pyridine probes both Zn²⁺ and EFAl sites.

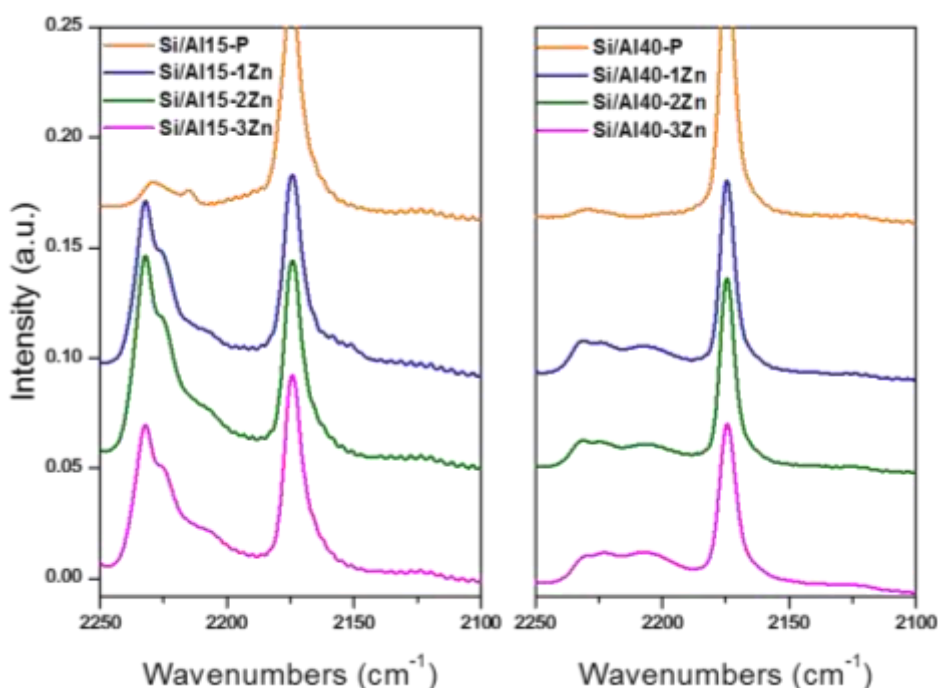


Fig. 2. Background-subtracted FTIR spectra of CO (0.1 mbar) adsorbed at -196 °C on parent H-ZSM-5 and Zn-containing HZSM-5 zeolites.

3.2.2 Solid state NMR spectroscopy

The single-pulse ^{29}Si MAS NMR spectra for the Si/Al=15 and Si/Al=40, shown in Fig. S5, allow distinction of SiO_4 environments of the type $\text{Q}^m[\text{nAl}]$ (m : number of bridging oxygens; n : number of connected Al tetrahedra). $\text{Q}^4[0\text{Al}]$ resonances in the range -110 ppm to -117 ppm dominate the spectra. Si bonded to one Al tetrahedron ($\text{Q}^4[1\text{Al}]$) is observed at -106.5 ppm and silanol groups ($\text{Q}^3[0\text{Al}]$) at -102 ppm to -103 ppm. Employing these resonances in deconvolutions of the ^{29}Si MAS NMR spectra and their resulting relative peak intensities (I) allows calculation of the Si/Al ratio of the framework structure (Table 4) by the relationship

$$\frac{\text{Si}}{\text{Al}} (^{29}\text{Si}) = \frac{3 \cdot I[\text{Q}^3[0\text{Al}]] + 4 \cdot I[\text{Q}^4[1\text{Al}]] + 4 \cdot I[\text{Q}^4[0\text{Al}]]}{I[\text{Q}^4[1\text{Al}]]} \quad (1)$$

which holds for silicious zeolites, considering the avoidance of Al-O-Al bonds (Loewenstein's rule [65]). A comparison of the ^{29}Si NMR spectra shows only slight differences for samples with Si/Al=40 in the amount of silanol sites. The Si/Al ratio is identical for the four samples of this series and very close to the Si/Al ratio determined by ICP, which shows that almost all Al is present in the framework structure. On contrary, for the Si/Al=15 series, a significant amount of Al is not present in the framework structure, as seen by the higher Si/Al ratios from NMR compared to those determined by ICP. Interestingly, Zn ion-exchanged samples with Si/Al=15 have a lower Si/Al ratio than the parent calcined sample, which suggests that the exchange treatment and subsequent calcination to some extent anneals structural defects and incorporates a part of the extra-framework Al species (EFAl) in the framework again.

Single-pulse ^{27}Al MAS NMR spectra for the two series are shown in Fig. 3 where central-transition resonances from four-, five- and six-fold coordinated Al are identified at chemical shifts of approx. 55, 30 and 0 ppm. The dominating signal at 55 ppm is related to tetrahedral framework sites whereas the five- and six-fold coordinated Al sites are ascribed to extra-framework aluminum from either cationic aluminum hydroxide species/hydroxylated alumina-like clusters inside the channel structure or as framework defects, where hydroxyl groups and water are partly bonded. The concentrations of the different Al environments have been determined by the relative intensities from simple simulations of the ^{27}Al MAS NMR spectra, using Voigt-type lineshapes, combined with the elemental concentrations from the ICP experiments (Table 4). For the series of samples with Si/Al=40, nearly all Al is present as Al(IV) (>95%) and, incorporated in the framework structure, as supported by the ^{29}Si NMR spectra. Only small amounts of Al are present as Al(VI). However, a significant amount of Al is present as Al(V) and Al(VI)

in the samples with Si/Al=15, especially in the parent calcined sample (17% Al(V) and 8.5% Al(VI)). This indicates that EFAl is incorporated into the framework structure and that defect sites are annealed by the ion exchange treatment, since the amounts of Al(V) and Al(VI) are reduced to below 3% and 4%, respectively. This observation is also in agreement with the indications from ^{29}Si NMR. Moreover, the high-field ^{27}Al NMR spectra reveal that a low-intensity shoulder is observed at the low-frequency side of the Al(IV) center bands for the samples with Si/Al=15, especially for the parent calcined sample. This site is tentatively ascribed to a distorted framework Al site in the vicinity of either a framework defect or a cationic EFAl species, which has replaced a Brønsted acid site [66, 67].

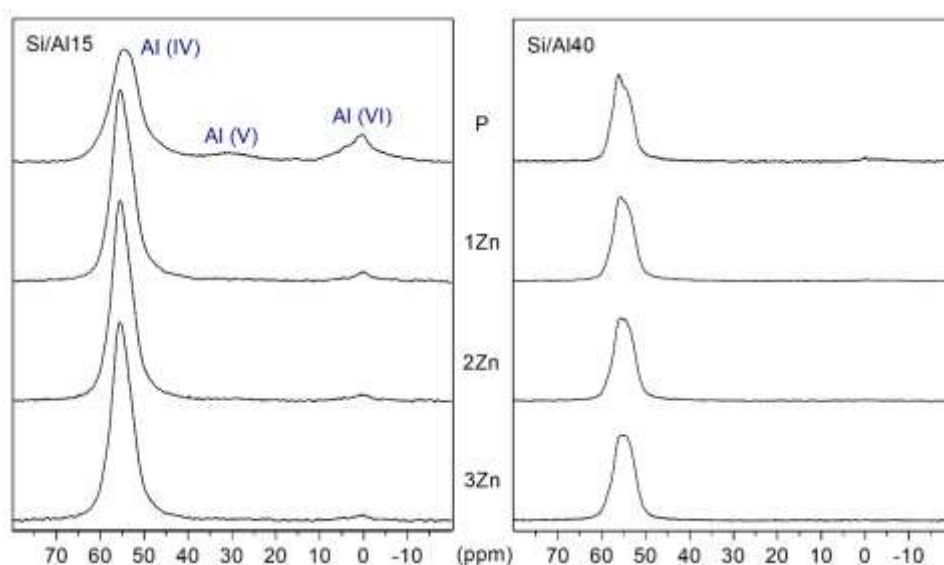


Fig. 3. Single-pulse ^{27}Al MAS NMR spectra of the two sample series Si/Al=15 and Si/Al=40 ($B_0 = 22.3$ T, $\nu_R = 30.0$ kHz). Four-, five- and six-fold surrounded ^{27}Al environments at chemical shifts of around 55, 30 and 0 ppm, respectively, are indicated in the spectrum from Si/Al15-P.

^1H DEPTH MAS NMR spectra of the two sample series are illustrated in Fig. 4 and reveal different types of proton species. Protons from Brønsted acid sites (BAS) have a resonance at about 3.5 to 5 ppm and signals of terminal and internal silanol sites appear

between 1.5 and 2.0 ppm. [68-71]. The resonances from Al-(OH)_x hydroxyl groups bonded to EFAl sites and/or framework defects and involved in hydrogen bonding are observed at approx. 2.5 ppm, whereas less distorted Al-(OH)_x units resonate at 0.8 ppm. Furthermore, it has recently been reported that Zn-OH sites in Zn-impregnated zeolites have chemical shifts at approx. 0.6 ppm [72, 73]. The concentration of BAS determined from simulations of the ¹H DEPTH MAS NMR spectra show an overall good agreement with the concentration of BAS determined by IR spectroscopy. The only exception is sample Si/Al15-P, which shows the double amount of BAS in the IR experiments. For both series it is observed that the Zn ions replace the BAS gradually with further ion exchange processes until approximately 0.2 mmol/g BAS remain. The Zn ions may be present as Zn²⁺, Zn(OH)⁺ or Zn(OH)₂. The existence of the latter two types of Zn environments is supported by the signal at 0.6 ppm, which increases slightly with further ion exchange processes [72, 73]. However, the presence of ¹H resonances from Al-(OH)_x sites with similar chemical shift may interfere with a clear identification of these sites.

The amount of BAS in the samples with Si/Al=15 is much lower than expected for such aluminum content. This may be related to the relatively high amount of EFAl, partially present as cationic EFAl species, which has replaced the protons at the ion-exchange positions. We note that a comparison of the concentrations of Al-(OH)_x and silanol sites from ¹H NMR and concentrations of EFAl and silanol sites from ²⁷Al and ²⁹Si NMR shows a discrepancy, since the concentrations of Al-(OH)_x and silanol sites from ¹H NMR are over- and underestimated, respectively. However, as expected, it is observed that the concentration of Al-(OH)_x sites reflects the amount of EFAl sites contained in the samples.

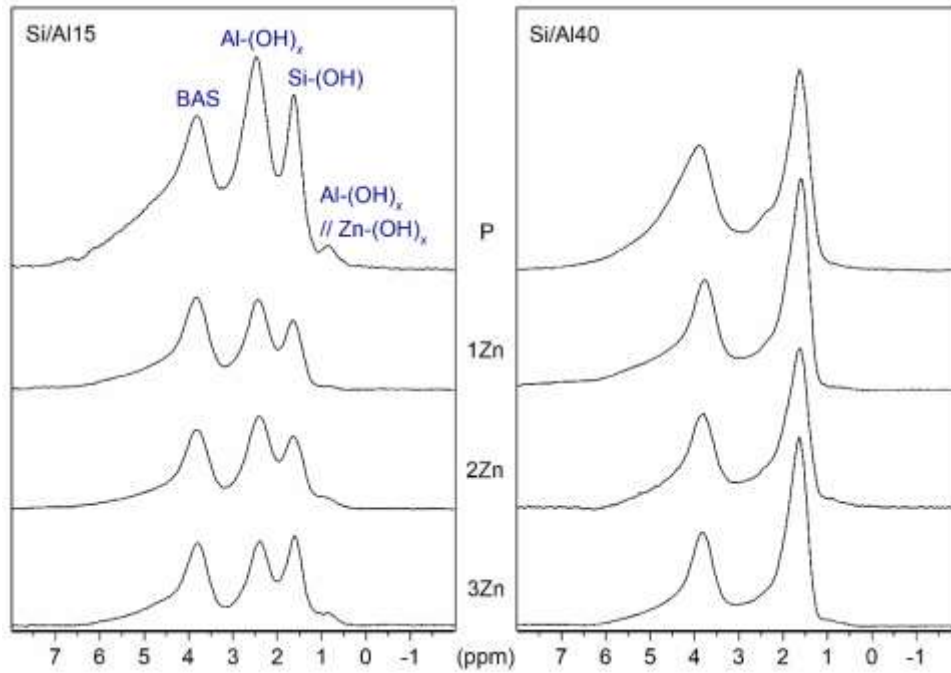


Fig. 4. ^1H DEPTH MAS NMR spectra ($B_0 = 14.1$ T, $\nu_R = 40.0$ kHz) of the two sample series Si/Al=15 (left) and Si/Al=40 (right).

Table 4. Si/Al ratios from ICP-OES and ^{29}Si NMR, and concentrations in mmol per gram of different sites determined by ^{29}Si , ^{27}Al and ^1H NMR.

Sample	Si/Al	Si/Al	silanol	Al(IV)	Al(V)	Al(VI)	EFAI	silanol	Al-(OH) _x	BAS
	ICP ^a	$^{29}\text{Si}^b$	(mmol/g)	(mmol/g)	(mmol/g)	(mmol/g)	(mmol/g)	(mmol/g)	(mmol/g)	(mmol/g)
			$^{29}\text{Si}^c$	$^{27}\text{Al}^c$	$^{27}\text{Al}^c$	$^{27}\text{Al}^c$	$^{27}\text{Al}^d$	$^1\text{H}^e$	$^1\text{H}^e$	$^1\text{H}^e$
Si/Al1	16	25	0.31	0.62	0.07	0.14	0.21	0.13	0.26	0.45
5-P										
Si/Al1		19	0.49	0.76	0.02	0.03	0.05	0.08	0.14	0.19
5-1Zn										
Si/Al1		19	0.43	0.79	0.02	0.02	0.04	0.13	0.22	0.16
5-2Zn										
Si/Al1		18	0.53	0.80	0.01	0.02	0.03	0.14	0.20	0.15
5-3Zn										
Si/Al4	40	42	0.60	0.33	---	0.02	0.02	0.22	0.04	0.33
0-P										

Si/Al4	42	0.86	0.34	---	0.01	0.01	0.28	0.05	0.24
0-1Zn									
Si/Al4	42	0.98	0.34	---	0.01	0.01	0.32	0.07	0.20
0-2Zn									
Si/Al4	42	0.87	0.34	---	0.01	0.01	0.35	0.06	0.18
0-3Zn									

^a Bulk Si/Al ratios; The same values as given for the calcined samples are obtained for the Zn ion-exchanged samples for each series of materials.

^b Si/Al (²⁹Si) ratios of the framework calculated from eq. (1).

^c Densities calculated with the relative intensities from deconvoluted NMR spectra and the elemental concentrations from ICP.

^d Relative intensity for EFAl, where Al(V) and Al(VI) resonances are considered as extra-framework aluminum.

3.2.3 XAS analysis of Zn speciation in the as-prepared and pretreated catalysts and under MTA conditions

To obtain further information on the nature of Zn sites in Zn-ZSM-5 catalysts at key process conditions, selected samples were characterized by Zn K-edge X-ray absorption spectroscopy (XAS). Fig. 5 shows the XANES and FT-EXAFS spectra for a Zn-ZSM-5 catalyst (Si/Al40-3Zn) in its hydrated state (RT in air), after pretreatment in air at 500 °C and cooling down to MTA operation temperature of 420 °C, and finally under MTA conditions. For comparison, the figure also reports the XAS spectra of Zn(NO₃)₂ aqueous solution, representative of fully hydrated Zn²⁺ ions, Zn metal and ZnO powder. These will guide the interpretation of the XAS features observed for the investigated Zn-ZSM-5 catalysts.

In the hydrated state at RT, the Zn species in the zeolite are almost indistinguishable from those observed in solution [74, 75], consisting of mobile [Zn(H₂O)_n]²⁺ or [Zn²⁺(H₂O)_{n-1}(OH)]⁺ complexes with n = 6, 5. The correspondent XANES spectra show an intense

peak at ca. 9668 eV (feature A in Fig. 5a). This feature is commonly referred to as "white line" peak it is and characteristic of Zn^{2+} cations in pseudo-octahedral coordination environments. Also, no trace of aggregated phases in the form of well-defined metal or oxide particles can be found, as it is clear from the comparison with the respective reference compounds.

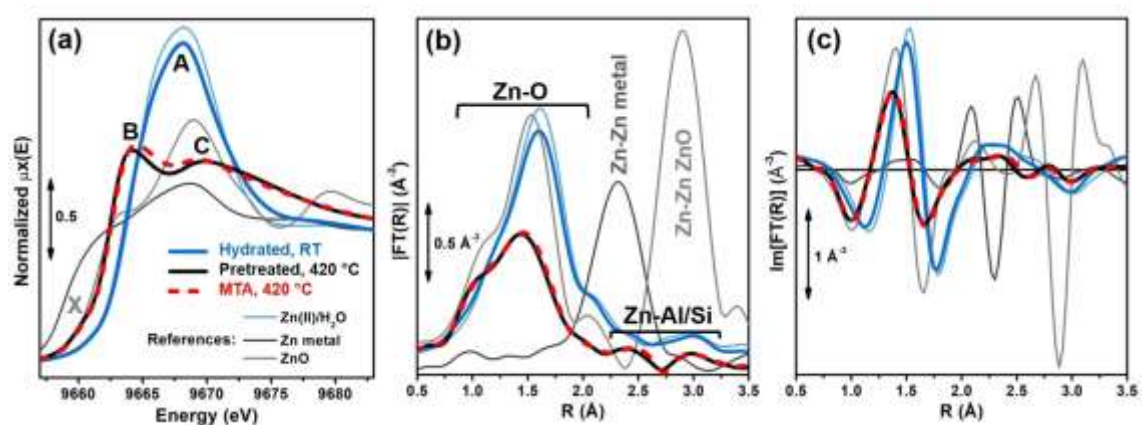


Fig. 5. Typical XAS fingerprints of Zn sites in Zn-ZSM-5 at key process steps. (a) Zn K-edge XANES for a representative Zn-ZSM-5 sample (Si/Al40-3Zn) in the hydrated state at RT, after pretreatment in air at 500 °C and cooling to 420 °C and during MTA at the same temperature. For comparison, the XANES spectra for a series of reference compounds (measured at RT) are also reported, i.e. $\text{Zn}(\text{NO}_3)_2$ aqueous solution, representative of fully hydrated Zn^{2+} ions, Zn metal foil and ZnO. (b, c) Magnitude (b) and imaginary part (c) of the FT-EXAFS spectra for the same sample and reference compounds reported in part (a), using the same colour code. All the FT-EXAFS spectra have been obtained by transforming the $k^2\chi(k)$ EXAFS functions in the 2.5-12.5 \AA^{-1} range.

High-temperature activation in air results in a decreased intensity of the first-shell EXAFS peak, indicating a lower coordination number to O neighbours. Simultaneously, this peak shifts to a significantly shorter distance with respect to the hydrated state, comparable to that observed for the ZnO reference. Also, the absorption edge position, found at 9662.6

eV (first maximum in the derivative spectrum, Fig. S6) in the activated catalyst significantly shifts to lower energies with respect to the hydrated material (9663.8 eV), approaching ZnO (9661.8 eV). However, the overall shape of the XANES spectrum is markedly different with respect to the ZnO reference. This sample has well-defined post-edge peaks at ca. 9664.1 eV (feature **B**) and 9669.8 eV (feature **C**) and no trace of the characteristic ZnO resonance at 9679.7 eV.

The characteristic XANES features of the activated material are in line with previous reports on Zn-zeolites [31, 76, 77] and Zn-enzymes [78-80]. The average coordination geometry for Zn-sites in activated Zn-ZSM-5 is hence expected to be still pseudo-tetrahedral, but significantly distorted. The EXAFS signature of the first-shell is rather broad, suggesting the presence of multiple Zn-O subshells at significantly different distances. In contrast to the hydrated sample, after pre-treatment we observe the presence of two additional well-defined maxima in the FT-EXAFS at ca. 2.4 and 3.0 Å in the phase-uncorrected spectra, which we attribute to Zn-Al/Si scattering single scattering (SS) paths in framework-interacting Zn²⁺ sites (see also the insets of Fig. 6b,c). This assignment is corroborated by EXAFS structural analysis reported below. The presence of aggregated Zn metal and ZnO nanoparticles cannot be reconciled with the overall XANES shape. The same conclusion is reached by the raw FT of the EXAFS spectra. Indeed, although Zn-metal and ZnO bulk compounds exhibit in the modulus of the FT a Zn-Zn and Zn-O-Zn component at ca. 2.3 Å and 3.0 Å, respectively (Fig. 5b) the phases of such components are completely different from that present in the corresponding spectra of the Zn-ZSM-5 catalysts, as clearly appreciable in Fig. 5c. These considerations definitively rule out the presence of any significant contribution from well-defined ZnO or Zn metal

nano-particles. Dealing with an ‘average’ technique, it is difficult, however, to conclude on the presence of very diluted dimeric Zn species.

Under MTA conditions, only slight modifications in the XAS spectra are observed compared to the activated catalysts. Interestingly, the XANES region shows clearer changes in the position and relative intensity of the **B** and **C** features (in Fig. 5a), whereas EXAFS remains almost unchanged. This could be consistent with interaction of H with framework-coordinated Zn-sites during the catalysis. After pretreatment and still during MTA, we observe the formation of a small shoulder in the rising-edge region at ca. 9660 eV (feature **X**). As discussed in more details in the Supporting Information, (Fig. S6) feature X can be tentatively connected with the formation of atomic-like Zn⁰ species [81] or Zn⁺-Zn⁺ dimers [76] in very low abundance (<10% total Zn), although further studies are required to clarify its origin.

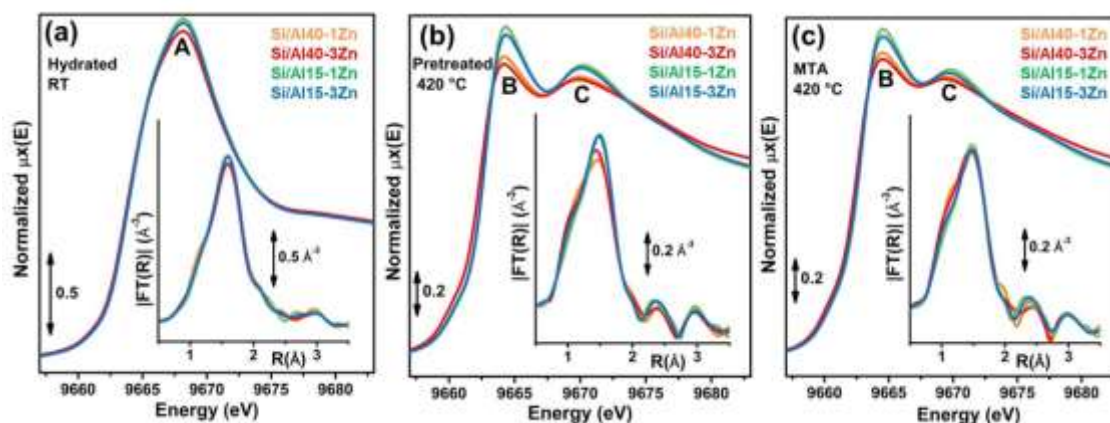


Fig. 6. Zn K-edge XANES (main panels) and FT-EXAFS spectra (insets) for Zn-ZSM-5 catalysts with different composition in term of Si/Al in the parent zeolite and Zn-loading (from 0.7 to 1.9 wt%) collected at key process steps: (a) hydrated catalysts at RT; (b) pretreated catalysts in air at 500 °C and cooled at 420 °C; (c) during MTA at 420 °C. All the FT-EXAFS spectra have been obtained by transforming the $k^2\chi(k)$ functions in the 2.5-12.5 Å⁻¹ range.

Fig. 6 compares the Zn K-edge XAS spectra of selected Zn-ZSM-5 catalysts with different Si/Al ratio of the parent zeolite and Zn-loadings, in their hydrated state, after pretreatment, and in *operando* conditions during MTA. As anticipated, the XAS features observed at the different process steps shows only minor modifications as a function of the composition. In particular, in the hydrated state, all the catalysts show equivalent XANES and EXAFS spectra, hardly distinguishable from the $\text{Zn}^{2+}/\text{H}_2\text{O}$ reference, as discussed above. Nevertheless, small but systematic differences can be appreciated in the XAS spectra of the pretreated catalysts. In particular, the intensity of the **B** and **C** features in the XANES appears to be significantly influenced by the Si/Al ratio: the catalysts with Si/Al = 15 show higher intensity at both peaks with respect to Zn-ZSM-5 samples with Si/Al = 40. A higher Al content in the framework seemingly also induces a small shift of the absorption edge to higher energy. In the FT-EXAFS, Si/Al = 15 samples exhibit a slight enhancement in the intensity of both first- and second-shell peaks with respect to the Si/Al = 40 series, pointing out a small increase in the average coordination numbers of Zn centers as a function of the framework composition. In addition, the first-shell peak sharpens and slightly shifts to higher R-values. The impact of the Zn-loading appears less pronounced: high- and low-loading catalysts give almost equivalent XANES and EXAFS spectra, except for a small enhancement in the intensity of the **B** peak systematically observed in low-loading materials. Under MTA conditions, although the XAS spectra of the samples appears still organized in two groups according to the Si/Al ratio, the differences in the XANES and in the EXAFS region among Si/Al = 15 and Si/Al = 40 are even less pronounced.

Considering the relatively low Al density in all our samples, having two Al sites in close proximity is quite unlikely, so it is possible to envisage $\mathbf{Z}[\text{Zn}(\text{OH})]$ complexes as the

dominant Zn-species in the activated materials. The presence of this kind of Zn sites is supported by the results obtained from ^1H DEPTH MAS NMR. To further support this hypothesis and obtain detailed structural information about the major Zn-species formed in the catalysts after activation, DFT was employed to predict the local geometry of $\text{Z}[\text{Zn}(\text{OH})]$ stabilized in the 5-membered ring (5R) of the ZSM-5 framework structure, previously proposed as the most favourable and accessible location for the Zn cations [82, 83]. The optimized geometry for the $\text{Z}[\text{Zn}(\text{OH})]$ complex is depicted in Fig. 7: the metal centre adopts a highly distorted pseudo-tetrahedral coordination, protruding towards the large 10-membered ring (10R), with the (OH^-) ligand pointing towards the centre of the 10R. As detailed in Fig. 7b, the Zn centre has four O neighbours in the first coordination shell, distributed over a rather broad distance range, from 1.78 Å for the O_{ef} atom of the OH group, to ca. 1.90 and 2.40 Å for the nearer O_{fw} and the two O'_{fw} atoms further away. In the second-shell region, the DFT-model foresees an Al_{fw} at ca. 2.97 Å, while the remaining Si and O atoms of the 5R fall in the 3.14-3.40 Å range from the Zn centre.

We fitted the EXAFS spectrum of the pretreated Si/Al40-3Zn catalyst using the DFT-optimized $\text{Z}[\text{Zn}(\text{OH})]$ geometry as the initial guess for the structural parameters, while fixing the coordination numbers for each coordination shell to the DFT ones (see experimental section for details on the fitting model). Fig. 7c and 7d compare the experimental and best fit curves, together with the principal scattering contributions to the EXAFS signal, while the best fit values for the parameters optimized in the fit are reported in Table 5. Overall, the $\text{Z}[\text{Zn}(\text{OH})]$ model allows a very good reproduction of the experimental spectrum, with physically reliable values for all the optimized parameters and Debye-Waller (DW) factors falling in the typical ranges expected for high-temperature data collection.

Table 5. Results from EXAFS fit for the Si/Al40-3Zn activated catalyst (data collection temperature: 420 °C) using the DFT-optimized Z[Zn(OH)] geometry shown in Fig. 7. The fit was performed in R-space, in the range (1.0 – 3.3) Å, employing the k-range (2.5 – 11.0) Å⁻¹ for the FT (number of independent point $N_{\text{ind}} = 2\Delta k\Delta R/\pi > 17$). Under round brackets, the DFT optimized distances are reported for comparison. Atoms are labelled with the subscript *fw* when are part of the framework and with *ef* in the case of extra framework

EXAFS Parameters	Best fit values
R-factor	0.021
N_{par} (N_{ind})	11 (17)
S_0^2	1.1 ± 0.2
ΔE (eV)	-9 ± 1
$R_{\text{O}(ef)}$ (Å)	1.83 ± 0.04 (1.78)
$\sigma_{\text{O}(ef)}^2$ (Å ²)	0.007 ± 0.003
$R_{\text{O}(fw)}$ (Å)	1.90 ± 0.05 (1.90)
$\langle R_{\text{O}'(fw)} \rangle$ (Å)	2.43 ± 0.02 (2.40)
$\sigma_{\text{O}(fw)}^2$ (Å ²)	0.006 ± 0.003
$R_{\text{T}(fw)}$ (Å)	2.96 ± 0.07 (2.97)
$\sigma_{\text{Al}(fw)}^2$ (Å ²)	0.007 ± 0.005
α_{fw}^a	0.01 ± 0.01
$ss_{fw} (\text{Å}^2)^a$	0.007 ± 0.005

^a*SS paths involving farer Si/O neighbors of the 5R (fw), with distances from the Zn absorber in the 3.14–3.40 Å range are modelled with a common contraction/expansion factor α_{fw} and DW factor σ_{fw}^2 increasing as the square root of the distance $R_{\text{eff},i}$ of the i^{th} atom from the absorber ($\Delta R_{fw,i} = \alpha_{fw} R_{\text{eff},i}$, $\sigma_{fw,i}^2 = ss_{fw} (R_{\text{eff},i}/R_0)^{1/2}$, where R_0 is the shortest R_{eff} for the group of paths). See also experimental section for details.*

These results support the presence of a major population of Z[Zn(OH)] species at Si/Al = 40. The similarity of the FT-EXAFS in activated catalysts with different composition also suggests an important contribution of such Zn²⁺ moieties at Si/Al = 15. In this respect, the higher density of Al in the framework is expected to result in a higher, although still limited, probability to have 5Rs with 2Al sites, or two 1Al sites in neighbouring 5Rs. A correspondingly larger fraction of 2Al Z₂Zn species can thus exist in these catalysts.

Interestingly, a previous DFT-based analysis of such sites by Yakovlev *et al.* [82] evidenced a less distorted environment with respect to what we observe in the $Z[Zn(OH)]$ geometry described above, with the four first-shell Zn-O_{fw} distances in the 1.93–2.04 Å range. A larger fraction of Z_2Zn sites in such configuration would be in line with the slight intensity increase and sharpening observed in the first-shell peak of the FT-EXAFS spectra of activated Zn-ZSM-5 catalysts with Si/Al = 15 (Fig. 6b).

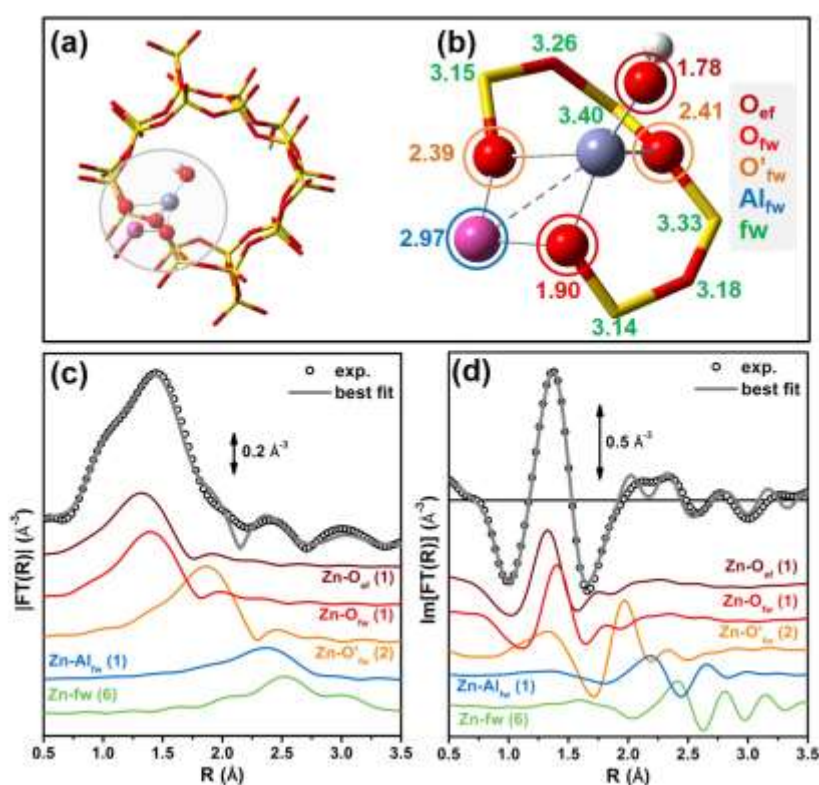


Fig. 7. (a) DFT-optimized geometry for the $Z[Zn(OH)]$ complex in the 5R of the ZSM-5 framework. (b) view of the full optimized cluster with the $Z[Zn(OH)]$ complex highlighted by a grey circle; magnified view of the Zn local environment in the 5R, with coloured circles indicating the different coordination shells included in the EXAFS fit and DFT distances of each atom from the Zn absorber expressed in Å. Atom colour code as follows: Zn, purple; H, white; O, red; Al, pink; Si, yellow. (c,d) Comparison between modulus (c) and imaginary part (d) of the experimental FT-EXAFS spectra (black circles) and corresponding best fit (solid grey lines) for Si/Al40-3Zn pretreated catalyst (data collection temperature: 420 °C) using the DFT-optimized $Z[Zn(OH)]$ geometry shown in Fig. 7. The principal contributions to the EXAFS signal are also

shown, vertically translated for clarity; the values of the parameters optimized in the fits are reported in Table 5. Atoms are labelled with the subscript *fw* when are part of the framework and with *ef* in the case of extra framework

3.3 Catalytic testing results

To evaluate the promoting effect of Zn as dehydrogenating agent and the synergy between Zn and Brønsted acid sites, all the described materials were tested as catalysts in the methanol to aromatics transformation at 400 °C at the conditions described in the experimental section. Under such conditions, all the samples reached full initial conversion of oxygenates (methanol and dimethyl ether) and the shape of the deactivation curves of the exchanged zeolites was similar to that of parent H-ZSM-5 (Fig. 8), with an initial period of complete conversion after which a rapid deactivation occurs leading to conversion values close to 0%. Nonetheless, a clear correlation between lifetime and sample composition in terms of both Al and Zn content was observed. First of all, comparing samples without Zn, H-ZSM-5 with higher density of acid sites (Si/Al15-P) showed a shorter lifetime than the parent zeolite with Si/Al=40. On the other hand, for the samples prepared with the same parent zeolite and increasing Zn content, it is clear that the stability decreases with increasing Zn content for both series of samples (Fig. 8). The loss in stability caused by the incorporation of Zn or other transition metals have been observed before [30, 84] and, in our case since no other remarkable changes are observed in the physicochemical properties, can be initially attributed to the variations in the acidity of the materials. For the correct interpretation of this observation it should be taken into account that both the Brønsted and Lewis acid sites play important roles in the methanol transformation and catalyst deactivation.

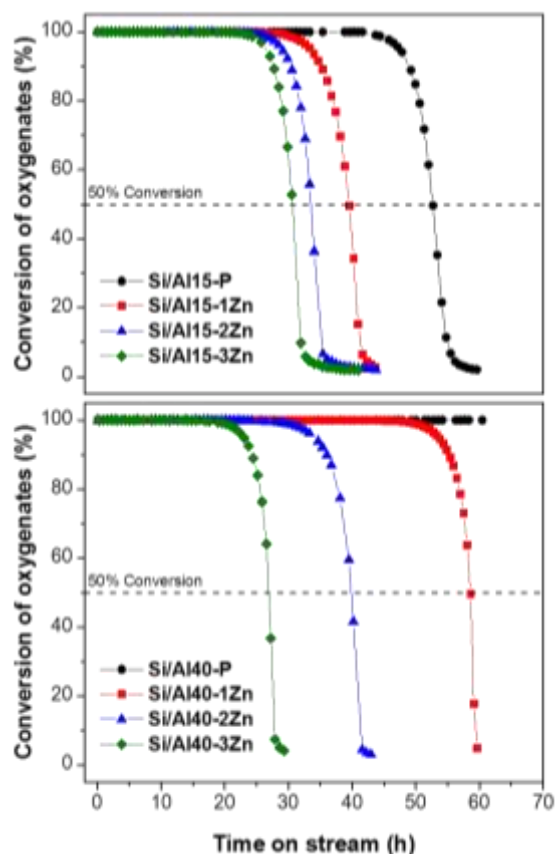


Fig. 8. Conversion of oxygenates (DME+MeOH) at 400 °C and WHSV = 3 h⁻¹ vs time on stream over the parent and exchanged zeolites with the two selected Si/Al ratios.

The resistance of the catalyst to deactivation can be quantified by calculating and comparing cumulative conversion capacities and the turnover number of each catalyst, described as the total weight of methanol converted per gram of catalysts or per accessible active center, respectively, until complete deactivation. Janssens has described deactivation as a loss in active catalyst mass with time, which may be perceived as a decrease in effective contact time [85]. Deactivation coefficients (α) and reaction rate constants (k), as described in Janssens' model, have been calculated for our samples and are collected, together with turnover numbers (TON) and conversion capacities, in Table 6. As expected from the time to the methanol breakthrough for the different catalysts in

the deactivation curves, the deactivation coefficient is within each catalyst series higher for samples with higher Zn content. The deactivation coefficient is also higher for the Si/Al=15 parent sample relative to the Si/Al=40 parent sample. These observations are also manifested by the calculated values of the cumulative conversion capacity, which are higher for both parent samples and far higher in the case of the pure zeolite with Si/Al=40. Comparing the turnover number of the parent samples, Si/Al40-P has a clearly higher TON than Si/Al15-P, which means that each Brønsted site is converting more grams of methanol in the former. It is surprising, though, that the turnover number of the exchanged zeolites decreases for high Zn contents in the Si/Al=40 materials, whereas it is quite independent on the Zn content in the materials with Si/Al=15. Clearly, for the Zn containing catalysts, the Brønsted acid site density alone cannot describe the catalyst performance. This is also reflected by the reaction rate constants (k). When these are normalized to Brønsted acid site density (Table 6), a quite profound increase is seen for the Zn containing catalysts in both series. A straightforward rationalization is that for the Zn containing catalysts, the Lewis acid sites introduced with the Zn contribute strongly to the catalytic activity, as one would also expect, based on the pronounced influence on both lifetime and selectivity. Beyond this, the relatively simple 2-parameter model of Janssens does not provide further, obvious insights.

Table 6. Cumulative conversion capacity, turnover number and calculated values of the deactivation coefficient and the constant rate of the zeolitic materials in MTA transformation at the experimental conditions tested.

Sample	Conversion capacity* (g _{CH₃OH} /g _{cat})	Deactivation coefficient «a» (mg _{cat} /mol _{CH₃OH})**	k (mol _{CH₃OH} /mmol _{BASH}) **	Turnover number (g _{CH₃OH} /mmol _{BAS})***
Si/Al15-P	154	197	2.80	169
Si/Al15-1Zn	116	263	6.43	429
Si/Al15-2Zn	100	310	10.67	556
Si/Al15-3Zn	91	341	10.29	455

Si/Al140-P	484	64	6.41	1564
Si/Al140-1Zn	173	177	13.57	721
Si/Al140-2Zn	116	261	10.47	644
Si/Al140-3Zn	79	388	12.25	464

*Cumulative conversion capacity calculated as the total grams of methanol converted per gram of catalyst from 100 to 0% conversion.

** Deactivation coefficient (a) and rate constant (K) were calculated according to Janssens' model [85]. Rates are calculated per mmol of Brønsted acid site (BAS)

*** Turnover number considered as the total grams of methanol converted per mmol of Brønsted acid sites until full deactivation.

The product distribution was found to be influenced by the composition of the samples, both by the density of acid sites of the parent zeolite and by the presence of Zn. Average selectivities to aromatics, collected in Table 7, were calculated at conversion levels between 100 and 50% in order to obtain a reliable measure not influenced by initial fluctuations in selectivities. Clearly, the presence of Zn in the catalysts has a promoting effect on the aromatics production (Table 7, Fig. 9). However, it must also be considered that when comparing parent samples, Si/Al15-P sample produces a considerably higher yield of aromatics, so, on a relative scale, the promoting effect of Zn is lower for the high Al sample (Table 7). For the parent samples, the differences found in yield of aromatics may be initially explained in terms of the greater density of available Brønsted and zeolite Lewis sites in SI/Al15-P sample which catalyze hydrogen transfer reactions leading to the formation of a higher amount of hydrogen rich molecules (such as alkanes). The goal of the incorporation of Zn is to avoid hydrogen transfer reactions and enhance dehydrogenation, thus producing aromatics together with molecular hydrogen instead of the mentioned alkanes. To evaluate to which extent aromatization is taking place via the dehydrogenation or hydrogen transfer mechanisms, the hydrogen transfer index (HTI) for

C4 hydrocarbons was calculated (Fig. 9) [23]. A low HTI would mean that the aromatization is taking place mostly by dehydrogenation. As shown in the plot in Fig. 9, the higher the Zn/Al ratio, the lower the HTI, despite the increasing yield of aromatics. Finally, the HTI is clearly higher in the Si/Al15 series of samples compared to samples with Si/Al=40.

Table 7. Average selectivity to aromatics calculated between 100 and 50% conversion for all the catalysts studied in MTA transformation at the experimental conditions described.

Sample	Benzene	Toluene	m/p-Xylene	o-Xylene	Tri-methylbenzenes	Total aromatics
Si/Al15-P	1	4	9	2	5	22
Si/Al15-1Zn	1	6	11	2	8	29
Si/Al15-2Zn	1	6	12	3	8	28
Si/Al15-3Zn	1	5	12	3	8	30
Si/Al40-P	0	2	4	1	4	11
Si/Al40-1Zn	1	4	9	2	9	27
Si/Al40-2Zn	1	4	11	2	10	28
Si/Al40-3Zn	0	4	11	2	12	31

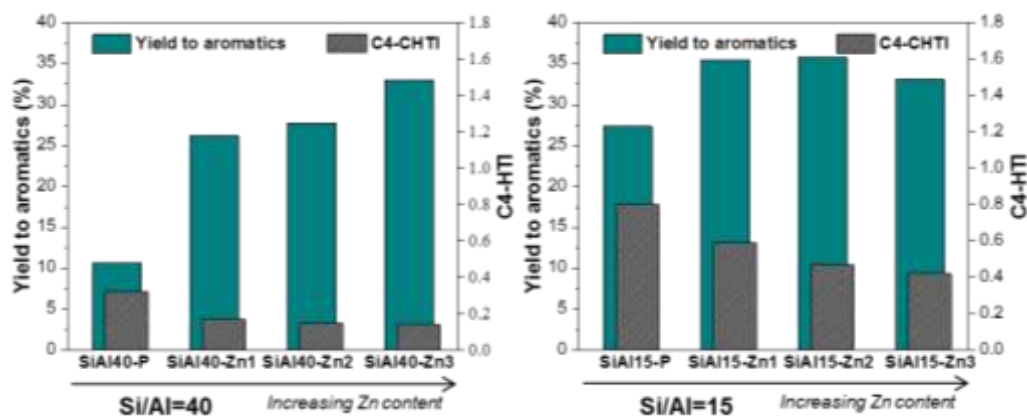


Fig. 9. Yield of aromatics and hydrogen transfer index after 4h on stream over parent and exchanged samples.

Fig. 10 shows the yield of aromatics as a function of time on stream. For all the catalyst with Si/Al=15, the yield of aromatics starts at its maximum and decreases with time on

stream. However, catalysts with Si/Al=40 have a rather different behavior, showing a gradual increase in the yield of aromatics with time on stream up to the methanol breakthrough.

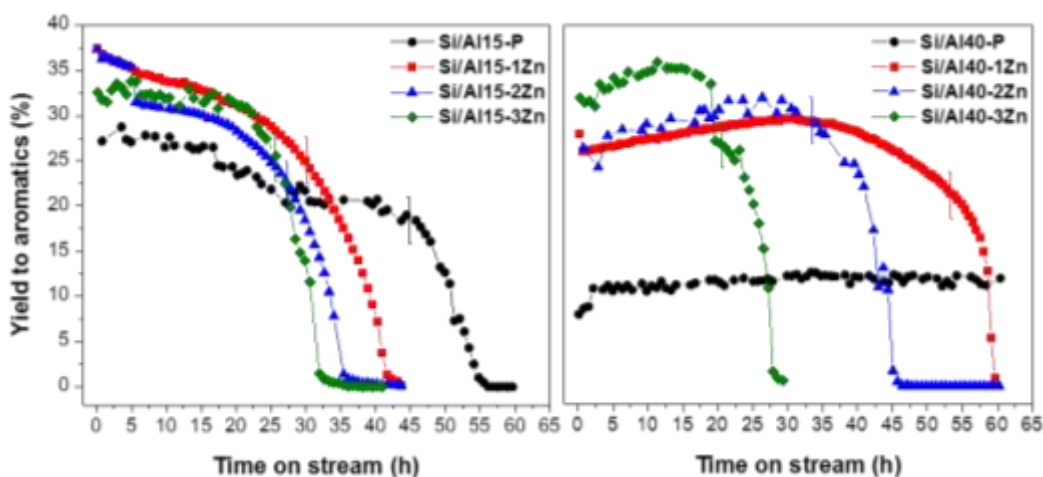


Fig. 10. Yield of aromatics as a function of time on stream using HZSM-5 and Zn-exchanged HZSM-5 materials as catalysts. Methanol breakthrough is marked with vertical lines for each catalyst.

To evaluate in detail how the product distribution evolves, the time evolution for the main reaction products (C_3 , C_4 and C_{6+} divided in aliphatic and aromatic compounds) is presented in Fig. 11. The product distribution obtained is comparable with that reported previously for ZSM-5 catalysts in methanol to hydrocarbon conversion [86], and Si/Al40-P sample is the only exception for which hydrocarbons with over 6 carbon atoms are not the main reaction products, and the catalysts produces predominantly C_3 and C_4 hydrocarbons instead. This can be explained considering that this material does not have a high density of active sites available for aromatization. Samples with Si/Al=15 show a gradual decrease in the yield of C_{6+} hydrocarbons with time on stream, and, for the Zn-

exchanged samples, the production of C₃ molecules becomes gradually higher up to the time where methanol breakthrough occurs.

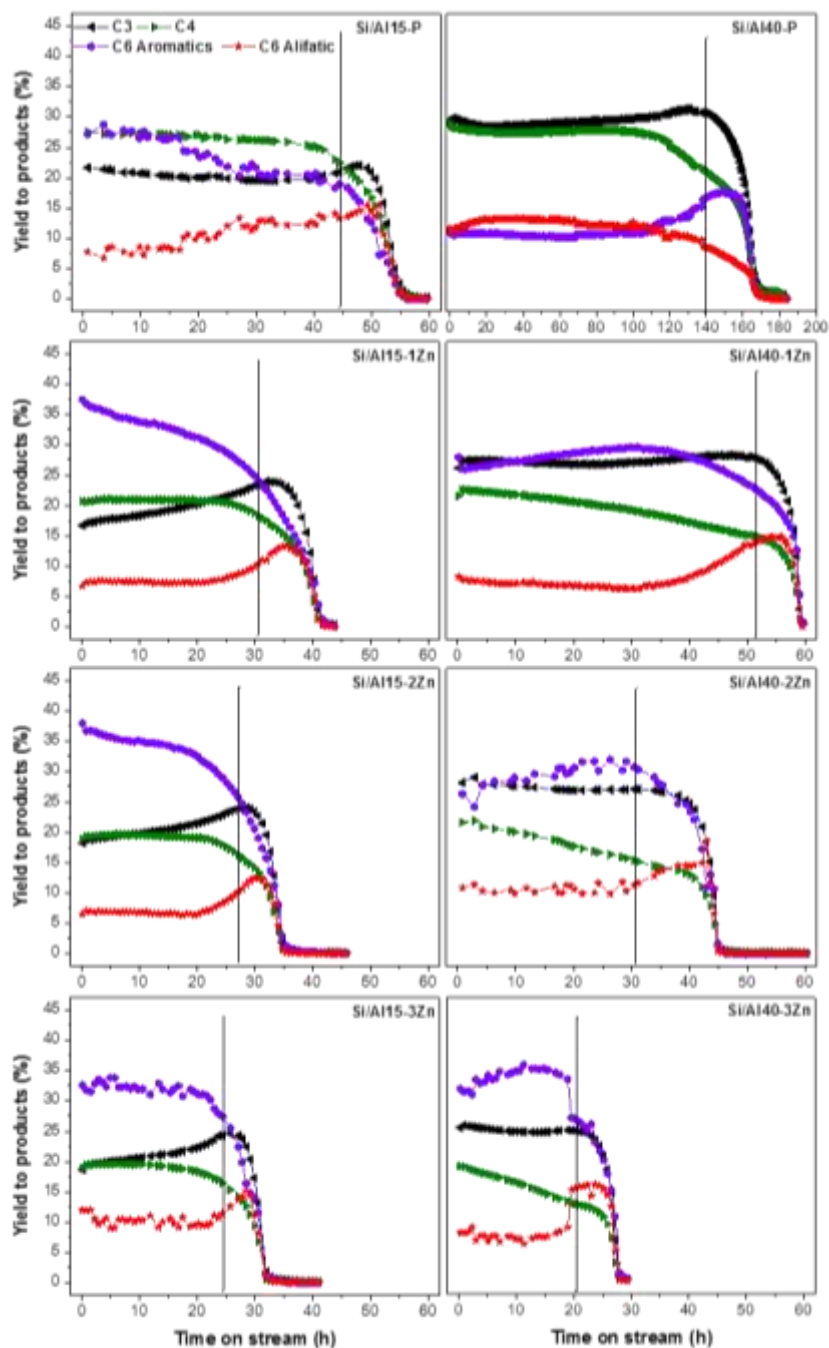


Fig. 11. Yield of hydrocarbon products vs of time on stream using HZSM-5 and Zn-exchanged HZSM-5 materials as catalysts. The same colour code is used in all the cases. Methanol breakthrough is marked with vertical lines.

4. Discussion

4.1 Distribution and nature of reactive Zn sites.

The type of Zn species that can be formed in zeolites and their role as active centers in the aromatization reaction are recurrent topics in the literature [28, 30, 31, 34, 35, 87]. For example, it has been reported that ionic species inside the microporous structure such as $\text{Zn}(\text{OH})^+$ are the responsible for the enhanced formation of aromatics rather than ZnO or metallic Zn agglomerates [28]. As described in Section 3.2, in this contribution we have used a wide range of techniques to characterize the Zn sites in MFI.

From the FTIR analysis with pyridine as probe molecule, it is clear that, in accordance with previous observations, the transition metal is incorporated at Brønsted sites in the zeolite, leading to a decrease in the density of Brønsted sites and the formation of new Lewis Zn sites (Fig. 1, Table 2). This observation is further corroborated by the solid-state NMR spectroscopy results for both series of samples (Table 4). Considering the higher probability of forming Zn sites coordinating to two Al sites in an Al rich zeolite, it is not surprising that the decrease in Brønsted acid sites is more pronounced in the case of the parent material with Si/Al=15. In the case of samples with Si/Al=40, the decrease correlates well with the amount of Zn incorporated, meaning that one Zn is incorporated per Brønsted site (Table 3). However, in the samples with Si/Al=15, the reduction in Brønsted acidity is 2.4-3.4 times higher than the amount of Zn exchanged into the samples, which clearly indicates that the Zn interacts with more than one Al center. As pointed out in Section 3.3.2, although the differences between the XAS spectra of the pre-treated samples are small, these support the presence of a slightly higher fraction of Z_2Zn^{2+} sites in the Si/Al = 15 series.

It also should be pointed out that, according to the IR results (Table 2), in both cases more Lewis acid sites are formed than Brønsted are consumed (Table 2). This interesting observation has also been reported in the literature [28, 30], and seems to be more pronounced when Zn is incorporated into the sample by ion exchange as compared to other methods, such as impregnation, direct synthesis or physical mixing [28]. The FTIR results obtained upon CO and pyridine adsorption, revealed the formation of a much higher than the stoichiometric amount of Lewis sites upon ion-exchange. From the NMR analysis it is evident that these extra Lewis sites are not related to EFAL, since extra framework Al species are being re-incorporated into the structure (especially in samples with Si/Al=15) during the ion exchange treatment. Accordingly, the most reasonable explanation is that Zn interacts with more than one probe molecule employed for the FTIR measurements. To verify that the changes in acidity are reversible and due to the presence of Zn in the materials, sample Si/Al40-1Zn was back-exchanged into the protonic form, which indeed resulted in very similar acidic properties than the parent sample (Fig. S7, Table S1).

Unlike other transition metals in zeolites, Zn is rather difficult to detect by spectroscopic means, but no significant contribution from Zn-Zn scattering was detected in EXAFS that could correspond to metallic Zn, ZnO or Zn dimers (Fig. 5). Therefore, these species are probably not formed in the material in substantial quantities. In addition to this, the increase of the signal at 0.6 ppm in the ^1H DEPTH MAS NMR spectra upon Zn incorporation, can be considered as an indication of the presence of Zn hydroxide species (Fig. 4). Thus, we can conclude that mainly monomeric Zn sites, predominantly in the form of $\text{Z}[\text{Zn}(\text{OH})]$ species are expected to be present in the samples (Fig. 5).

4.2 Catalytic performance trends and role of Zn on the catalytic behaviour of the samples

We have observed that increasing amounts of Zn leads to significant variations in product distribution and life time of the catalysts. For both sample series the presence of Zn causes an increase in the aromatization activity of the catalysts (Fig. 9, Fig. 10). According to the decrease in the calculated hydrogen transfer indices of C₄ compounds (C₄-HTI) at increasing Zn content, it is quite clear that Zn promotes the dehydrogenation pathway.

On the other hand, there is a reduction in lifetime with increasing Zn content in the zeolites. This negative effect has been observed before [84, 88] and might simply be related to the higher probability of forming deactivating coke molecules from the increased content of aromatics in the micropores of the zeolite. In fact, the most stable catalyst is Si/Al40-P, which is the one that displays the lowest yield of aromatics, because it does not have Zn sites or high enough density of Brønsted or Lewis acid sites to enhance the aromatization. As mentioned above, Lewis sites can also catalyze fast hydrogen transfer reactions leading to aromatics and oxygen containing coke species via formaldehyde [24] thus further promoting the deactivation. The extra Lewis sites formed after the ion exchange (observed with FTIR) might also be able to catalyze similar reactions. In summary, the faster deactivation of samples with same Al content but increasing Zn exchanged can be explained in terms of both the higher concentration of aromatics in this samples, and the higher number of Lewis sites. However, it should be pointed out that it is unknown at this point whether or not the Lewis sites generated by Zn display the same tendency for the detrimental formaldehyde formation as the 'regular' Lewis acid sites present in non-exchanged catalysts.

So, at this point it is quite clear that the catalytic behaviour of the samples is influenced not just by the presence (and amount) of Zn but also by the Zn/Al ratio in each case. Frequently, the outstanding catalytic performance of bifunctional metal-zeolite systems is attributed to the synergic effect of different active sites, in this case, Zn sites and Brønsted sites. This active-site synergy, which is proposed to tune the properties of the active centers and considered to be caused by proximity and interaction, has been described in previous literature as the origin of the unique catalytic properties of metal-containing zeolites [73, 89-91]. In this section, however, it is important to highlight that it is not necessarily the synergy between the Zn sites and the zeolite protonic acid sites, but simply an optimum concentration of both sites controlling which of the two possible mechanistic pathways for the aromatization that is dominating in each case. This implies that the impact of a certain amount of Zn can be different depending on the density of Brønsted acid sites. It is therefore the relative content of Al and Zn which defines if the reaction is dominated by the Zn or the zeolite and whether or not the aromatization is taking place by dehydrogenation or hydrogen transfer reactions.

To illustrate this situation, it is useful to compare the parent samples, where Si/Al15-P produced a clearly higher yield of aromatics than Si/Al40-P at the same experimental conditions. Since the parent samples do not have any dehydrogenation sites, aromatics will be formed by hydrogen transfer reactions, resulting in a high yield of alkanes and, thus, in a high hydrogen transfer index value (Fig. 9). In addition to that, we observe that samples obtained from the same parent zeolite display a systematic increase in yield of aromatics with the gradual increase Zn content. Also in both series of samples the HTI decreases markedly when Zn is incorporated in the samples (Fig. 9), but all the Zn-exchanged samples from the Si/Al=40 series have lower values of HTI than any of the

Zn-exchanged samples with Si/Al=15 at any yield of aromatics. Taking into account that the Zn content is in all the samples around 1-2%, one would expect that the effect of Zn in both materials is the same, but the marked difference in HTI indicates that the aromatization reaction mechanism is governed by the relation between the Zn content and the Si/Al ratio. In materials for which it is less likely to find a Zn site compared to a zeolite acid site, the aromatization reaction takes place mainly through the traditional hydrogen transfer reaction, as it happens in the parent samples. We may then, conclude that it is the high Si/Al ratio and therefore the higher Zn/Al ratios that lead to the lower degree of hydrogen transfer in the Si/Al40 series. This becomes very obvious comparing samples Si/Al15-3Zn and Si/Al40-3Zn: the yield of aromatics for both catalysts is around 33%, but the C4-HTI obtained with the first one is two times higher compared to the latter one (Fig.8).

To verify to which extent the presence of Zn in the materials drives the reaction through the dehydrogenation pathway, the hydrogen transfer indices as a function of the yield of aromatics are presented in Fig. 12. For the Si/Al=15 series, at the same yield of aromatics, the HTI is lower for samples with higher Zn content, while the HTI is in general lower for samples belonging to the series with Si/Al=40, which highlights the important influence that the Zn/Al has in the formation of aromatics. Nonetheless, the hydrogen transfer index is never zero, not even for the samples with a high Zn/Al ratio, which indicates that alkanes can result from hydrogen transfer reaction or hydrogenation of olefins. It can be hypothesized, then, that either the aromatization is not taking place only over the Zn sites or that there are some reaction steps involving hydrogen transfer that Zn sites are not able to catalyze by dehydrogenation.

At the same time, the incorporation of Zn by ion exchange entails a loss of Brønsted sites in the zeolitic framework, that are necessary to catalyze the formation of the hydrocarbons needed for the dehydrogenation reaction. Thus, the quantity of Zn introduced in the zeolite should be controlled maintaining an optimum ratio between Zn sites, that enhance the aromatization by dehydrogenation, and Brønsted sites which are responsible for the conversion of methanol into hydrocarbons.

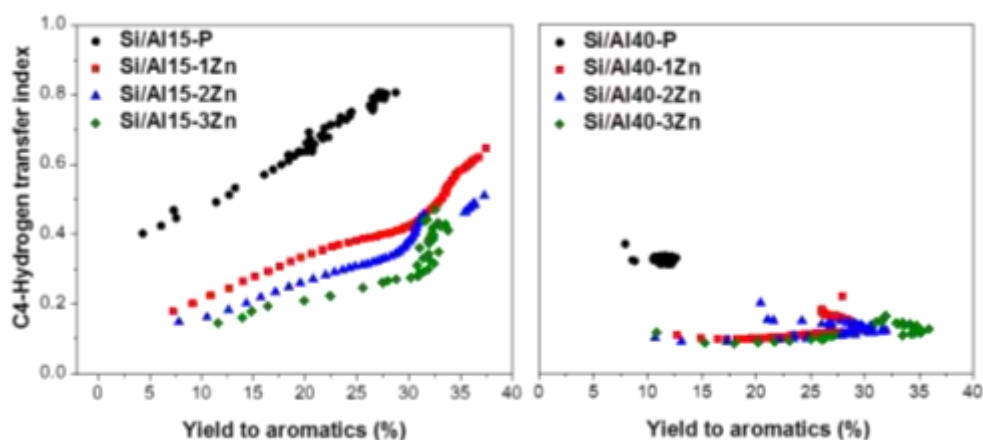


Fig. 12. Hydrogen transfer index for C₄ hydrocarbons (C₄-HTI) as a function of the yield of aromatics obtained in the reaction using HZSM-5 and Zn-HZSM-5 materials.

4.3 Structure-reactivity relationships. Role of Zn-sites in the MTA transformation.

The fact that Zn and other transition metals are active as dehydrogenating agents in metal-zeolite bifunctional catalysts is well known, but an exact mechanistic understanding is often missing. Generally, metal sites act polarizing on neighbouring molecular bonds, and Zn sites in particular show a certain affinity for hydrogen, which could accelerate the combination of surface hydrogen via dehydrogenation or dehydrocyclization [31, 92-94]. This capacity has been attributed to the presence of pairs of Lewis acid (Zn²⁺) and basic (O²⁻) species in ZnO material. Taking this into account, it can be hypothesized that the

Zn-O bond in the Zn species formed in the zeolites (primarily the $Z[Zn(OH)]$ complexes identified by XAS in the pretreated catalysts) can be comparable to those in ZnO. Thus, it is plausible that the dehydrogenation starts with the interaction of hydrogen with the formed Zn species. These Zn species can be considered, then, as a vessel for the recombinative desorption of hydrogen as molecular H_2 to happen. Noteworthy, in MTA conditions, only slight modifications in the XAS spectra are observed with respect to the activated catalysts. Interestingly, the XANES region shows clearer changes in the position and relative intensity of the two **B** and **C** peaks, whereas EXAFS remains almost unchanged (see Fig. 6). These evidences could be consistent with interaction of H with framework-coordinated Zn-sites during the catalysis. Indeed, it is known that XANES provides a highest sensitivity in such case, probing the perturbation in the electronic structure of the metal centre upon the formation of H-adducts [85, 86], whereas the presence of H in the coordination sphere of Zn is poorly detectable by EXAFS due to its very low scattering amplitude.

As a summary, $Z[Zn(OH)]$ species in zeolites are most likely responsible for the dehydrogenation reactions that leads to aromatization. However, to be able to obtain an aromatics rich mixture of hydrocarbons, inhibiting the formation of undesired light alkane by-products it is necessary to control not just the Zn content in the zeolite but also the density of Brønsted sites of the zeolite. As discussed in the previous section, some Brønsted acidity is necessary to build olefins as precursors of the formation of aromatics, but if the density of Brønsted sites is too high, the reaction will be dominated by the regular zeolite sites and not the Zn sites, and the hydrogen transfer mechanism would be favoured.

5. Conclusions

The catalytic performance of bifunctional Zn-ZSM-5 catalysts was studied in the methanol to aromatics transformation. Zn-ZSM-5 samples were produced by ion exchanging two parent materials with Si/Al=15 and 40 up to three times, thus increasing the amount of Zn incorporated in the material. Zn speciation in the materials was also studied. Characterization results evidence that the dominant Zn-species formed in the material are $\text{Zn}(\text{OH})^+$ complexes and there is no evidence for the formation of large ZnO or metallic aggregates. Zn is incorporated in the structure mainly over the Brønsted sites, forming new Lewis sites and consuming Brønsted acid centers.

The presence of Zn in the catalyst clearly influences the catalytic performance of the samples in MTA both in terms of product selectivity and catalyst lifetime. With increasing Zn content, the yield of aromatics increases in comparison with the corresponding parent zeolite, but the catalyst lifetime progressively decreases. ~~It is also clear from our data that the Zn/Al ratio can be considered as a good descriptor of the aromatization activity.~~ Comparing samples with similar Zn content, a higher density of acid sites (lower Si/Al ratio in the parent sample) favors the hydrogen transfer reactions, which translates into a higher yield of aromatics but also a higher production of alkanes. On the other hand, in samples with lower density of Al sites, the aromatization is governed by the Zn sites rather than by the zeolite acid centers and the aromatization takes place mainly by dehydrogenation, thus increasing the yield of aromatics while keeping the hydrogen transfer index and the production of alkanes low.

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Supplementary Material

Supplementary data associated with this article can be found in the online version.

References

- [1] M. Bender, *New Technologies and Alternative Feedstocks in Petrochemistry and Refining*, Preprints of the DGMK, (2013), 59.
- [2] E. Lalik, X. Liu, J. Klinowski, *J. Phys. Chem.*, 96 (1992) 805.
- [3] Y. Ono, K. Kanae, *J. Chem. Soc., Faraday Trans.*, 87 (1991) 669.
- [4] C.D. Chang, A.J. Silvestri, *J. Catal.*, 47 (1977) 249.
- [5] D. Xiang, Y. Qian, Y. Man, S. Yang, *Appl. Energy*, 113 (2014) 639.
- [6] A.M. Niziolek, O. Onel, C.A. Floudas, *AIChE J.*, 62 (2016) 1531.
- [7] G.A. Olah, A. Goepfert, G.K.S. Prakash, *J. Org. Chem.*, 74 (2009) 487.
- [8] M. Martín, *Engineering*, 3 (2017) 166.
- [9] U. Olsbye, M. Bjørgen, S. Svelle, K.-P. Lillerud, S. Kolboe, *Catal. Today*, 106 (2005) 108.
- [10] S. Svelle, F. Joensen, J. Nerlov, U. Olsbye, K.-P. Lillerud, S. Kolboe, M. Bjørgen, *J. Am. Chem. Soc.*, 128 (2006) 14770.
- [11] M. Bjørgen, S. Svelle, F. Joensen, J. Nerlov, S. Kolboe, F. Bonino, L. Palumbo, S. Bordiga, U. Olsbye, *J. Catal.*, 249 (2007) 195.
- [12] T.V.W. Janssens, S. Svelle, U. Olsbye, *J. Catal.*, 308 (2013) 122.
- [13] M. Stöcker, *J. Catal.*, 29 (1999) 3.
- [14] A.G. Gayubo, A.T. Aguayo, A. Alonso, A. Atutxa, J. Bilbao, *Catal. Today*, 106 (2005) 112.
- [15] D. Rojo-Gama, M. Signorile, F. Bonino, S. Bordiga, U. Olsbye, K.P. Lillerud, P. Beato, S. Svelle, *J. Catal.*, 351 (2017) 33.
- [16] I. Pinilla-Herrero, U. Olsbye, C. Márquez-Álvarez, E. Sastre, *J. Catal.*, 352 (2017) 191.
- [17] J.S. Martinez-Espin, M. Morten, T.V.W. Janssens, S. Svelle, P. Beato, U. Olsbye, *Catal. Sci. Technol.*, 7 (2017) 2700.

- [18] H. Schulz, *Catal. Today*, 154 (2010) 183.
- [19] S. Müller, Y. Liu, M. Vishnuvarthan, X. Sun, A.C. van Veen, G.L. Haller, M. Sanchez-Sanchez, J.A. Lercher, *J. Catal.*, 325 (2015) 48.
- [20] U. Olsbye, S. Svelle, K.P. Lillerud, Z.H. Wei, Y.Y. Chen, J.F. Li, J.G. Wang, W.B. Fan, *Chem. Soc. Rev.*, 44 (2015) 7155.
- [21] A. Galadima, O. Muraza, *Ind. Eng. Chem. Res.*, 54 (2015) 4891.
- [22] U. Olsbye, S. Svelle, M. Bjørgen, P. Beato, T.V.W. Janssens, F. Joensen, S. Bordiga, K.P. Lillerud, *Angew. Chem. Int. Ed.*, 51 (2012) 5810.
- [23] S. Ilias, A. Bhan, *ACS Catal.*, 3 (2013) 18.
- [24] S. Müller, Y. Liu, F.M. Kirchberger, M. Tonigold, M. Sanchez-Sanchez, J.A. Lercher, *J. Am. Chem. Soc.*, 138 (2016) 15994.
- [25] C. Song, K. Liu, D. Zhang, S. Liu, X. Li, S. Xie, L. Xu, *Appl. Catal., A*, 470 (2014) 15.
- [26] M. Conte, J.A. Lopez-Sanchez, Q. He, D.J. Morgan, Y. Ryabenkova, J.K. Bartley, A.F. Carley, S.H. Taylor, C.J. Kiely, K. Khalid, G.J. Hutchings, *Catal. Sci. Technol.*, 2 (2012) 105.
- [27] R. Barthos, T. Bánsági, T. Süli Zakar, F. Solymosi, *J. Catal.*, 247 (2007) 368.
- [28] X. Niu, J. Gao, Q. Miao, M. Dong, G. Wang, W. Fan, Z. Qin, J. Wang, *Microporous Mesoporous Mater.*, 197 (2014) 252.
- [29] T. Gong, L. Qin, J. Lu, H. Feng, *Phys. Chem. Chem. Phys.*, 18 (2016) 601.
- [30] X. Niu, J. Gao, K. Wang, Q. Miao, M. Dong, G. Wang, W. Fan, Z. Qin, J. Wang, *Fuel Process. Technol.*, 157 (2017) 99.
- [31] J.A. Biscardi, G.D. Meitzner, E. Iglesia, *J. Catal.*, 179 (1998) 192.
- [32] J. Zhang, W. Qian, C. Kong, F. Wei, *ACS Catal.*, 5 (2015) 2982.
- [33] G. Qi, J. Xu, J. Su, J. Chen, X. Wang, F. Deng, *J. Am. Chem. Soc.*, 135 (2013) 6762.
- [34] H. Berndt, G. Lietz, J. Völter, *Appl. Catal., A*, 146 (1996) 365.
- [35] H. Berndt, G. Lietz, B. Lücke, J. Völter, *Appl. Catal., A*, 146 (1996) 351.
- [36] A.A. Gabrienko, S.S. Arzumanov, A.V. Toktarev, I.G. Danilova, I.P. Prosvirin, V.V. Kriventsov, V.I. Zaikovskii, D. Freude, A.G. Stepanov, *ACS Catal.*, 7 (2017) 1818.
- [37] S. Bordiga, C. Lamberti, F. Bonino, A. Travert, F. Thibault-Starzyk, *Chem. Soc. Rev.*, 44 (2015) 7262.
- [38] S. Bordiga, E. Groppo, G. Agostini, J.A. van Bokhoven, C. Lamberti, *Chem. Rev.*, 113 (2013) 1736.
- [39] K.K. Matthias Thommes, Alexander V. Neimark, James P. Olivier, Francisco Rodriguez-Reinoso, Jean Rouquerol and Kenneth S. W. Sing, *Pure Appl. Chem.*, 87 (2015) 1051.
- [40] B.C. Lippens, J.H. de Boer, *J. Catal.*, 4 (1965) 319.
- [41] E. Selli, L. Forni, *Microporous Mesoporous Mater.*, 31 (1999) 129.
- [42] O. Mathon, A. Beteva, J. Borrel, D. Bugnazet, S. Gatla, R. Hino, I. Kantor, T. Mairs, M. Munoz, S. Pasternak, F. Perrin, S. Pascarelli, *J. Synchrotron Radiat.*, 22 (2015) 1548.
- [43] D. Bellet, B. Gorges, A. Dallery, P. Bernard, E. Pereiro, J. Baruchel, *J. Appl. Crystallogr.*, 36 (2003) 366.
- [44] C. Lamberti, S. Bordiga, F. Bonino, C. Prestipino, G. Berlier, L. Capello, F. D'Acapito, F. Xamena, A. Zecchina, *Phys. Chem. Chem. Phys.*, 5 (2003) 4502.

- [45] C. Lamberti, S. Bordiga, D. Arduino, A. Zecchina, F. Geobaldo, G. Spanó, F. Genoni, G. Petrini, A. Carati, F. Villain, G. Vlaic, *J. Phys. Chem. B*, 102 (1998) 6382.
- [46] B. Ravel, M. Newville, *J. Synchrotron Radiat*, 12 (2005) 537.
- [47] A. Martini, E. Borfecchia, K.A. Lomachenko, I.A. Pankin, C. Negri, G. Berlier, P. Beato, H. Falsig, S. Bordiga, C. Lamberti, *Chem. Sci.*, 8 (2017) 6836.
- [48] E. Borfecchia, K.A. Lomachenko, F. Giordanino, H. Falsig, P. Beato, A.V. Soldatov, S. Bordiga, C. Lamberti, *Chem. Sci.*, 6 (2015) 548.
- [49] Ch. Baerlocher, L.B. McCusker, Database of Zeolite Structures: <<http://www.iza-structure.org/databases/>>.
- [50] G. te Velde, F.M. Bickelhaupt, E.J. Baerends, C. Fonseca Guerra, S.J.A. van Gisbergen, J.G. Snijders, T. Ziegler, *J. Comput. Chem.*, 22 (2001) 931.
- [51] E. Van Lenthe, E.J. Baerends, *J. Comput. Chem.*, 24 (2003) 1142.
- [52] J.P. Perdew, K. Burke, M. Ernzerhof, *Phys. Rev. Lett.*, 77 (1996) 3865.
- [53] E.v. Lenthe, A. Ehlers, E.-J. Baerends, *J. Chem. Phys.*, 110 (1999) 8943.
- [54] H. Lerner, M. Draeger, J. Steffen, K.K. Unger, *Zeolites*, 5 (1985) 131.
- [55] S.I. Zabinsky, J.J. Rehr, A. Ankudinov, R.C. Albers, M.J. Eller, *Phys. Rev. B*, 52 (1995) 2995.
- [56] J.J. Rehr, R.C. Albers, *Rev. Mod. Phys.*, 72 (2000) 621.
- [57] D.K. Pappas, E. Borfecchia, M. Dybala, I.A. Pankin, K.A. Lomachenko, A. Martini, M. Signorile, S. Teketel, B. Arstad, G. Berlier, C. Lamberti, S. Bordiga, U. Olsbye, K.P. Lillerud, S. Svelle, P. Beato, *J. Am. Chem. Soc.*, 139 (2017) 14961.
- [58] F. Bonino, S. Chavan, J.G. Vitillo, E. Groppo, G. Agostini, C. Lamberti, P.D.C. Dietzel, C. Prestipino, S. Bordiga, *Chem. Mater.*, 20 (2008) 4957.
- [59] P. Sazama, B. Wichterlova, J. Dedecek, Z. Tvaruzkova, Z. Musilova, L. Palumbo, S. Sklenak, O. Gonsiorova, *Microporous Mesoporous Mater.*, 143 (2011) 87.
- [60] M.B. Sayed, R.A. Kydd, R.P. Cooney, *J. Catal.*, 88 (1984) 137.
- [61] B. Onida, F. Geobaldo, F. Testa, F. Crea, E. Garrone, *Microporous Mesoporous Mater.*, 30 (1999) 119.
- [62] V.B. Kazansky, V.Y. Borovkov, A.I. Serikh, R.A. van Santen, B.G. Anderson, *Catal. Lett.*, 66 (2000) 39.
- [63] P. D'Angelo, M. Benfatto, S. Della Longa, N.V. Pavel, *Phys. Rev. B*, 66 (2002).
- [64] J.E. Penner-Hahn, *Coord. Chem. Rev.*, 249 (2005) 161.
- [65] A. Oda, T. Ohkubo, T. Yumura, H. Kobayashi, Y. Kuroda, *Dalton Trans.*, 44 (2015) 10038.
- [66] S.M.T. Almutairi, B. Mezari, P. Magusin, E.A. Pidko, E.J.M. Hensen, *ACS Catal.*, 2 (2012) 71.
- [67] V. Yachandra, L. Powers, T.G. Spiro, *J. Am. Chem. Soc.*, 105 (1983) 6596.
- [68] K. Peariso, Z.H.S. Zhou, A.E. Smith, R.G. Matthews, J.E. Penner-Hahn, *Biochemistry*, 40 (2001) 987.
- [69] K. Clark-Baldwin, D.L. Tierney, N. Govindaswamy, E.S. Gruff, C. Kim, J. Berg, S.A. Koch, J.E. Penner-Hahn, *J. Am. Chem. Soc.*, 120 (1998) 8401.
- [70] A. Oda, H. Torigoe, A. Itadani, T. Ohkubo, T. Yumura, H. Kobayashi, Y. Kuroda, *Angew. Chem. Int. Ed.*, 51 (2012) 7719.
- [71] A.L. Yakovlev, A.A. Shubin, G.M. Zhidomirov, R.A. van Santen, *Catal. Lett.*, 70 (2000) 175.
- [72] E.A. Pidko, R.A. van Santen, *J. Phys. Chem. C*, 111 (2007) 2643.
- [73] Y. Xin, P. Qi, X. Duan, H. Lin, Y. Yuan, *Catal. Lett.*, 143 (2013) 798.
- [74] T.V.W. Janssens, *J. Catal.*, 264 (2009) 130.

- [75] M. Bjørgen, F. Joensen, M. Spangsborg Holm, U. Olsbye, K.-P. Lillerud, S. Svelle, *Appl. Catal., A*, 345 (2008) 43.
- [76] A.L. Yakovlev, A.A. Shubin, G.M. Zhidomirov, R.A. van Santen, *Catal. Lett.*, 70 (2000) 175.
- [77] Y. Ni, A. Sun, X. Wu, G. Hai, J. Hu, T. Li, G. Li, *Microporous Mesoporous Mater.*, 143 (2011) 435.
- [78] P. Meriaudeau, C. Naccache, *J. Mol. Catal.*, 59 (1990) L31.
- [79] A.G. Stepanov, S.S. Arzumanov, A.A. Gabrienko, V.N. Parmon, I.I. Ivanova, D. Freude, *ChemPhysChem*, 9 (2008) 2559.
- [80] G. Qi, Q. Wang, J. Xu, J. Trébosc, O. Lafon, C. Wang, J.-P. Amoureux, F. Deng, *Angew. Chem. Int. Ed.*, 55 (2016) 15826.
- [81] G. Buckles, G.J. Hutchings, C.D. Williams, *Catal. Lett.*, 11 (1991) 89.
- [82] L.M. Lubango, M.S. Scurrall, *Appl. Catal., A*, 235 (2002) 265.
- [83] C. Lamberti, A. Zecchina, E. Groppo, S. Bordiga, *Chem. Soc. Rev.*, 39 (2010) 4951.
- [84] D. Scarano, S. Bertarione, G. Spoto, A. Zecchina, C. Otero Areán, *Thin Solid Films*, 400 (2001) 50.
- [85] A.L. Bugaev, A.A. Guda, A. Lazzarini, K.A. Lomachenko, E. Groppo, R. Pellegrini, A. Piovano, H. Emerich, A.V. Soldatov, L.A. Bugaev, V.P. Dmitriev, J.A. van Bokhoven, C. Lamberti, *Catal. Today* 283 (2017) 119.
- [86] A.L. Bugaev, A.A. Guda, K.A. Lomachenko, V.V. Srabionyan, L.A. Bugaev, A.V. Soldatov, C. Lamberti, V.P. Dmitriev, J.A. van Bokhoven, *J. Phys. Chem. C*, 118 (2014) 10416.

TABLE CAPTIONS

Table 1. Experimental details of the Zn incorporation in two parent ZSM-5 zeolites with different Si/Al ratios.

Table 2. Elemental composition of the samples measured by ICP and density of acid sites detected from the FTIR experiments using pyridine as probe molecule.

Table 3. Adsorption capacity and textural properties of the parent and Zn-exchanged ZSM-5 samples*.

Table 4. Si/Al ratios from ICP-OES and ^{29}Si NMR, and concentrations in mmol per gram of different sites determined by ^{29}Si , ^{27}Al and ^1H NMR.

Table 5. Results from EXAFS fit for the Si/Al40-3Zn activated catalyst (data collection temperature: 420 °C) using the DFT-optimized $\mathbf{Z}[\text{Zn}(\text{OH})]$ geometry shown in Fig. 7. The fit was performed in R-space, in the range (1.0 – 3.3) Å, employing the k-range (2.5 – 11.0) Å $^{-1}$ for the FT (number of independent point $N_{\text{ind}} = 2\Delta k\Delta R/\pi > 17$). Under round brackets, the DFT optimized distances are reported for comparison. Atoms are labelled with the subscript *fw* when are part of the framework and with *ef* in the case of extra framework

Table 6. Cumulative conversion capacity, turnover number and calculated values of the deactivation coefficient and the constant rate of the zeolitic materials in MTA transformation at the experimental conditions tested.

Table 7. Average selectivity to aromatics calculated between 100 and 50% conversion for all the catalysts studied in MTA transformation at the experimental conditions described.

FIGURE CAPTIONS

Fig. 1. FTIR spectra of the parent H-ZSM-5 and the Zn-containing HZSM-5 zeolites. **a)** FTIR spectra of the pre-treated samples in the OH stretching region. **b)** FTIR spectra of pyridine adsorbed on the zeolites.

Fig. 2. Background-subtracted FTIR spectra of CO (0.1 mbar) adsorbed at -196 °C on parent H-ZSM-5 and Zn-containing HZSM-5 zeolites.

Fig. 3. Single-pulse ^{27}Al MAS NMR spectra of the two sample series Si/Al=15 and Si/Al=40 ($B_0 = 22.3$ T, $\nu_R = 30.0$ kHz). Four-, five- and six-fold surrounded ^{27}Al environments at chemical shifts of around 55, 30 and 0 ppm, respectively, are indicated in the spectrum from Si/Al15-P.

Fig. 4. ^1H DEPTH MAS NMR spectra ($B_0 = 14.1$ T, $\nu_R = 40.0$ kHz) of the two sample series Si/Al=15 (left) and Si/Al=40 (right).

Fig. 5. Typical XAS fingerprints of Zn sites in Zn-ZSM-5 at key process steps. (a) Zn K-edge XANES for a representative Zn-ZSM-5 sample (Si/Al40-3Zn) in the hydrated state

at RT, after pretreatment in air at 500 °C and cooling to 420 °C and during MTA at the same temperature. For comparison, the XANES spectra for a series of reference compounds (measured at RT) are also reported, i.e. $\text{Zn}(\text{NO}_3)_2$ aqueous solution, representative of fully hydrated Zn^{2+} ions, Zn metal foil and ZnO. (b, c) Magnitude (b) and imaginary part (c) of the FT-EXAFS spectra for the same sample and reference compounds reported in part (a), using the same colour code. All the FT-EXAFS spectra have been obtained by transforming the $k^2\chi(k)$ EXAFS functions in the 2.5-12.5 \AA^{-1} range.

Fig. 6. Zn K-edge XANES (main panels) and FT-EXAFS spectra (insets) for Zn-ZSM-5 catalysts with different composition in term of Si/Al in the parent zeolite and Zn-loading (from 0.7 to 1.9 wt%) collected at key process steps: (a) hydrated catalysts at RT; (b) pretreated catalysts in air at 500 °C and cooled at 420 °C; (c) during MTA at 420 °C. All the FT-EXAFS spectra have been obtained by transforming the $k^2\chi(k)$ functions in the 2.5-12.5 \AA^{-1} range.

Fig. 7. (a) DFT-optimized geometry for the $\text{Z}[\text{Zn}(\text{OH})]$ complex in the 5R of the ZSM-5 framework. (b) view of the full optimized cluster with the $\text{Z}[\text{Zn}(\text{OH})]$ complex highlighted by a grey circle; magnified view of the Zn local environment in the 5R, with coloured circles indicating the different coordination shells included in the EXAFS fit and DFT distances of each atom from the Zn absorber expressed in \AA . Atom colour code as follows: Zn, purple; H, white; O, red; Al, pink; Si, yellow. (c,d) Comparison between modulus (c) and imaginary part (d) of the experimental FT-EXAFS spectra (black circles) and corresponding best fit (solid grey lines) for Si/Al40-3Zn pretreated catalyst (data collection temperature: 420 °C) using the DFT-optimized $\text{Z}[\text{Zn}(\text{OH})]$ geometry shown in Fig. 7. The principal contributions to the EXAFS signal are also shown, vertically translated for clarity; the values of the parameters optimized in the fits are reported in Table 5. Atoms are labelled with the subscript fw when are part of the framework and with ef in the case of extra framework

Fig. 8. Conversion of oxygenates (DME+MeOH) at 400 °C and $\text{WHSV} = 3 \text{ h}^{-1}$ vs time on stream over the parent and exchanged zeolites with the two selected Si/Al ratios.

Fig. 9. Yield of aromatics and hydrogen transfer index after 4h on stream over parent and exchanged samples.

Fig. 10. Yield of aromatics as a function of time on stream using HZSM-5 and Zn-exchanged HZSM-5 materials as catalysts. Methanol breakthrough is marked with vertical lines for each catalyst.

Fig. 11. Yield of hydrocarbon products vs of time on stream using HZSM-5 and Zn-exchanged HZSM-5 materials as catalysts. The same colour code is used in all the cases. Methanol breakthrough is marked with vertical lines.

Fig. 12. Hydrogen transfer index for C₄ hydrocarbons (C₄-HTI) as a function of the yield of aromatics obtained in the reaction using HZSM-5 and Zn-HZSM-5 materials.

- [1] M. Bender, *New Technologies and Alternative Feedstocks in Petrochemistry and Refining*, Desden, Germany, 2013.
- [2] E. Lalik, X. Liu, J. Klinowski, *The Journal of Physical Chemistry*, 96 (1992) 805.
- [3] Y. Ono, K. Kanae, *Journal of the Chemical Society, Faraday Transactions*, 87 (1991) 669.
- [4] C.D. Chang, A.J. Silvestri, *Journal of Catalysis*, 47 (1977) 249.
- [5] D. Xiang, Y. Qian, Y. Man, S. Yang, *Applied Energy*, 113 (2014) 639.
- [6] A. M. Niziolek, O. Onel, C. A. Floudas, *Production of benzene, toluene, and xylenes from natural gas via methanol: Process synthesis and global optimization*, 2015.
- [7] G.A. Olah, A. Goepfert, G.K.S. Prakash, *The Journal of Organic Chemistry*, 74 (2009) 487.
- [8] M. Martín, *Engineering*, 3 (2017) 166.
- [9] U. Olsbye, M. Bjørgen, S. Svelle, K.-P. Lillerud, S. Kolboe, *Catalysis Today*, 106 (2005) 108.
- [10] S. Svelle, F. Joensen, J. Nerlov, U. Olsbye, K.-P. Lillerud, S. Kolboe, M. Bjørgen, *Journal of the American Chemical Society*, 128 (2006) 14770.
- [11] M. Bjørgen, S. Svelle, F. Joensen, J. Nerlov, S. Kolboe, F. Bonino, L. Palumbo, S. Bordiga, U. Olsbye, *Journal of Catalysis*, 249 (2007) 195.
- [12] T.V.W. Janssens, S. Svelle, U. Olsbye, *Journal of Catalysis*, 308 (2013) 122.
- [13] M. Stöcker, *Microporous and Mesoporous Materials*, 29 (1999) 3.
- [14] A.G. Gayubo, A.T. Aguayo, A. Alonso, A. Atutxa, J. Bilbao, *Catalysis Today*, 106 (2005) 112.
- [15] D. Rojo-Gama, M. Signorile, F. Bonino, S. Bordiga, U. Olsbye, K.P. Lillerud, P. Beato, S. Svelle, *Journal of Catalysis*, 351 (2017) 33.
- [16] I. Pinilla-Herrero, U. Olsbye, C. Márquez-Álvarez, E. Sastre, *Journal of Catalysis*, 352 (2017) 191.
- [17] J.S. Martinez-Espin, M. Morten, T.V.W. Janssens, S. Svelle, P. Beato, U. Olsbye, *Catalysis Science & Technology*, (2017).
- [18] H. Schulz, *Catalysis Today*, 154 (2010) 183.
- [19] S. Müller, Y. Liu, M. Vishnuvarthan, X. Sun, A.C. van Veen, G.L. Haller, M. Sanchez-Sanchez, J.A. Lercher, *Journal of Catalysis*, 325 (2015) 48.
- [20] U. Olsbye, S. Svelle, K.P. Lillerud, Z.H. Wei, Y.Y. Chen, J.F. Li, J.G. Wang, W.B. Fan, *Chemical Society Reviews*, 44 (2015) 7155.

- [21] A. Galadima, O. Muraza, *Industrial & Engineering Chemistry Research*, 54 (2015) 4891.
- [22] U. Olsbye, S. Svelle, M. Bjørngen, P. Beato, T.V.W. Janssens, F. Joensen, S. Bordiga, K.P. Lillerud, *Angewandte Chemie International Edition*, 51 (2012) 5810.
- [23] S. Ilias, A. Bhan, *ACS Catalysis*, 3 (2013) 18.
- [24] S. Müller, Y. Liu, F.M. Kirchberger, M. Tonigold, M. Sanchez-Sanchez, J.A. Lercher, *Journal of the American Chemical Society*, 138 (2016) 15994.
- [25] C. Song, K. Liu, D. Zhang, S. Liu, X. Li, S. Xie, L. Xu, *Applied Catalysis A: General*, 470 (2014) 15.
- [26] M. Conte, J.A. Lopez-Sanchez, Q. He, D.J. Morgan, Y. Ryabenkova, J.K. Bartley, A.F. Carley, S.H. Taylor, C.J. Kiely, K. Khalid, G.J. Hutchings, *Catalysis Science & Technology*, 2 (2012) 105.
- [27] R. Barthos, T. Bánsági, T. Süli Zakar, F. Solymosi, *Journal of Catalysis*, 247 (2007) 368.
- [28] X. Niu, J. Gao, Q. Miao, M. Dong, G. Wang, W. Fan, Z. Qin, J. Wang, *Microporous and Mesoporous Materials*, 197 (2014) 252.
- [29] T. Gong, L. Qin, J. Lu, H. Feng, *Physical Chemistry Chemical Physics*, 18 (2016) 601.
- [30] X. Niu, J. Gao, K. Wang, Q. Miao, M. Dong, G. Wang, W. Fan, Z. Qin, J. Wang, *Fuel Processing Technology*, 157 (2017) 99.
- [31] J.A. Biscardi, G.D. Meitzner, E. Iglesia, *Journal of Catalysis*, 179 (1998) 192.
- [32] J. Zhang, W. Qian, C. Kong, F. Wei, *ACS Catalysis*, 5 (2015) 2982.
- [33] G. Qi, J. Xu, J. Su, J. Chen, X. Wang, F. Deng, *Journal of the American Chemical Society*, 135 (2013) 6762.
- [34] H. Berndt, G. Lietz, J. Völter, *Applied Catalysis A: General*, 146 (1996) 365.
- [35] H. Berndt, G. Lietz, B. Lücke, J. Völter, *Applied Catalysis A: General*, 146 (1996) 351.
- [36] A.A. Gabrienko, S.S. Arzumanov, A.V. Toktarev, I.G. Danilova, I.P. Prosvirin, V.V. Kriventsov, V.I. Zaikovskii, D. Freude, A.G. Stepanov, *ACS Catalysis*, 7 (2017) 1818.
- [37] S. Bordiga, C. Lamberti, F. Bonino, A. Travert, F. Thibault-Starzyk, *Probing zeolites by vibrational spectroscopies*, 2015.
- [38] S. Bordiga, E. Groppo, G. Agostini, J.A. van Bokhoven, C. Lamberti, *Chemical Reviews*, 113 (2013) 1736.
- [39] K.K. Matthias Thommes, Alexander V. Neimark, James P. Olivier, Francisco Rodriguez-Reinoso, Jean Rouquerol and Kenneth S. W. Sing, *Pure and Applied Chemistry*, 87 (2015) 1051.
- [40] B.C. Lippens, J.H. de Boer, *Journal of Catalysis*, 4 (1965) 319.
- [41] E. Selli, L. Forni, *Microporous and Mesoporous Materials*, 31 (1999) 129.
- [42] D.G. Cory, W.M. Ritchey, *Journal of Magnetic Resonance* (1969), 80 (1988) 128.
- [43] H. Huo, L. Peng, C.P. Grey, *The Journal of Physical Chemistry C*, 113 (2009) 8211.
- [44] O. Mathon, A. Beteva, J. Borrel, D. Bugnazet, S. Gatla, R. Hino, I. Kantor, T. Mairs, M. Munoz, S. Pasternak, F. Perrin, S. Pascarelli, *Journal of Synchrotron Radiation*, 22 (2015) 1548.
- [45] D. Bellet, B. Gorges, A. Dallery, P. Bernard, E. Pereiro, J. Baruchel, *Journal of Applied Crystallography*, 36 (2003) 366.
- [46] C. Lamberti, S. Bordiga, F. Bonino, C. Prestipino, G. Berlier, L. Capello, F. D'Acapito, F. Xamena, A. Zecchina, *Physical Chemistry Chemical Physics*, 5 (2003) 4502.
- [47] C. Lamberti, S. Bordiga, D. Arduino, A. Zecchina, F. Geobaldo, G. Spanó, F. Genoni, G. Petrini, A. Carati, F. Villain, G. Vlaic, *The Journal of Physical Chemistry B*, 102 (1998) 6382.
- [48] B. Ravel, M. Newville, *Journal of Synchrotron Radiation*, 12 (2005) 537.
- [49] A. Martini, E. Borfecchia, K.A. Lomachenko, I.A. Pankin, C. Negri, G. Berlier, P. Beato, H. Falsig, S. Bordiga, C. Lamberti, *Chemical Science*, 8 (2017) 6836.
- [50] E. Borfecchia, K.A. Lomachenko, F. Giordanino, H. Falsig, P. Beato, A.V. Soldatov, S. Bordiga, C. Lamberti, *Chemical Science*, 6 (2015) 548.
- [51] D.o.Z.s.h.w.i.-s.o.d. Ch. Baerlocher and L.B. McCusker.
- [52] G. te Velde, F.M. Bickelhaupt, E.J. Baerends, C. Fonseca Guerra, S.J.A. van Gisbergen, J.G. Snijders, T. Ziegler, *Journal of Computational Chemistry*, 22 (2001) 931.

- [53] E. Van Lenthe, E.J. Baerends, *Journal of Computational Chemistry*, 24 (2003) 1142.
- [54] J.P. Perdew, K. Burke, M. Ernzerhof, *Physical Review Letters*, 77 (1996) 3865.
- [55] E.v. Lenthe, A. Ehlers, E.-J. Baerends, *The Journal of Chemical Physics*, 110 (1999) 8943.
- [56] H. Lermer, M. Draeger, J. Steffen, K.K. Unger, *Zeolites*, 5 (1985) 131.
- [57] S.I. Zabinsky, J.J. Rehr, A. Ankudinov, R.C. Albers, M.J. Eller, *Physical Review B*, 52 (1995) 2995.
- [58] J.J. Rehr, R.C. Albers, *Reviews of Modern Physics*, 72 (2000) 621.
- [59] D.K. Pappas, E. Borfecchia, M. Dyballa, I.A. Pankin, K.A. Lomachenko, A. Martini, M. Signorile, S. Teketel, B. Arstad, G. Berlier, C. Lamberti, S. Bordiga, U. Olsbye, K.P. Lillerud, S. Svelle, P. Beato, *Journal of the American Chemical Society*, 139 (2017) 14961.
- [60] F. Bonino, S. Chavan, J.G. Vitillo, E. Groppo, G. Agostini, C. Lamberti, P.D.C. Dietzel, C. Prestipino, S. Bordiga, *Chemistry of Materials*, 20 (2008) 4957.
- [61] P. Sazama, B. Wichterlova, J. Dedecek, Z. Tvaruzkova, Z. Musilova, L. Palumbo, S. Sklenak, O. Gonsiorova, *Microporous and Mesoporous Materials*, 143 (2011) 87.
- [62] M.B. Sayed, R.A. Kydd, R.P. Cooney, *Journal of Catalysis*, 88 (1984) 137.
- [63] B. Onida, F. Geobaldo, F. Testa, F. Crea, E. Garrone, *Microporous and Mesoporous Materials*, 30 (1999) 119.
- [64] V.B. Kazansky, V.Y. Borovkov, A.I. Serikh, R.A. van Santen, B.G. Anderson, *Catalysis Letters*, 66 (2000) 39.
- [65] W. Loewenstein, *American Mineralogist*, 39 (1954) 92.
- [66] N. Malicki, G. Mali, A.-A. Quoineaud, P. Bourges, L.J. Simon, F. Thibault-Starzyk, C. Fernandez, *Microporous and Mesoporous Materials*, 129 (2010) 100.
- [67] W. Luo, E. R. H. van Eck, P. Bruijninx, B. Weckhuysen, Influence of Levulinic Acid Hydrogenation on Aluminum Coordination in Zeolite-Supported Ruthenium Catalysts: A 27 Al 3QMAS Nuclear Magnetic Resonance Study, 2017.
- [68] M. Hunger, *Solid State Nuclear Magnetic Resonance*, 6 (1996) 1.
- [69] M. Hunger, S. Ernst, S. Steuernagel, J. Weitkamp, *Microporous Materials*, 6 (1996) 349.
- [70] W. Zhang, M. Sun, R. Prins, *The Journal of Physical Chemistry B*, 106 (2002) 11805.
- [71] Y. Jiang, J. Huang, W. Dai, M. Hunger, *Solid State Nuclear Magnetic Resonance*, 39 (2011) 116.
- [72] J.-F. Wu, S.-M. Yu, W.D. Wang, Y.-X. Fan, S. Bai, C.-W. Zhang, Q. Gao, J. Huang, W. Wang, *Journal of the American Chemical Society*, 135 (2013) 13567.
- [73] G. Qi, Q. Wang, J. Xu, J. Trébosc, O. Lafon, C. Wang, J.-P. Amoureux, F. Deng, *Angewandte Chemie International Edition*, 55 (2016) 15826.
- [74] P. D'Angelo, M. Benfatto, S. Della Longa, N.V. Pavel, *Physical Review B*, 66 (2002).
- [75] J.E. Penner-Hahn, *Coordination Chemistry Reviews*, 249 (2005) 161.
- [76] A. Oda, T. Ohkubo, T. Yumura, H. Kobayashi, Y. Kuroda, *Dalton Transactions*, 44 (2015) 10038.
- [77] S.M.T. Almutairi, B. Mezari, P. Magusin, E.A. Pidko, E.J.M. Hensen, *Acs Catalysis*, 2 (2012) 71.
- [78] V. Yachandra, L. Powers, T.G. Spiro, *Journal of the American Chemical Society*, 105 (1983) 6596.
- [79] K. Peariso, Z.H.S. Zhou, A.E. Smith, R.G. Matthews, J.E. Penner-Hahn, *Biochemistry*, 40 (2001) 987.
- [80] K. Clark-Baldwin, D.L. Tierney, N. Govindaswamy, E.S. Gruff, C. Kim, J. Berg, S.A. Koch, J.E. Penner-Hahn, *Journal of the American Chemical Society*, 120 (1998) 8401.
- [81] A. Oda, H. Torigoe, A. Itadani, T. Ohkubo, T. Yumura, H. Kobayashi, Y. Kuroda, *Angewandte Chemie-International Edition*, 51 (2012) 7719.
- [82] A.L. Yakovlev, A.A. Shubin, G.M. Zhidomirov, R.A. van Santen, *Catalysis Letters*, 70 (2000) 175.

- [83] E.A. Pidko, R.A. van Santen, *The Journal of Physical Chemistry C*, 111 (2007) 2643.
- [84] Y. Xin, P. Qi, X. Duan, H. Lin, Y. Yuan, *Catalysis Letters*, 143 (2013) 798.
- [85] T.V.W. Janssens, *Journal of Catalysis*, 264 (2009) 130.
- [86] M. Bjørgen, F. Joensen, M. Spangsbørg Holm, U. Olsbye, K.-P. Lillerud, S. Svelle, *Applied Catalysis A: General*, 345 (2008) 43.
- [87] A.L. Yakovlev, A.A. Shubin, G.M. Zhidomirov, R.A. van Santen, *Catalysis Letters*, 70 (2000) 175.
- [88] Y. Ni, A. Sun, X. Wu, G. Hai, J. Hu, T. Li, G. Li, *Microporous and Mesoporous Materials*, 143 (2011) 435.
- [89] P. Meriaudeau, C. Naccache, *Journal of Molecular Catalysis*, 59 (1990) L31.
- [90] A.G. Stepanov, S.S. Arzumanov, A.A. Gabrienko, V.N. Parmon, I.I. Ivanova, D. Freude, *ChemPhysChem*, 9 (2008) 2559.
- [91] G. Buckles, G.J. Hutchings, C.D. Williams, *Catalysis Letters*, 11 (1991) 89.
- [92] L.M. Lubango, M.S. Scurrell, *Applied Catalysis A: General*, 235 (2002) 265.
- [93] C. Lamberti, A. Zecchina, E. Groppo, S. Bordiga, *Chemical Society Reviews*, 39 (2010) 4951.
- [94] D. Scarano, S. Bertarione, G. Spoto, A. Zecchina, C. Otero Areán, *Thin Solid Films*, 400 (2001) 50.